

U2 Module 2 Video Notes: Intensive vs Extensive

BIG IDEAS

DETAILS

Intensive properties	Intensive properties are things that <u>do not depend</u> on the amount of matter present. For example, two different amounts of water will both boil at <u>100°C</u> Examples of intensive properties: Boiling/melting pt. Density Hardness color Flamability combustibility QUESTION: True or false: if you break a sample of a material into pieces, the intensive properties of the material will change False																				
Extensive properties	Extensive properties are things that <u>Do depend</u> on the amount of matter present. Examples of extensive properties: volume mass length weight																				
Practice	Is it an extensive or intensive property? <table><tr><td>Volume</td><td><u>Extensive</u></td><td>Mass</td><td><u>Extensive</u></td></tr><tr><td>Density</td><td><u>Intensive</u></td><td>Flammability</td><td><u>Intensive</u></td></tr><tr><td>Corrosiveness</td><td><u>Intensive</u></td><td>Boiling Point</td><td><u>Intensive</u></td></tr><tr><td>Combustibility</td><td><u>Intensive</u></td><td>Freezing point</td><td><u>Intensive</u></td></tr><tr><td>Melting point</td><td><u>Intensive</u></td><td></td><td></td></tr></table>	Volume	<u>Extensive</u>	Mass	<u>Extensive</u>	Density	<u>Intensive</u>	Flammability	<u>Intensive</u>	Corrosiveness	<u>Intensive</u>	Boiling Point	<u>Intensive</u>	Combustibility	<u>Intensive</u>	Freezing point	<u>Intensive</u>	Melting point	<u>Intensive</u>		
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