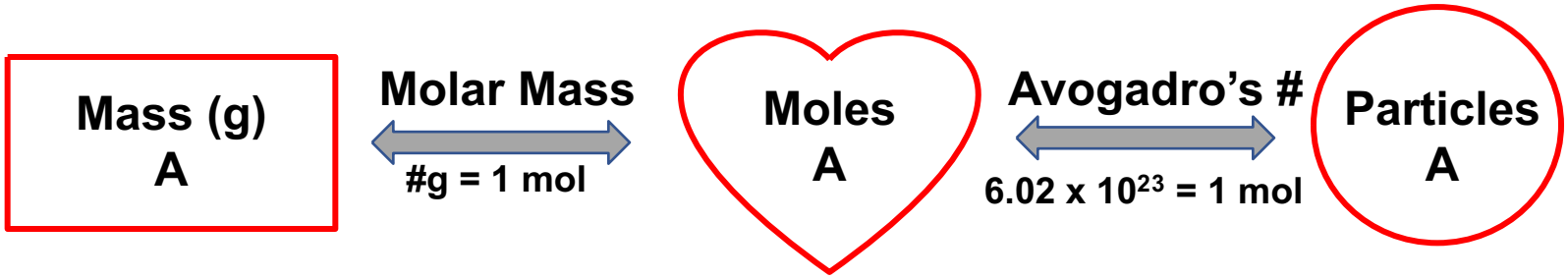
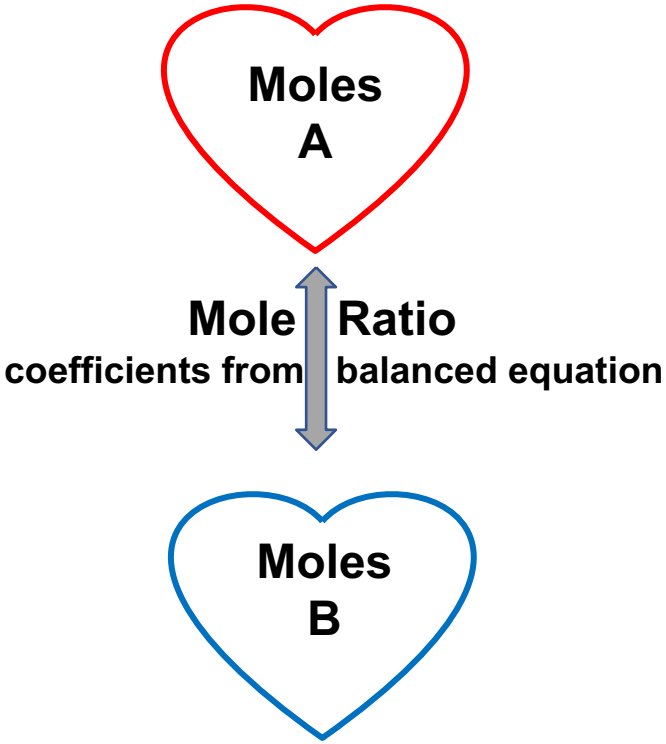
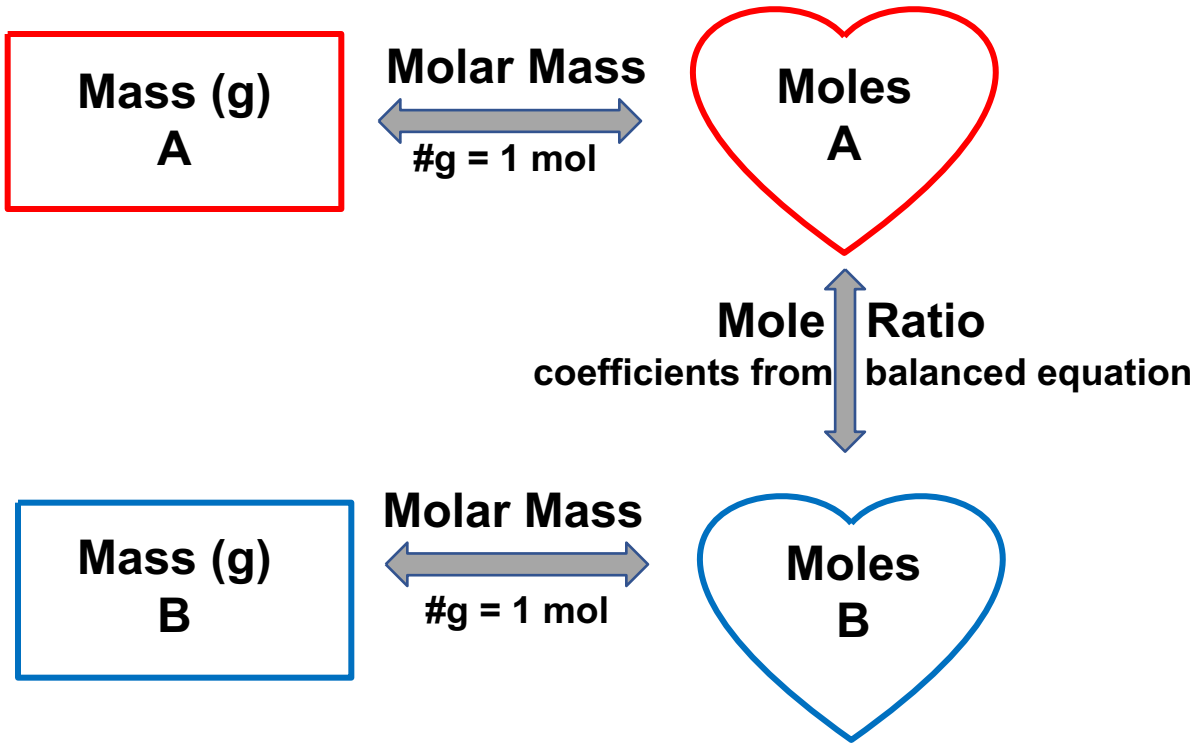
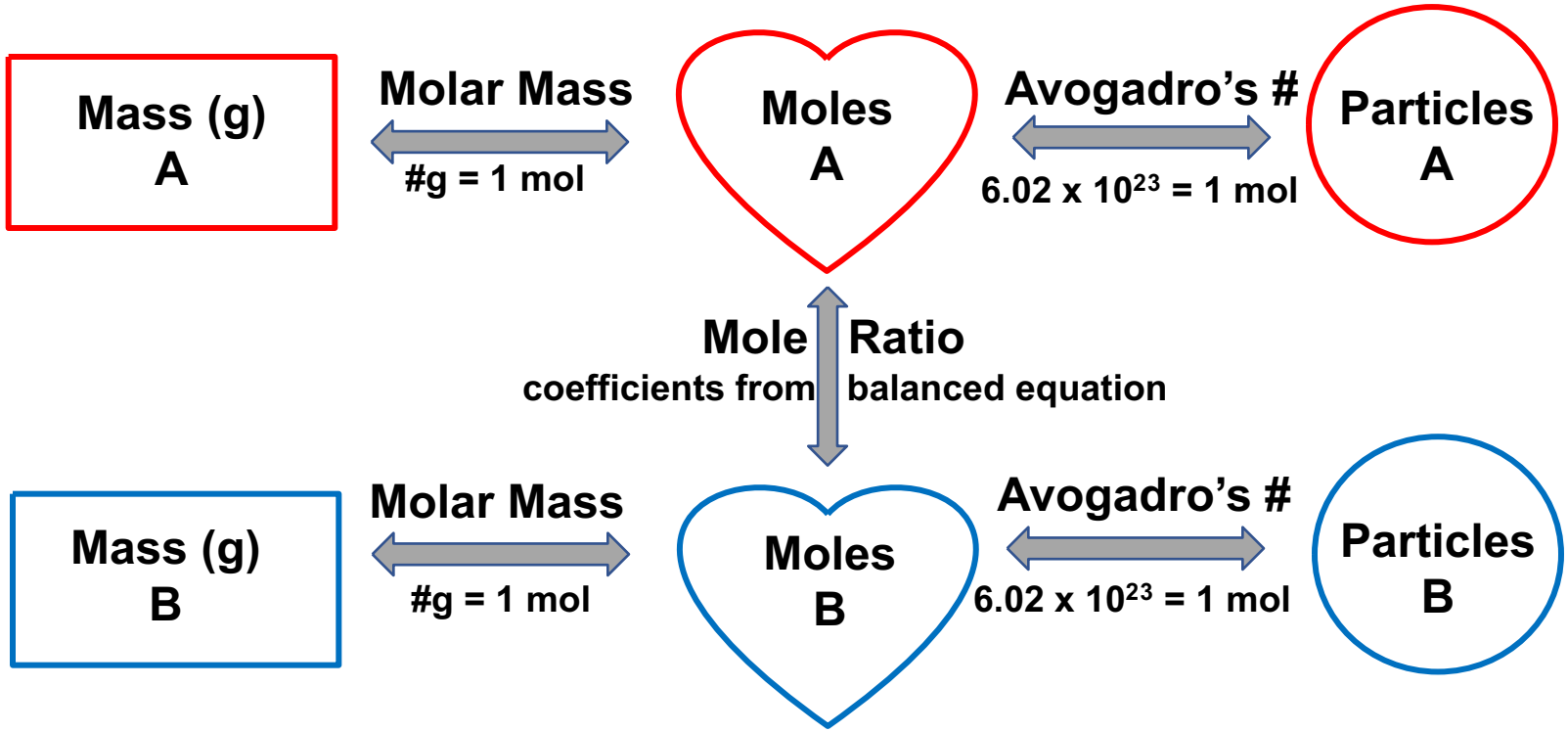
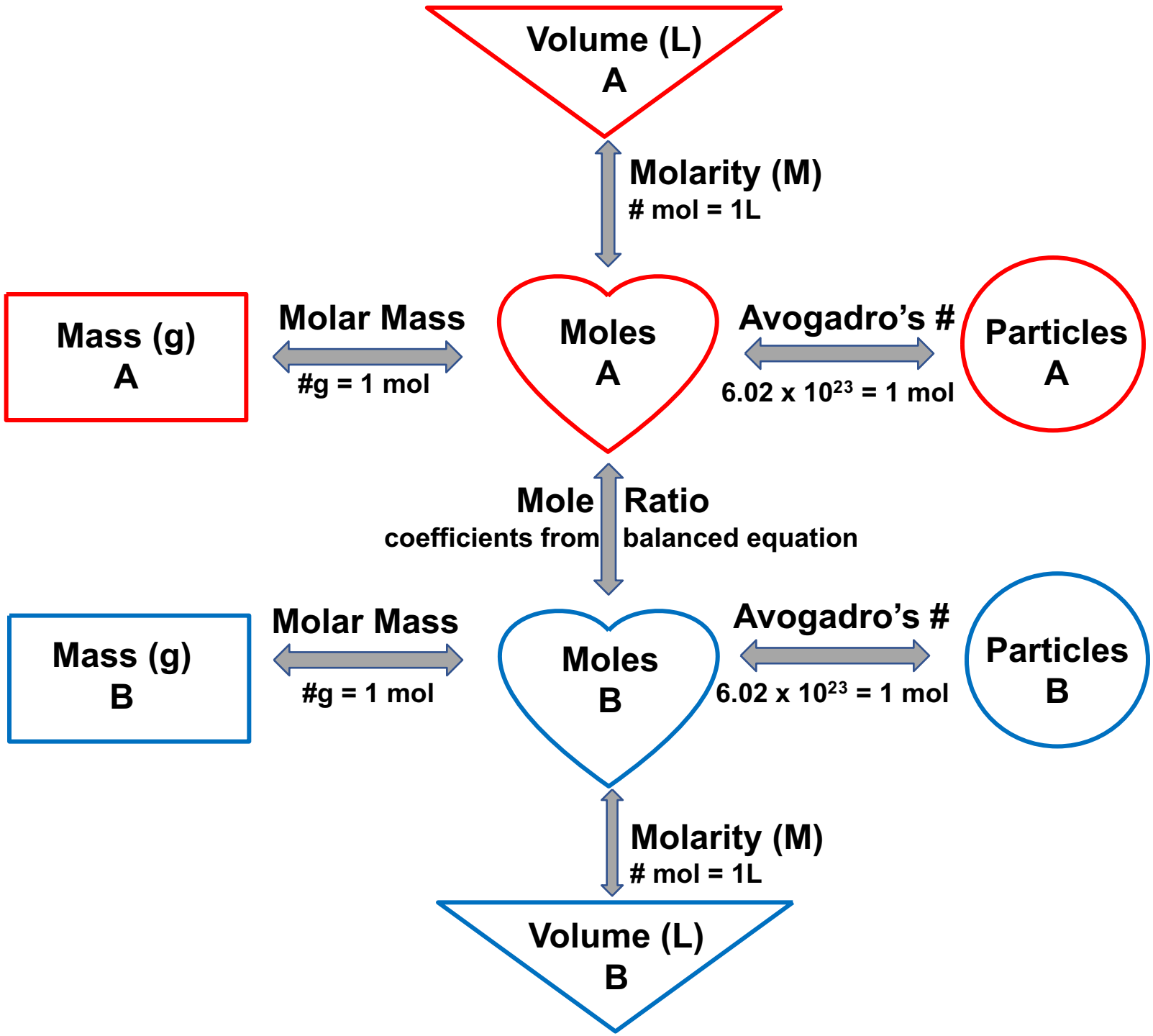
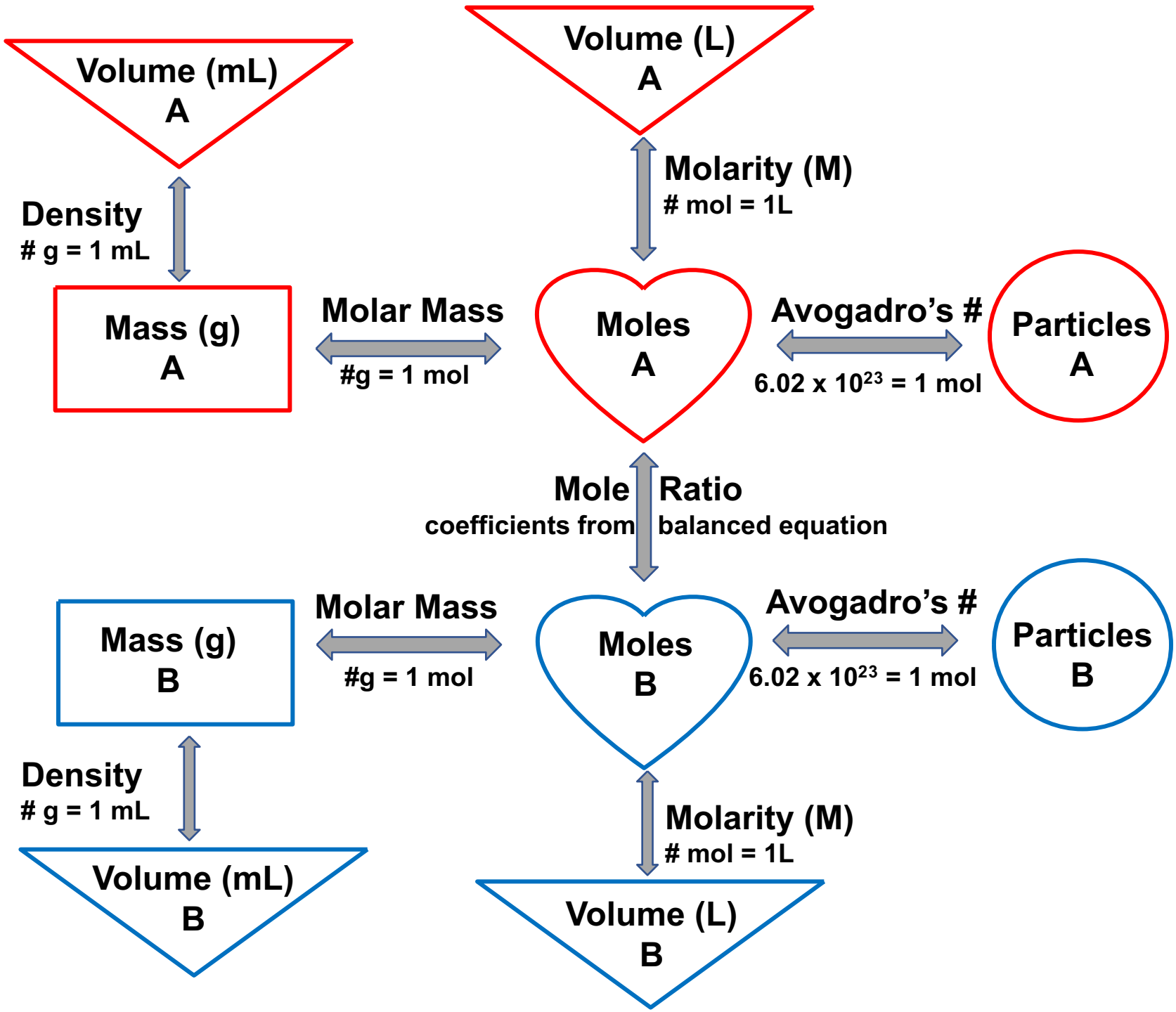

Mass (g)
A
Molar Mass
#g = 1 mol
Moles
A
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
A


Moles
A
Mole Ratio
