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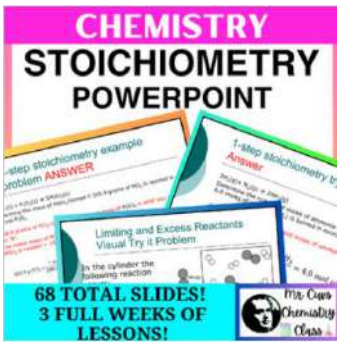
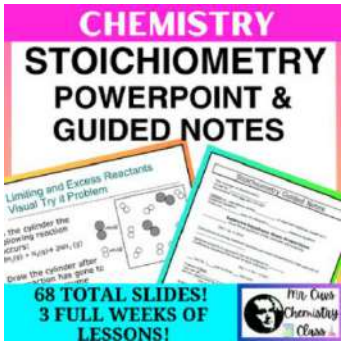
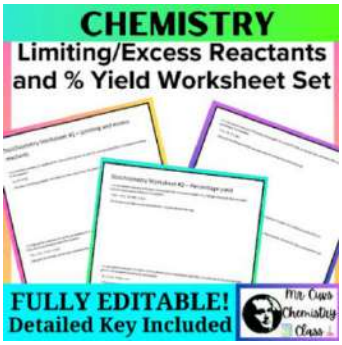
*Interested in more resources related to stoichiometry? Check out my other resources below – just click the pictures!

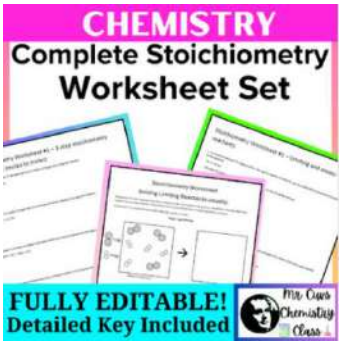
Thanks for your support and I welcome any feedback you may have.

