# WELCOME TO SCHOOL

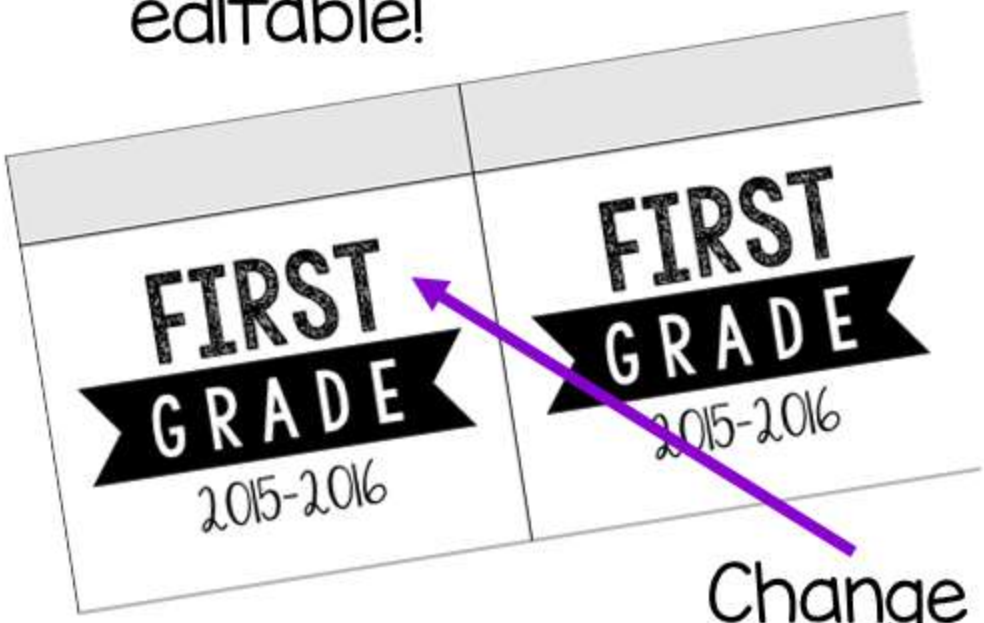
# FLIPBOOK


Completely editable!

