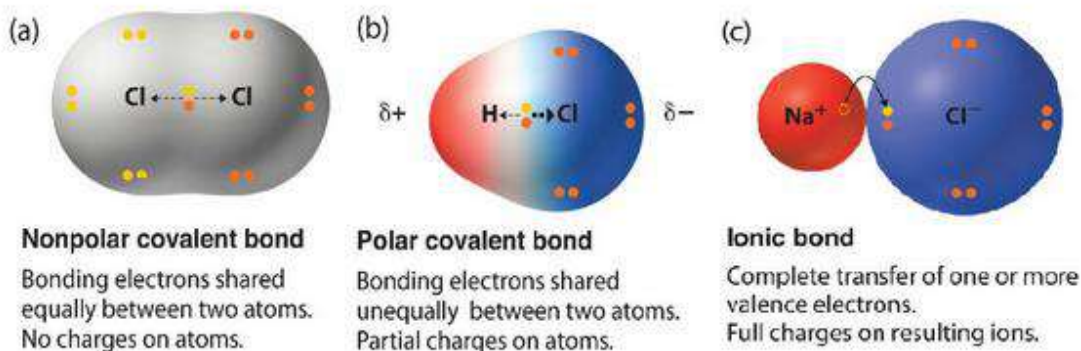
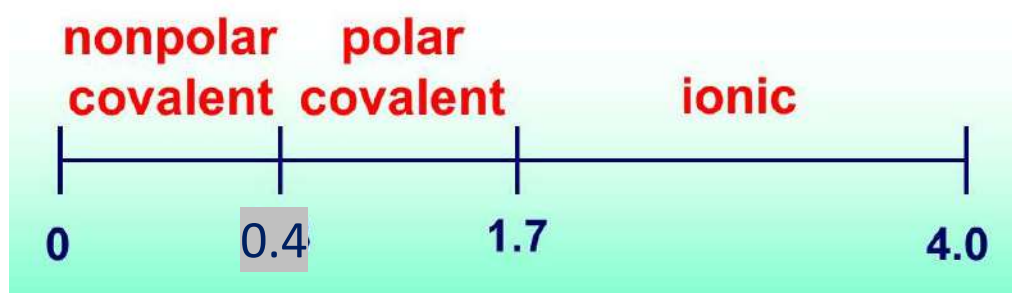


Molecular Geometry

# of bonding groups/domains on 'central' atom	# of lone pair electrons on 'central' atom	Example	Molecular Geometry	Bond Angle
2	0	CO ₂	linear	180
3	0	CH ₂ O	trigonal planar	120
2	1	SF ₂	bent	less than 120
4	0	CH ₄	tetrahedral	109.5
3	1	NH ₃	trigonal pyramidal	less than 109.5
2	2	H ₂ O	bent	less than 109.5

Electronegativity Difference



0 to 0.39 = nonpolar covalent bond
 0.40 to <1.7 = polar covalent bond
 >1.7 = ionic bond