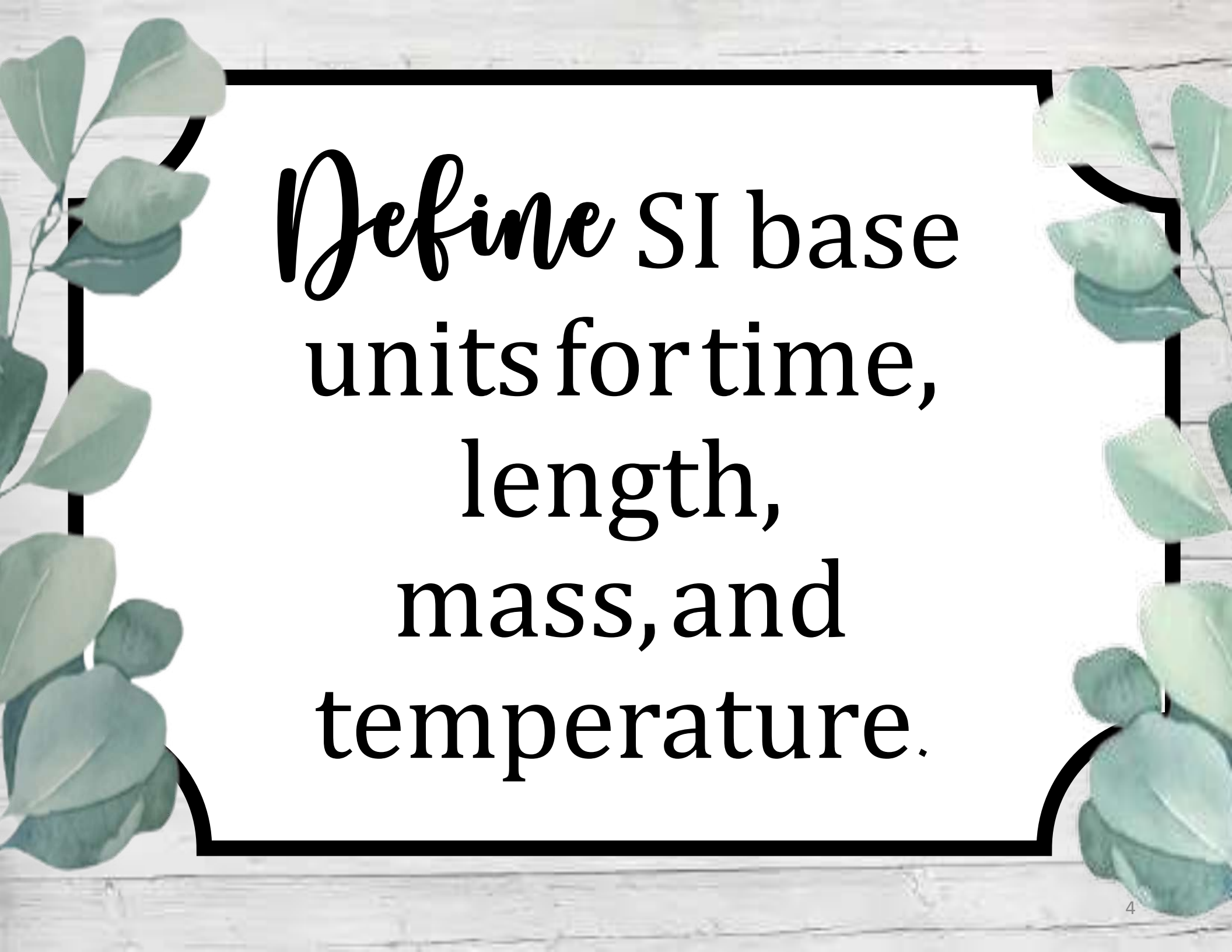


*I can...*

**I can...**



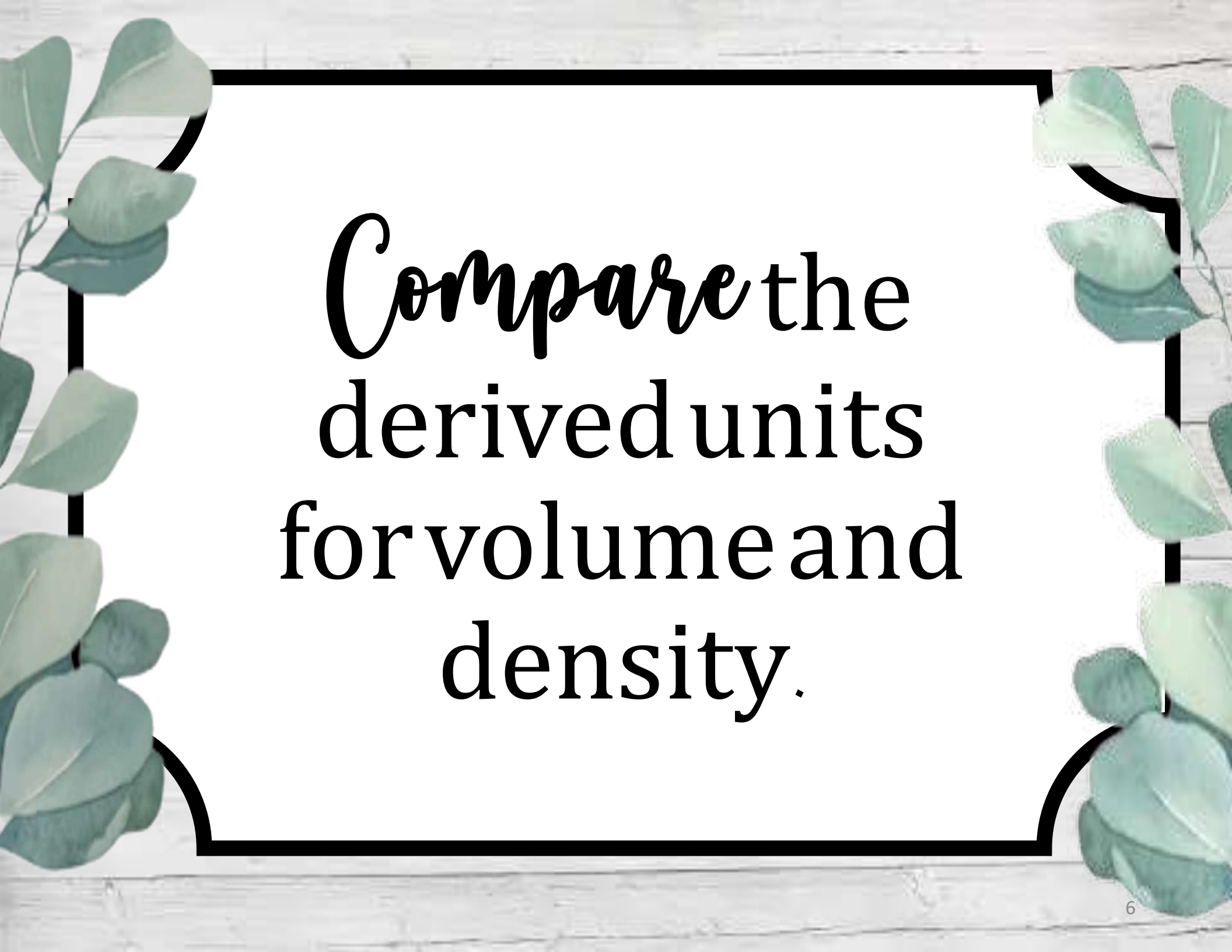
# Measurement & Calculations