By Jessica Vicknair
Tech and Teachability

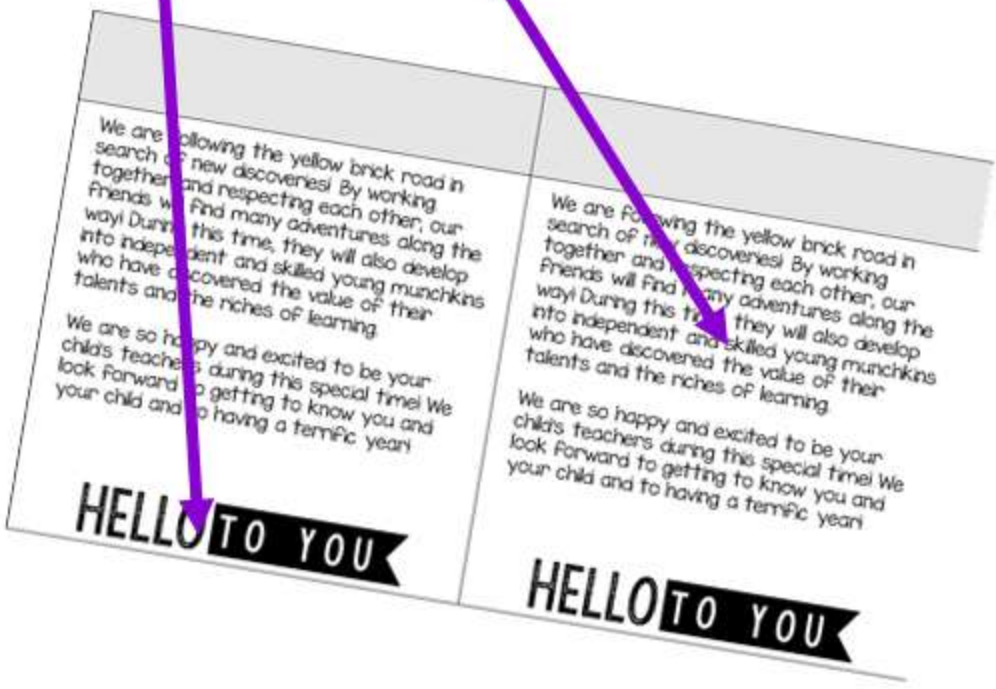
*Freebie*

Completely editable!




Change your grade

Change the text on the flaps



# Use as many flaps as you need.

There are lots of ways to get in touch with me! The most effective and quick way would be text message or Class Dojo. Below you will find all of my contact information. Please do not hesitate to get in touch with me if you have any questions.

Email: jvicknair@mayfairlabschool.com
Cell Phone: (318)792-3119
Classroom Blog: www.prideandprimary.com
Classroom Twitter: @proudpalsmist
Classroom Drive Account:
http://bit.do/mayfairlab18

School Facebook:
https://www.facebook.com/MayfairLabSchool
School phone: (225)761-7850

## CONTACT ME

| me | Monday | Tuesday | Wednesday | Thursday | Friday |
|---|---|---|---|---|---|
| | MORNING MATH/MEETING | | | | |
| | ELA - BLOCK 1 (PHONICS, SPELLING, READING GROUPS) | | | | |
| | ELA - BLOCK 2 (READING STRATEGIES, GRAMMAR LESSONS, READING GROUPS) | | | | |
| | GRAMMAR/WRITING | | | | |
| | MATH | | | | |
| | READ ALOUD | | | | |
| | LUNCH | | | | |
| | FIRST ANCILLARY | | | | FINISH IT FRIDAY |
| | RECESS | | | | |
| | DEAR | | | | |
| | SECOND ANCILLARY | | | | |
| | THIRD ANCILLARY | | | | |
| 2:45-3:00 | SCIENCE/ SOCIAL STUDIES | | | | |
| | PACK UP/READ ALOUD/DISMISSAL | | | | |

## CLASS SCHEDULE

lass Dojo as our
m behavior system. Class
os to encourage positive
om behaviors.

lso an excellent way to stay
tact with you! We can instantly
and receive messages, updates, and
os from class via the smartphone
s. It is the simplest way to get in
tact with me and to keep up with
ur child's day.

lease check out the resources at the
nk/QR code below and let me know if you
have any questions!

http://bit.do/dojoforparents

## CLASS DOJO

### Mayfair Lab Schou.

School begins promptly at 8:30 AM. A child must go to the office and receive a tardy slip after this time.

For the safety of your child, he/she must go home every day in the usual way

**UNLESS we have a NOTE from you.**

For an unexpected change, please notify the office and we will receive the message.

Ancillary - we have music, PE, French, Library, and Guidance.

Notes home - notes are sent home periodically, please make sure to check your child's OZ binder EVERY night so you are up to date on school and classroom events.

## SCHOOL PROCEDURES

Big Brown Envelope - This contains your child's graded papers and is sent home every two weeks. **Please sign and return folder and all papers.**

Book Orders - sent home monthly and are optional.

Homework - We send home weekly homework Packets on Fridays. They ontain the entire week's homework and ay be done at a time you and your d decide. They are due back the wing Friday.

ices - Please send an excuse with hild when they return. Any make k will be sent home when your urns.

Field Trips - You will be notified and sent a permission slip to sign and return for field trip.

