



Chapter Review

Properties of Atoms and the Periodic Table

Part A. Vocabulary Review

Directions: On the space at the left, write the term that correctly completes each statement. Use each term once.

metals

isotopes

average atomic mass

electron cloud

groups

metalloids

transition elements

atomic number

electrons

nucleus

mass number

periods

chemical symbol

quarks

periodic table

chemical symbol

periods

average atomic mass

groups

transition elements

atomic number

quarks

nucleus

mass number

isotopes

metals

metalloids

electrons

electron cloud

periodic table

1. A capital letter or a combination of a capital letter and a small letter that is used to represent an element is called a(n) _____.

2. The horizontal rows of elements are called _____.

3. An average of the masses of all the isotopes that occur in nature for an element is the _____.

4. Vertical columns of elements are called _____.

5. Elements in the middle of the periodic table, groups 4 through 7, are called the _____.

6. The number of protons in an atom is the _____.

7. Protons and neutrons can be subdivided into _____ by colliding them.

8. The center of an atom where protons and neutrons are located is the _____.

9. A total count of the neutrons and protons in an atom is the _____.

10. Atoms of the same element but with different numbers of neutrons are _____.

11. Elements that are found on the left side of the periodic table are _____.

12. Elements that have some properties of both metals and nonmetals are _____.

13. The particles that move about the nucleus and have a negative charge are _____.

14. The region around the nucleus occupied by electrons is a(n) _____.

15. A chart that shows the classification of elements is called the _____.

Part B. Concept Review

Name _____ Date _____ Class _____

Directions: Refer to the periodic table below and the boxes at the right of the table to answer questions 11–15.

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11 Na 23	26 Fe 55.8	79 Au 197
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