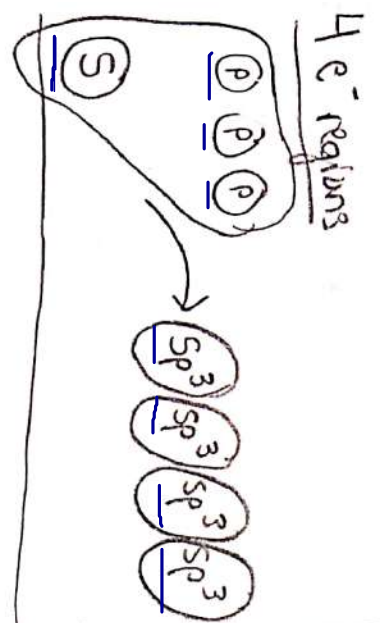
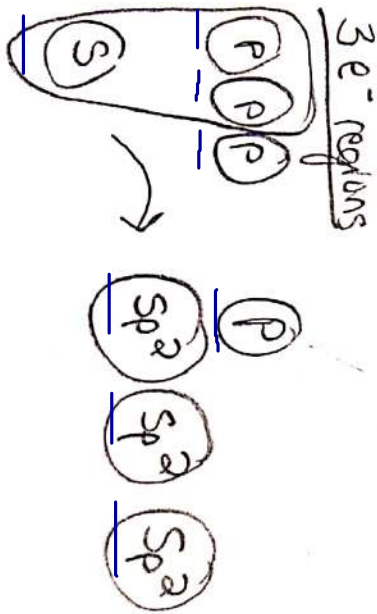




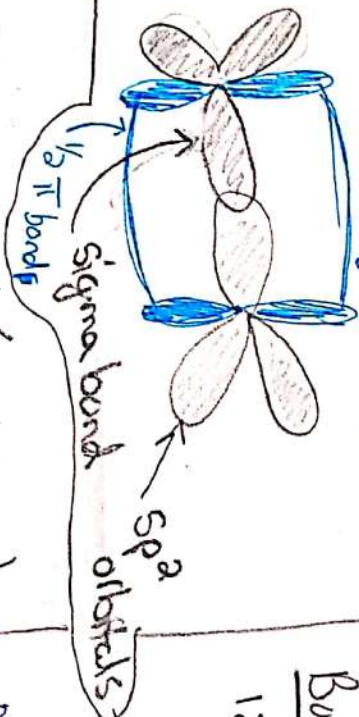
we get: 4 equal energy
sp³ orbitals



molecular geometry:
0 LP = tetrahedral
1 LP = trigonal pyramidal
2 LP = bent
bond angles
109.5°
* all single bonds and/or L.P.

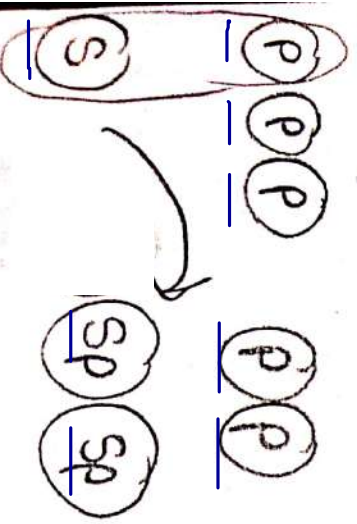


we get: 3 equal energy sp²
orbitals, 1 unhybridized p
orbital

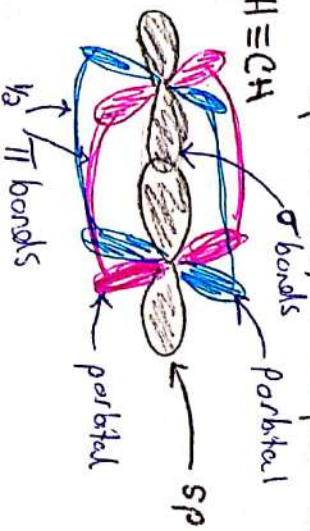
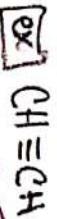


molecular geometry
0 LP = trigonal planar
1 LP = bent
Bond angles
120°
* double bond

2 e⁻ regions



we get: 2 hybridized (equal energy)
sp orbitals, 2 unhybridized p orb.



molecular geometry
linear
bond angles
180°
* Triple bond