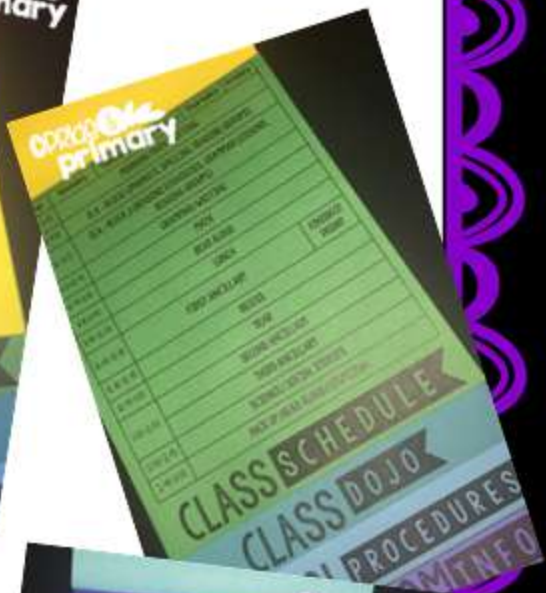
OZ Binder - This binder will keep your child Organized with Zeal! Check it everyday!

## CLASSROOM INFO

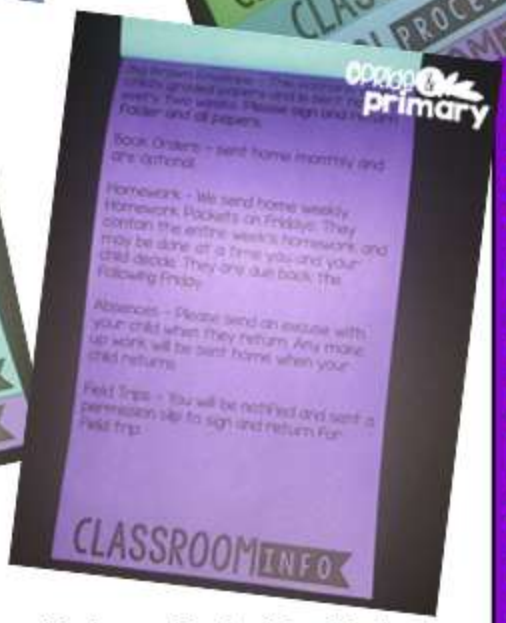
# Use Astrobrights Paper to make your flip book POP!


Fold welcome topper and staple all together.

