

MM	2.016	31.998	18.015
Balanced Eqn	2H_2	$+ \text{O}_2 \rightarrow$	$2\text{H}_2\text{O}$
coefficient- (Mol/Mol Ratio)	2	1	2
Moles	14.67		14.67
given amount	? g 29.57g		264.36g

Example problem:
grams

How many ~~moles~~
of hydrogen gas
does it take to produce
264.36 g H_2O ?

Step 1 - identify given & unknown

2 - convert given to moles (if not already)

3 - use mole ratio to determine mole of unknown

4 - final answer