coefficients from balanced equation
Moles
B


Mass (g)
A
Molar Mass
#g = 1 mol
Moles
A
Mole Ratio
coefficients from balanced equation
Mass (g)
B
Molar Mass
#g = 1 mol
Moles
B


Mass (g)
A
Molar Mass
#g = 1 mol
Moles
A
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
A
Mole Ratio
coefficients from balanced equation
Mass (g)
B
Molar Mass
#g = 1 mol
Moles
B
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
B


Volume (L)
A
Molarity (M)
mol = 1L
Mass (g)
A
Molar Mass
#g = 1 mol
Moles
A
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
A
Mole Ratio
coefficients from balanced equation
Mass (g)
B
Molar Mass
#g = 1 mol
Moles
B
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
B
Molarity (M)
mol = 1L
Volume (L)
B


Volume (mL)
A
Volume (L)
A
Density
g = 1 mL
Molarity (M)
mol = 1L
Mass (g)
A
Molar Mass
#g = 1 mol
Moles
A
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
A
Mole Ratio
coefficients from balanced equation
Mass (g)
B
Molar Mass
#g = 1 mol
Moles
B
Avogadro's #
6.02 x 10^23 = 1 mol
Particles
B
Density
g = 1 mL
Molarity (M)
mol = 1L
Volume (mL)
B
Volume (L)
B