## **Practice Quiz - Periodic Table**

1.	A Lewis dot (e- d a) atomic mass c) number of vale	ot) structure shows th ence electrons	e for	an element b) number of pro d) number of neu	otons utrons		
2.	A column on the a) row	periodic table is know b) period c) cl	n as a ass	d) group or fami	ly		
3.	Atoms are arrang a) mass b	ed in order on the peri ) atomic number	odic table, c) nu	based on mber of neutrons	d) ato	mic radius	
4.	In the diagram to number of dots sh a) Sulfur b) Carbon	the right, X represents nown, which of the fol c) Fl d) C	s a certain lowing cou uorine alcium	element. Based o ild be element X?	n the	X	
5.	Based on the diag a) lose 6 electrons	gram in question 4, ho b) gain 6 electro	w would el ons c)	ement X tend to f lose 2 electrons	form bonds? d) gain 2	electrons	
6.	Alkaline earth me a) 1A	etals are found in what b) 2A c) 7A	column o A	f the periodic table d) 8A	e?		
7.	Which of these el a) Calcium	ements should have p b) Carbon	roperties m	nost similar to thos c) Boron	se of Berylliur d) Lit	n? hium	
8.	Groups possess elements with a) the same number of electrons b) the same number of protons			<ul><li>c) similar properties</li><li>d) the same number of neutrons</li></ul>			
9.	When an atom gains an electron, it becomes a a) heavy isotope b) light isotope			c) positive ion	d) neg	gative ion	
10.	How many electrona) 7	ons does Phosphorus l b) 5 c) 3	nave in its	outer shell? d) 1			
11.	The scientist who a) Galileo	proposed the modern b) Mendeleev	periodic ta	able be arranged b c) Moseley	y atomic num d) Ne	ber was wton	
12.	The group "B" elements a) can have varying oxidation numbersb) are also called transition metalsc) both answers						
13.	Which element is a) Krypton	an example of a noble b) Chlorine	e gas? c) Iridi	um	d) Nitrogen		

14.	Valence electron a) nucleus	s are electrons in the b) outer shell	c) inner shell	d) ionizat	tion level				
15.	Which of these i a) 5	s the most stable number b) 6	r of electrons to h c) 7	have in an outer she d) 8	ll (valence)?				
16.	Which elements a) halogens	generally form ions by g b) noble gases	giving up one ele c) alkali metal	ctron? s d) transit	ion metals				
17.	Non-metals are f a) top	found in what area of the b) bottom	e periodic table? c) left side	d) right s	ide				
18.	An alkali metal l a) none	nas one valence electron, b) two	, while an alkalin c) three	e earth metal has d) four					
19.	Which list of elements contains only metals?a) helium, carbon, goldc) iodine, iron, nickelb) sodium, chromium, copperd) phosphorus, nitrogen, oxygen								
20.	As you move lef a) decreases	t to right across a period b) stays the same	, the number of v c) increases	valence electrons					
21.	Which elements are the "staircase elements, sharing metallic and non-metallic properties?a) metalloidsb) halogensc) chalcogensd) transition elements								
22.	The scientist who proposed the periodic table be arranged by atomic mass wasa) Galileob) Mendeleevc) Moseleyd) Newton								
23.	Neutral elements a) alkali metals	s which have complete o b) transition m	uter electron she	lls are in the group c) halogens	called d) noble gases				
24.	Which element has a) Ar	two electrons in its vale b) P c) Ba	ence shell? d) Si						
25.	25. The name of the Group 7A family of elements is called								
26.	An atom with eight	t electrons in the valence gas.	e shell is very stal	ble. This element w	rould be called a				
27.	Look at period 2 or ionization energ a) increases	n the Periodic Table. As y from left to right? b) decreases	the atomic numb c) Ren	er increases, what l	happens to the overall				
28.	The atomic number of the Group 5A elements that is also in the 4 <sup>th</sup> period is								
29.	Elements with similar properties has a similar arrangement of electrons.								
30.	In general, as non-	metals form ions, they te	end to	electrons.					

31. How many periods does the periodic table have? \_\_\_\_\_