Canadian HS Chemistry: Naming Compounds Review

Each page has 10 compounds. Write the formula for the first 5 and give the name for the rest. Click the mouse to go to the next compound.

Created by Mr. Hood (November 2005)

Page 1-Binary Ionic Compounds

- sodium chloride
- calcium flouride
- copper (I) sulfide
- iron (III) oxide
- magnesium oxide

- NaCl
- CaF_2
- Cu_2S
- Fe_2O_3
- MgO

- Na_2O • sodium oxide
- CaS • calcium sulfide
- $AlCl_3$ • aluminum chloride
- SrF_2 • strontium flouride
- Ca_3P_2 • calcium phosphide

Page 2-Binary Ionic Compounds

- potassium chloride • KCl
- calcium nitride • Ca_3N_2
- copper (II) oxide • CuO
- iron (II) chloride • $FeCl_2$
- sodium oxide • Na_2O
- NaF • sodium flouride
- BaS • barium sulfide
- Al_2O_3 • aluminum oxide
- PbO_2 • lead (IV) oxide
- FeN • iron (III) nitride

Page 3-Polyatomic Ions

- potassium nitrate • KNO_3
- calcium hydroxide • $Ca(OH)_2$
- copper (II) sulfate • $CuSO_4$
- iron (II) nitrite • $Fe(NO_2)_2$
- sodium bicarbonate or sodium hydrogen carbonate • $NaHCO_3$
- NH_4F • ammonium flouride
- KCN • potassium cyanide
- Na_2O_2 • sodium peroxide
- $Pb(NO_3)_2$ • lead (II) nitrate
- $Fe(ClO)_2$ • iron (II) hypochlorite

Page 4-Polyatomic Ions

- Sodium acetate — $NaCH_3COO$
- calcium bromate — $Ca(BrO_3)_2$
- copper (I) perchlorate — $CuClO_4$
- iron (III) dichromate — $Fe_2(Cr_2O_7)_3$
- potassium bisulfate or potassium hydrogen sulfate — $KHSO_4$
- $Ca_3(PO_4)_2$ — calcium phosphate
- MgC_2O_4 — magnesium oxalate
- H_2O_2 — hydrogen peroxide
- NH_4Br — ammonium bromide
- $Fe(ClO_2)_3$ — iron (III) chlorite

Page 5-Covalent Compounds

- carbon monoxide
- dinitrogen monoxide
- diarsenic trioxide
- carbon tetrachloride
- tetraphosphorous decoxide
- SO_3
- CO_2
- PCl_5
- N_2O_5
- SF_6

- CO
- N_2O
- As_2O_3
- CCl_4
- P_4O_{10}
- sulfur trioxide
- carbon dioxide
- phosphorous pentachloride
- dinitrogen pentoxide
- sulfur hexaflouride

Page 6-Mixed Review

- potassium perchlorate
- $KClO_4$
- ammonium nitrite
- NH_4NO_2
- lead (II) flouride
- PbF_2
- strontium chloride
- $SrCl_2$
- sodium dihydrogen phosphate
- NaH_2PO_4
- SO_2
- sulfur dioxide
- CaO_2
- Calcium peroxide
- Na_2CO_3
- sodium carbonate
- $PbSO_4$
- lead (II) sulfate
- $NaClO_3$
- sodium chlorate

Page 7-Mixed Review

- ammonium phosphate
- $(NH_4)_3PO_4$
- sodium sulfite
- Na_2SO_3
- beryllium chloride
- $BeCl_2$
- strontium carbonate
- $SrCO_3$
- calcium hydrogen phosphate or calcium biphosphate
- $CaHPO_4$
- SO_3
- sulfur trioxide
- K_2CrO_4
- potassium chromate
- $MgCO_3$
- magnesium carbonate
- Cu_3PO_4
- copper (I) phosphate
- $Ba(ClO_2)_2$
- barium chlorite

Page 8-Mixed Review

- dinitrogen tetroxide
- N_2O_4
- aluminum sulfate
- $Al_2(SO_4)_3$
- Chromium (III) bromide
- $CrBr_3$
- lithium chloride
- LiCl
- calcium bicarbonate or calcium hydrogen carbonate
- $Ca(HCO_3)_2$
- PCl_3
- phosphorous trichloride
- $K_2Cr_2O_7$
- potassium dichromate
- $KMnO_4$
- Potassium permanganate
- CuS
- copper (II) sulfide
- $Ba(Na)_2$
- Banana ☺

Polyatomic Ions-Review

- hydroxide • OH^-
- acetate • CH_3COO^-
- cyanide • CN^-
- nitrate • NO_3^-
- nitrite • NO_2^-
- bromate • BrO_3^-
- bicarbonate (hydrogen carbonate) • HCO_3^-
- dihydrogen phosphate • $H_2PO_4^-$
- perchlorate • ClO_4^-
- chlorate • ClO_3^-
- chlorite • ClO_2^-
- hypochlorite • ClO^-