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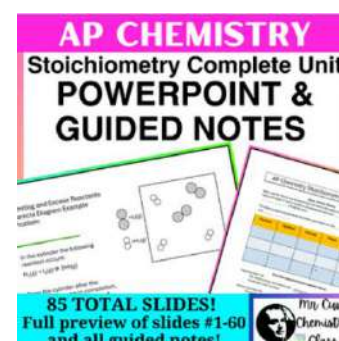
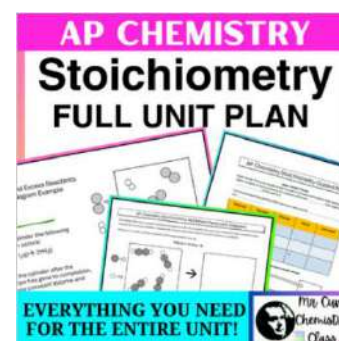
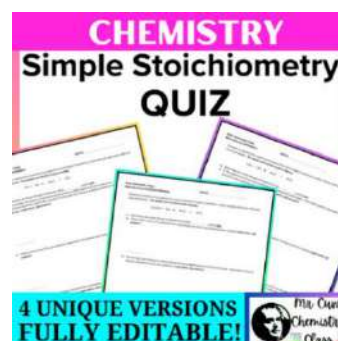
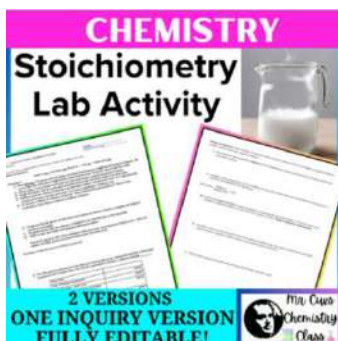
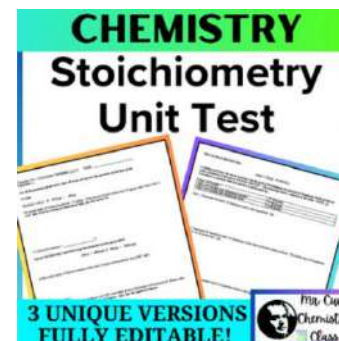
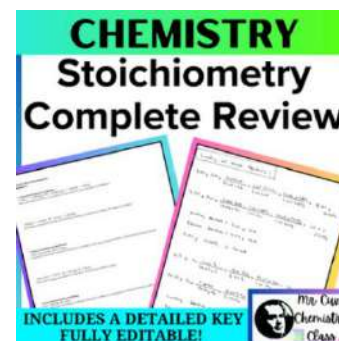
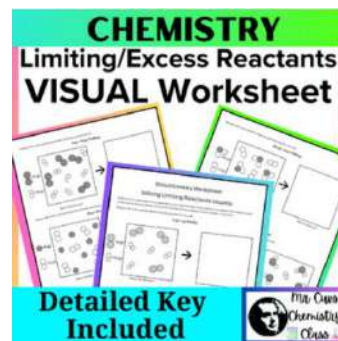
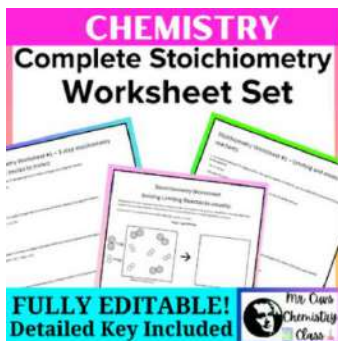
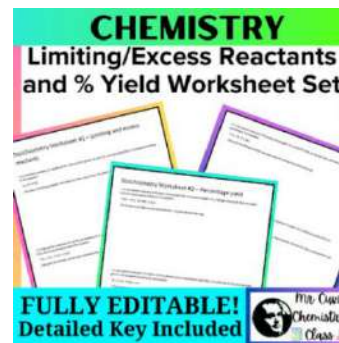
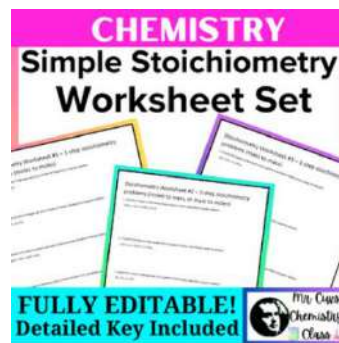
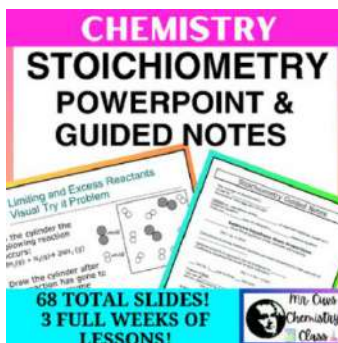
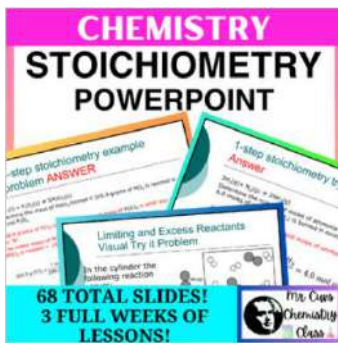
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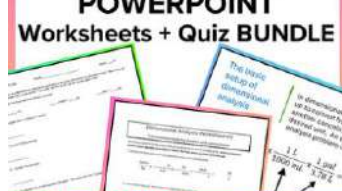
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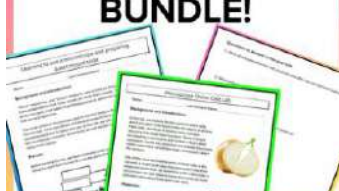
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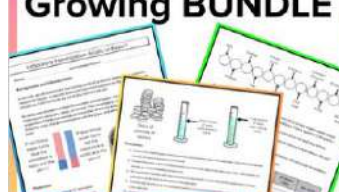
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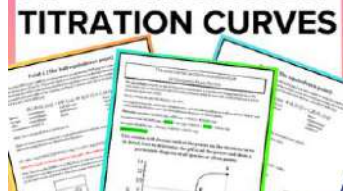
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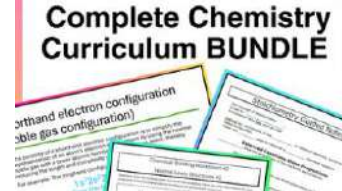
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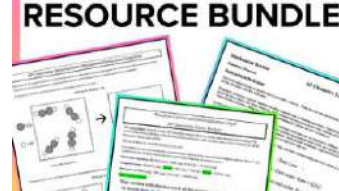