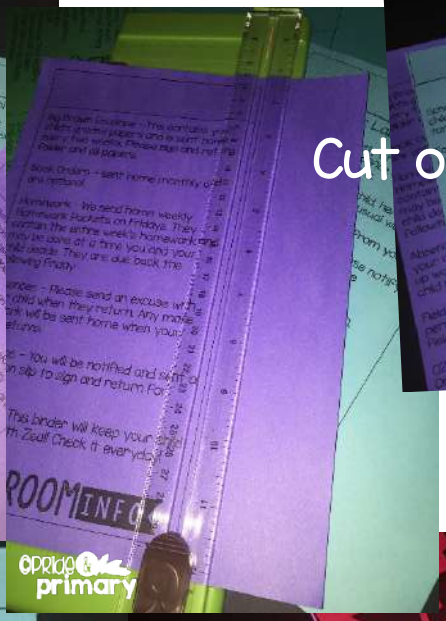




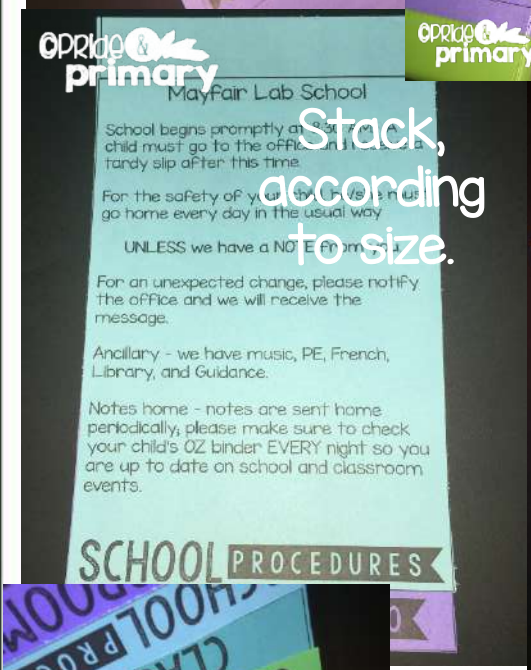



The Astrobrights colors I used: Solar Yellow, Pulsar Pink, Re-Entry Red, Orbit Orange, Gamma Green, Terrestrial Teal, Lunar Blue, Venus Violet



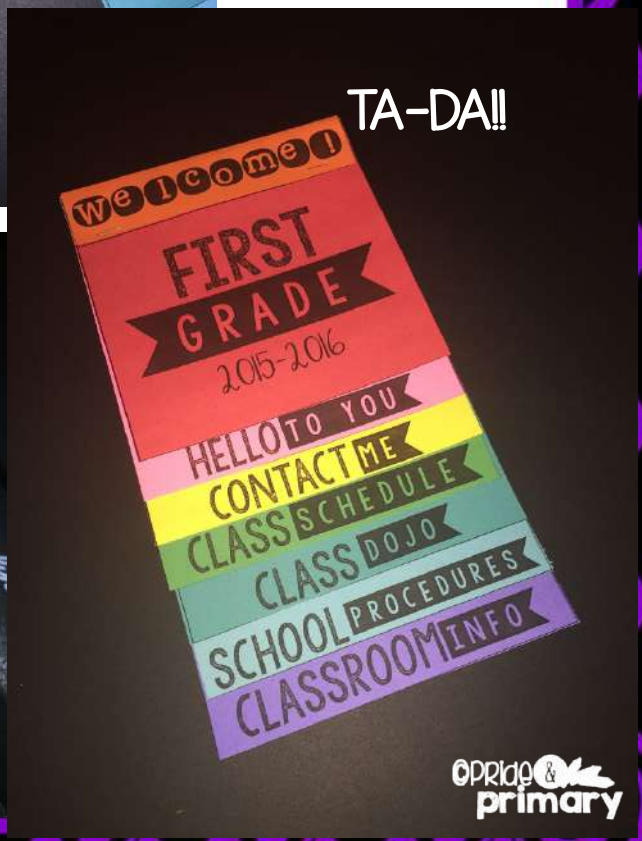



Stack, according to size.


Fold topper.



Line up, tap a few times to make even.

# The fonts I used are:

KG Miss Kindergarten

KG DEFYING GRAVITY

KG ALL OF ME

KG Eyes Wide Open

KG Wake me UP

TO USE KG DEFYING GRAVITY TO MAKE THE ] BANNER SIDES [ OR THE BLANK SPACE |

| = | the bar key

] = [ the left bracket key

[ = ] the right bracket key

THANK YOU FOR DOWNLOADING! I am so happy you want to try out my product ☺

Please test it out, make it match your classroom, and let me know by sending some feedback my way!

If you post this product on social media, please feel free to tag me (@techandteachability) so I can find you and see how you've used my products to add to your wonderful ideas!

Jessica

- **Diatomic Elements**
  $N_2$, $O_2$, $F_2$, $Cl_2$, $Br_2$, $I_2$, & $H_2$

- **Common Polyatomic Ions**
  $OH^-$ (hydroxide)
  $NO_3^{1-}$ (nitrate)
  $CO_3^{2-}$ (carbonate)
  $CrO_4^{2-}$ (chromate)
  $SO_4^{2-}$ (sulfate)
  $PO_4^{3-}$ (phosphate)
  $NH_4^{1+}$ (ammonium)

