

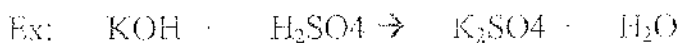
Solubility Rules Worksheet

1. Name or give the chemical formula for each of the following compounds.
2. State whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution. Use solubility rules.

Chemical Formula	Name	Solubility
1. $\text{NH}_4\text{CH}_3\text{COO}$		
2. $\text{Ba}(\text{OH})_2$		
3. _____	Iron (III) Carbonate	
4. NaOH		
5. RbNO_3		
6. _____	Cesium Sulfate	
7. MgSO_4		
8. ZnCl_2		
9. _____	Zinc Hydroxide	
10. $\text{Zn}_3(\text{PO}_4)_2$		
11. AgBr		
12. KNO_3		
13. Al_2S_3		
14. _____	Silver Acetate	
15. Sr_2CrO_4		
16. _____	Aluminum Phosphate	
17. BaSO_4		
18. $\text{Ca}(\text{OH})_2$		
19. BaCO_3		
20. MgCO_3		
21. _____	Iron (III) sulfide	
22. $\text{Sn}(\text{OH})_2$		
23. _____	Silver Iodide	
24. Hg_2SO_4		
25. _____	Mercury Chloride	

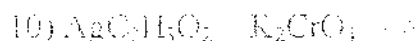
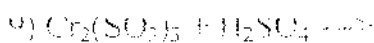
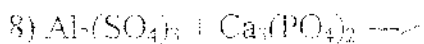
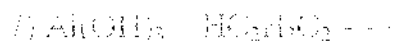
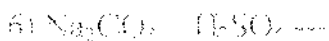
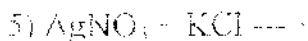
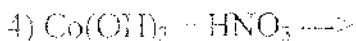
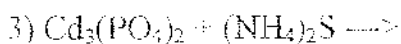
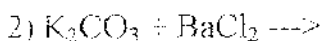
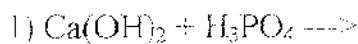
Double Replacement Reactions

The Cations and Anions switch partners : $AB + XY \rightarrow AY + BX$



Check the oxidation number to calculate the correct number of atoms.

Write correct formulas for the products in these double replacement reactions



Oxidation
States:

Ca^{+2}	OH^{-1}
K^{+1}	Pb_4
H^{+1}	CO_3^{-2}
Ba^{+2}	Cl^{-1}
Cd^{+2}	S^{-2}
Co^{+2}	NO_3^{-1}
Ag^{+1}	SO_4^{-2}
N_4^{+1}	
Al^{+3}	$C_2H_3O_2^{-}$
Cr^{+3}	SO_3^{-2}
NH_4^{+1}	CrO_4^{-2}