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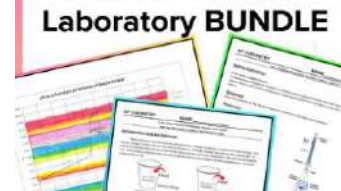
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## One Step Stoichiometric calculation FLOWCHART

$$\text{given mol A} \times \frac{\blacksquare \text{ moles of B}}{\blacksquare \text{ moles of A}} = \text{moles of B}$$

Where A is the starting substance and B is what you are solving for.

■ = the stoichiometric coefficient from the balanced equation

## Two Step Stoichiometric calculations FLOWCHART

$$\text{given mol A} \times \frac{\blacksquare \text{ moles of B}}{\blacksquare \text{ moles of A}} \times \frac{\text{molar mass of B}}{1 \text{ mol B}} = \text{grams of B}$$

Or

$$\text{mass of A} \times \frac{1 \text{ mol A}}{\text{molar mass of A}} \times \frac{\blacksquare \text{ moles of B}}{\blacksquare \text{ moles of A}} = \text{moles of B}$$

## Three Step Stoichiometric calculations FLOWCHART

$$\text{mass of A} \times \frac{1 \text{ mol A}}{\text{molar mass of A}} \times \frac{\blacksquare \text{ moles of B}}{\blacksquare \text{ moles of A}} \times \frac{\text{molar mass of B}}{1 \text{ mol B}} = \text{grams of B}$$



Tip: There is always a “1” attached to the molar mass steps