- **Common Charges**

| 1 | 2 |
|---|---|
| $Li^+$ | |
| $Na^+$ | $Mg^{2}_{+}$ |
| $K^+$ | $Ca^{2+}$ |
| $Rb^+$ | $Sr^{2+}$ |
| $Cs^+$ | $Ba^{2+}$ |

| 13 | 14 | 15 | 16 | 17 | 18 |
|---|---|---|---|---|---|
| | | $N^{3-}$ | $O^{2-}$ | $F^-$ | |
| $Al^{3+}$ | | $P^{3-}$ | $S^{2-}$ | $Cl^-$ | |
| | | | $Se^{2-}$ | $Br^-$ | |
| | | | | $I^-$ | |
| | | | | | |

EXTRAS [redacted]

- **Complete Combustion**
  Carbon Dioxide is Produced
  Blue Flame
  Hydrocarbon + $O_2 \rightarrow CO_2 + H_2O$
  $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$

- **Incomplete Combustion**
  Carbon Monoxide is Produced
  Yellow Flame
  Hydrocarbon + $O_2 \rightarrow CO + H_2O$
  $2CH_4 + 3O_2 \rightarrow 2CO + 4H_2O$

# COMBUSTION

- Also Called Metathesis

- Double Replacement

  $AX + BY \rightarrow AY + BX$

  - A & B are Cations
  - x & y are Anions

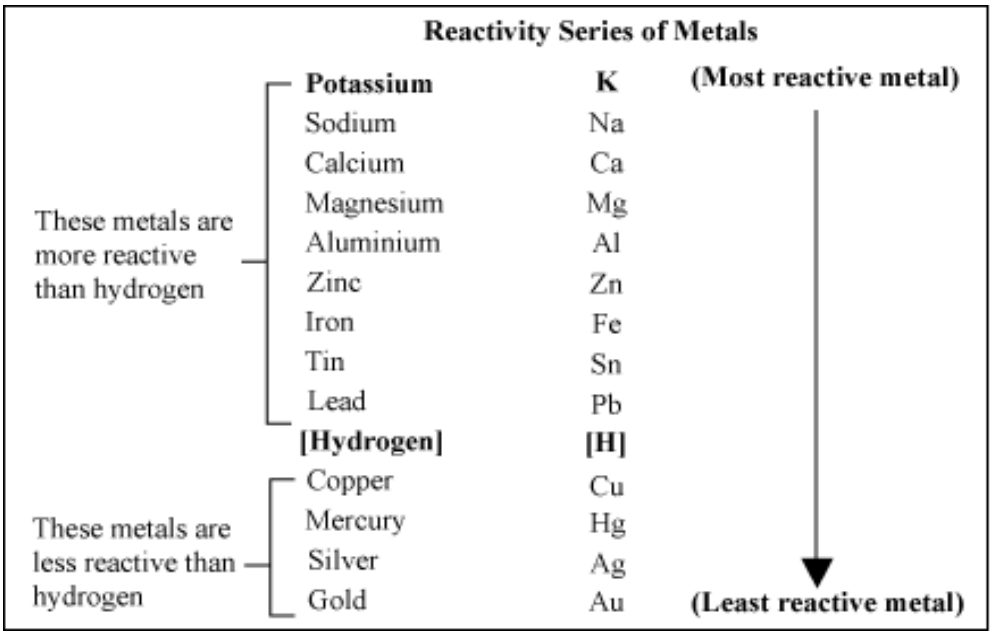
  $BaCl_2 + 2NaBr \rightarrow 2NaCl + BaBr_2$