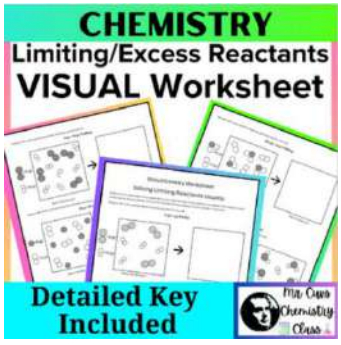
Mr Cuvs Chemistry Class

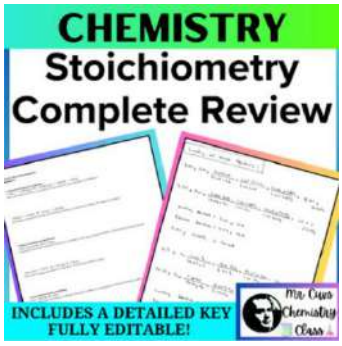


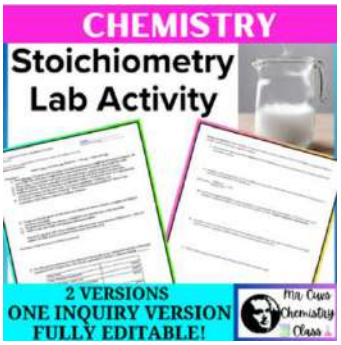


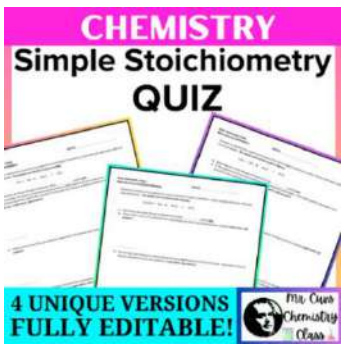


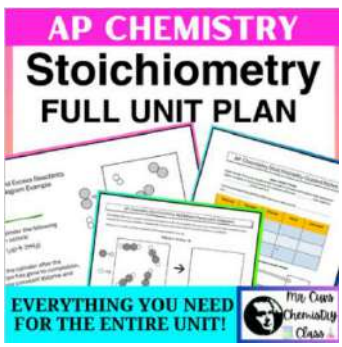


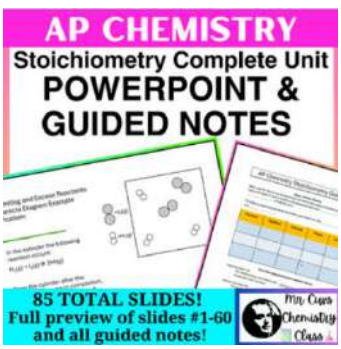




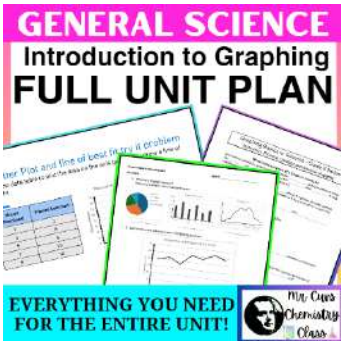


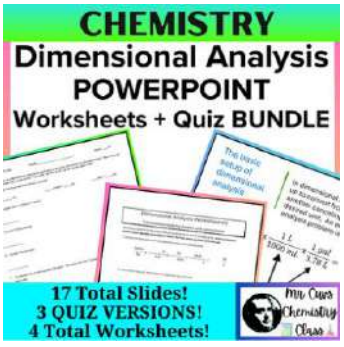


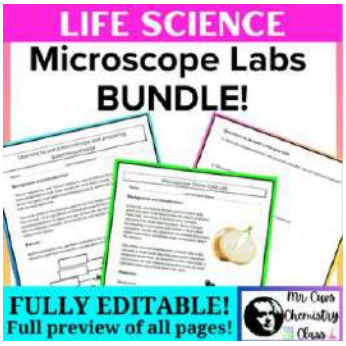


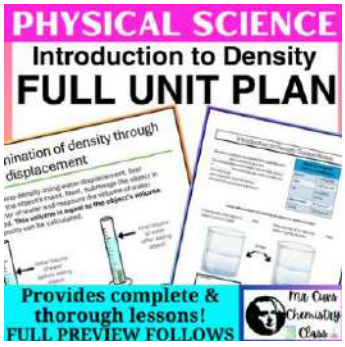


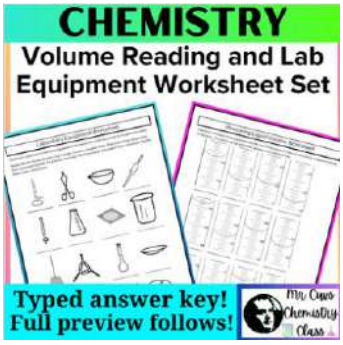
Many other great resources are available as well!
Just click the pictures below or check out my store!

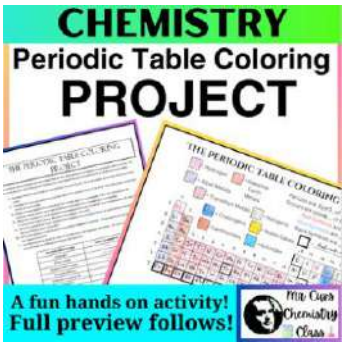


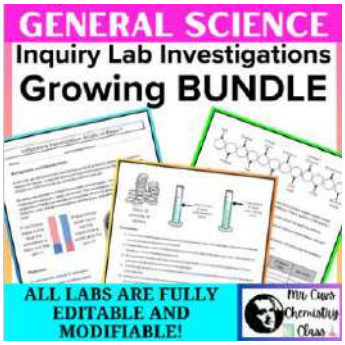


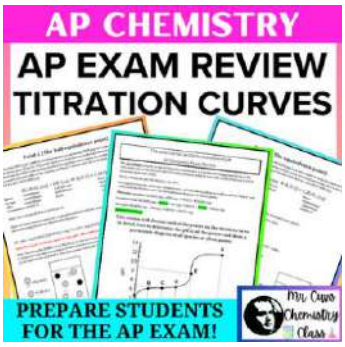


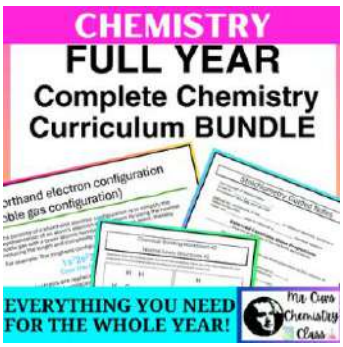


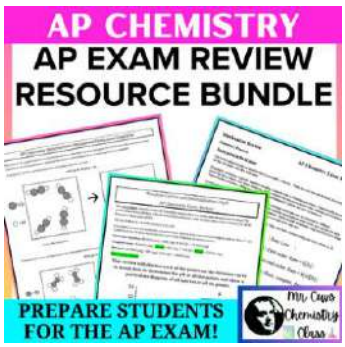


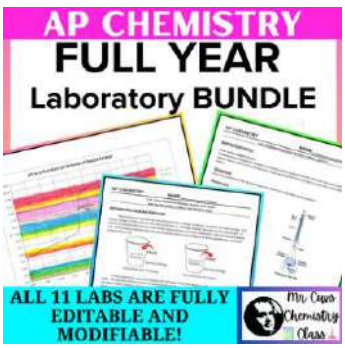












One Step Stoichiometric calculation FLOWCHART

$$given\ mol\ A\ x\ \frac{■ moles\ of\ B}{■ moles\ of\ A} = moles\ of\ B$$

Where A is the starting substance and B is what you are solving for.

■ = the stoichiometric coefficient from the balanced equation

Two Step Stoichiometric calculations FLOWCHART

$$given\ mol\ A\ x\ \frac{■ moles\ of\ B}{■ moles\ of\ A}\ x\ \frac{molar\ mass\ of\ B}{1\ mol\ B} = grams\ of\ B$$

Or

$$mass\ of\ A\ x \frac{1\ mol\ A}{molar\ mass\ of\ A}\ x \frac{■ moles\ of\ B}{■ moles\ of\ A} = moles\ of\ B$$

Three Step Stoichiometric calculations FLOWCHART

$$mass\ of\ A\ x \frac{1\ mol\ A}{molar\ mass\ of\ A}\ x \frac{■ moles\ of\ B}{■ moles\ of\ A}\ x\ \frac{molar\ mass\ of\ B}{1\ mol\ B} = grams\ of\ B$$

Tip: There is always a “1” attached to the molar mass steps