| **Soluble Compounds (aqueous)** | **Insoluble Exceptions (solids/gases)** |
|---|---|
| **Compounds containing alkali metal ions ($Li^+$, $Na^+$, $K^+$, $Rb^+$, $Cs^+$) and the ammonium ion ($NH_4^+$)** | |
| **Nitrates ($NO_3^-$), bicarbonates ($HCO_3^-$), and chlorates ($ClO_3^-$)** | |
| **Halides ($Cl^{1-}$, $Br^{1-}$, $I^{1-}$)** | Halides of $Ag^+$, $Hg_2^{2+}$, and $Pb^{2+}$ |
| **Sulfates ($SO_4^{2-}$)** | Sulfates of $Ag^+$, $Ca^{2+}$, $Sr^{2+}$, $Ba^{2+}$, $Hg_2^{2+}$, and $Pb^{2+}$ |
| **Insoluble (solids/gases)** | Soluble Exceptions (aqueous) |
| **Carbonates ($CO_3^{2-}$), phosphates ($PO_4^{3-}$), chromates ($CrO_4^{2-}$), and sulfides ($S^{2-}$)** | Compounds containing alkali metal ions (Column 1 Elements) |
| **Hydroxides ($OH^{1-}$)** | Compounds containing alkali metal ions (Column 1 Elements) and the $Ba^{2+}$ ion |

DOUBLE-REPLACE

- **Single-Replacement with Cation**
  $A + Bx \rightarrow B + Ax$
  $H_2 + Cu(NO_3)_2 \rightarrow Cu + 2NHO_3$

- **Single-Replacement with Anion**
  $y + Ax \rightarrow x + Ay$
  $Cl_2 + 2NaBr \rightarrow Br_2 + 2NaCl$




# SINGLE-REPLACE

- **Binary Compound**
  $AX \rightarrow A + X$
  $2MgO \rightarrow 2Mg + O_2$
- **Metal Carbonate**
  Metal $CO_3 \rightarrow$ Metal Oxide + $CO_2$
  $MgCO_3 \rightarrow MgO + CO_2$
- **Metal Hydroxide**
  Metal $OH \rightarrow$ Metal Oxide + $H_2O$
  $Mg(OH)_2 \rightarrow MgO + H_2O$
- **Metal Chlorate**
  Metal $ClO_3 \rightarrow$ Metal Chloride + $O_2$
  $Mg(ClO_3)_2 \rightarrow MgCl_2 + O_2$
- **Acid**
  Acid $\rightarrow$ Nonmetal Oxide + $H_2O$
  $H_2CO_3 \rightarrow CO_2 + H_2O$

# DECOMPOSITION


- **Binary Compound**
  $AX \rightarrow A + X$
  $2MgO \rightarrow 2Mg + O_2$
- **Metal Carbonate**
  Metal $CO_3 \rightarrow$ Metal Oxide + $CO_2$
  $MgCO_3 \rightarrow MgO + CO_2$
- **Metal Hydroxide**
  Metal $OH \rightarrow$ Metal Oxide + $H_2O$
  $Mg(OH)_2 \rightarrow MgO + H_2O$
- **Metal Chlorate**
  Metal $ClO_3 \rightarrow$ Metal Chloride + $O_2$
  $Mg(ClO_3)_2 \rightarrow MgCl_2 + O_2$
- **Acid**
  Acid $\rightarrow$ Nonmetal Oxide + $H_2O$
  $H_2CO_3 \rightarrow CO_2 + H_2O$

DECOMPOSITION

- **Group A Metal + Nonmetal**
  Metal + Nonmetal → Binary Compound
  $Mg + Cl_2 \rightarrow MgCl_2$
- **Nonmetal Oxide + Water**
  Nonmetal Oxide + $H_2O \rightarrow$ Acid
  $CO_2 + H_2O \rightarrow H_2CO_3$
- **Metallic Oxide + Water**
  Metallic Oxide + $H_2O \rightarrow$ Base
  $MgO + H_2O \rightarrow Mg(OH)_2$

# SYNTHESIS


- **Group A Metal + Nonmetal**
  Metal + Nonmetal → Binary Compound
  $Mg + Cl_2 \rightarrow MgCl_2$
- **Nonmetal Oxide + Water**
  Nonmetal Oxide + $H_2O \rightarrow$ Acid
  $CO_2 + H_2O \rightarrow H_2CO_3$
- **Metallic Oxide + Water**
  Metallic Oxide + $H_2O \rightarrow$ Base
  $MgO + H_2O \rightarrow Mg(OH)_2$

SYNTHESIS

- Synthesis
  $A + X \rightarrow AX$
- Decomposition
  $AX \rightarrow A + X$
- Single-Replacement
  $A + BX \rightarrow B + AX$
- Double-Replacement
  $AX + BY \rightarrow BX + AY$
- Combustion
  Hydrocarbon + $O_2 \rightarrow H_2O + CO_2$

REACTION TYPES


- Synthesis
  $A + X \rightarrow AX$
- Decomposition
  $AX \rightarrow A + X$
- Single-Replacement
  $A + BX \rightarrow B + AX$
- Double-Replacement
  $AX + BY \rightarrow BX + AY$
- Combustion
  Hydrocarbon + $O_2 \rightarrow H_2O + CO_2$

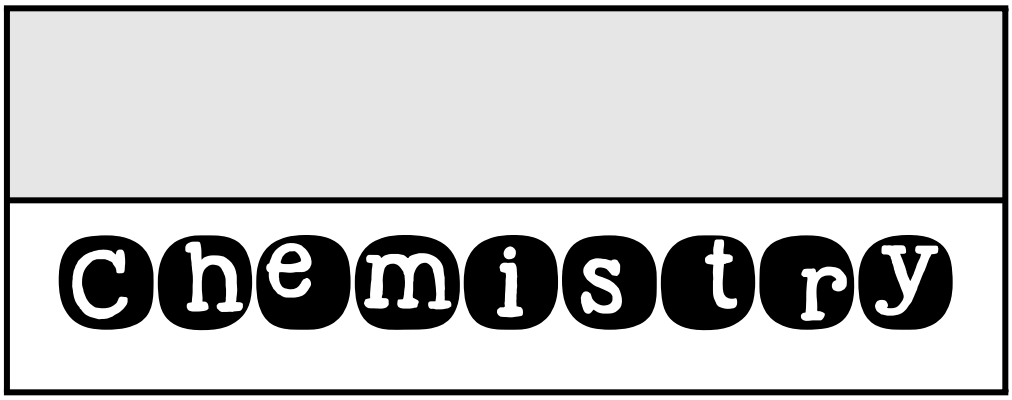
REACTION TYPES

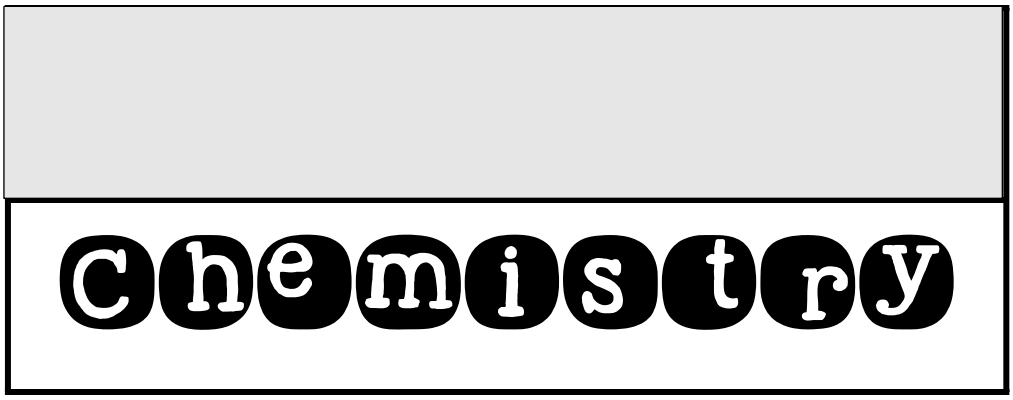
TYPES OF

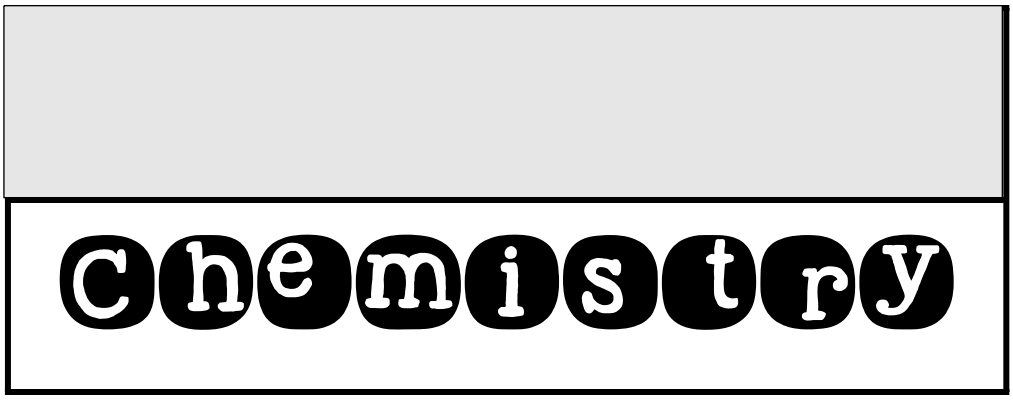
Reactions

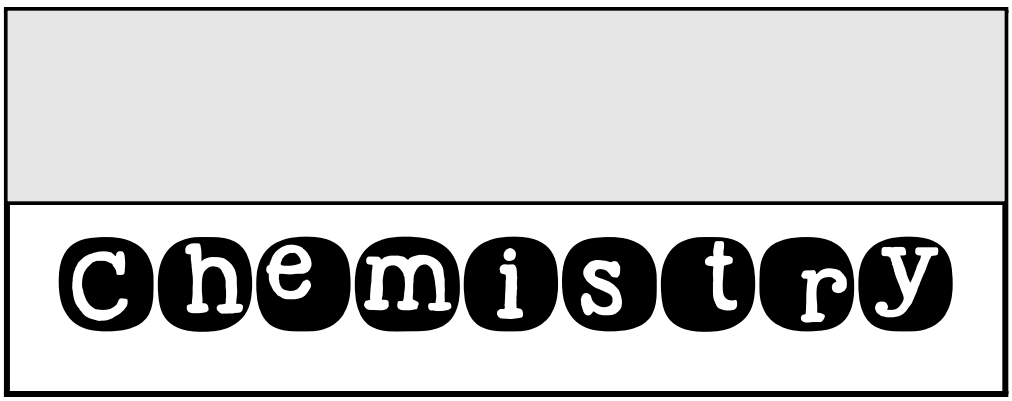
TYPES OF

Reactions


Chemistry


Chemistry


Chemistry


Chemistry

# Thank You!

Thank you so much for downloading this product. I hope you love using and implementing it into your classroom. If you have any questions at all, find an error, or just a suggestion please email me at Jessica@techandteachability.com. Please follow me on TpT to receive updates when new products are posted!

## TERMS OF USE

- Every page of this product is copyrighted.
- You have purchased a single classroom use. Please do not redistribute. If you wish to share this product with others, please visit your My Purchases page and purchase additional licenses for 50% off.
- Please do not post this document or parts of this document online (even a personal/classroom website). Doing so is a violation of the Digital Millennium Copyright Act (DMCA).
- Permission is granted to copy pages specifically designed for student or teacher use by the original purchaser or licensee. The reproduction of any other part of this product is strictly prohibited.
- Feel free to share about *your use* of this product, but credit back to my store is required. Please link back to my store.

## STAY CONNECTED

Facebook
Twitter
Instagram
Bloglovin'
Pinterest
TpT

## CREDITS TO:

SELLING STRATEGIES

CREATIVE Clips KRISTA WALLDEN

KG Fonts