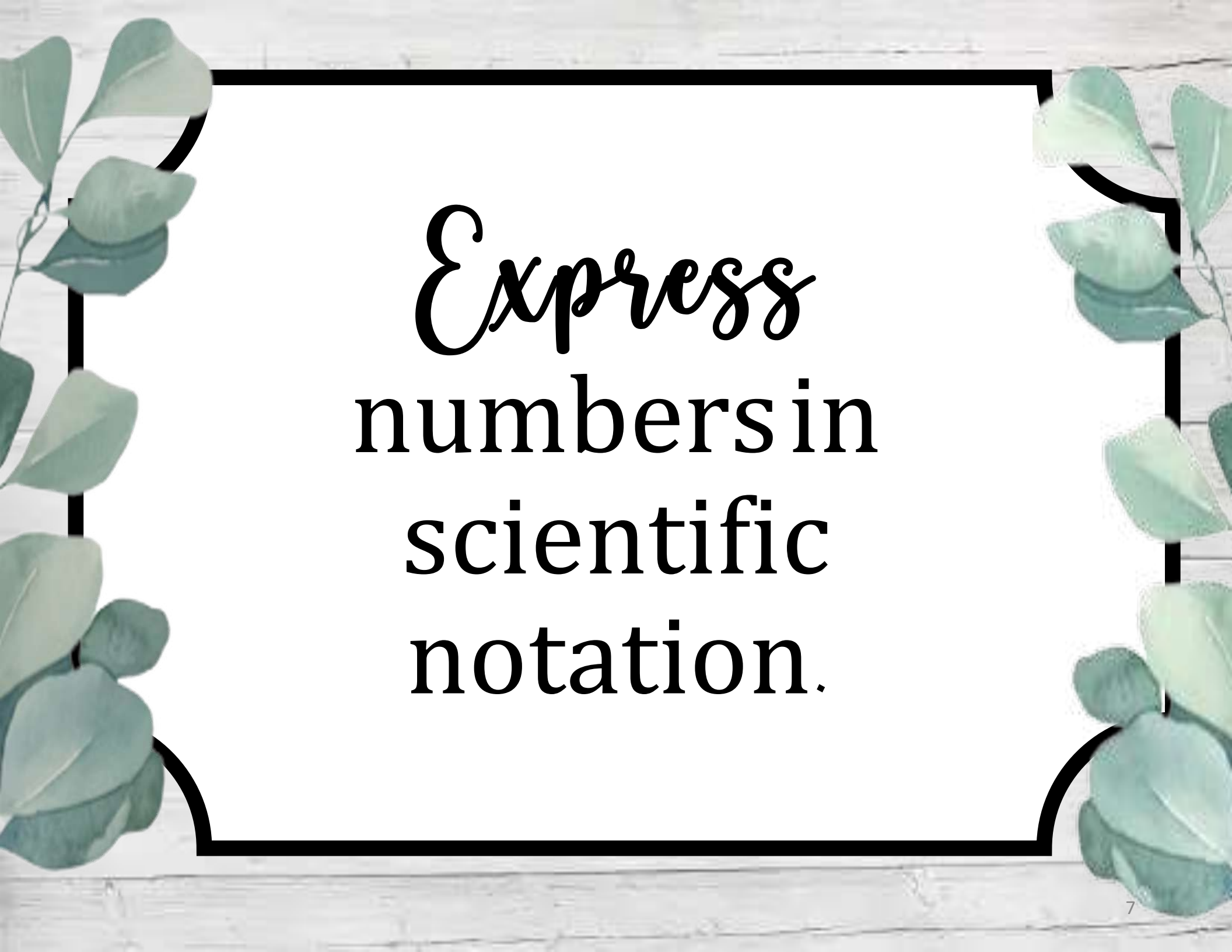
*Define* SI base  
units for time,  
length,  
mass, and  
temperature.



*Explain* how  
adding a prefix  
changes a unit.

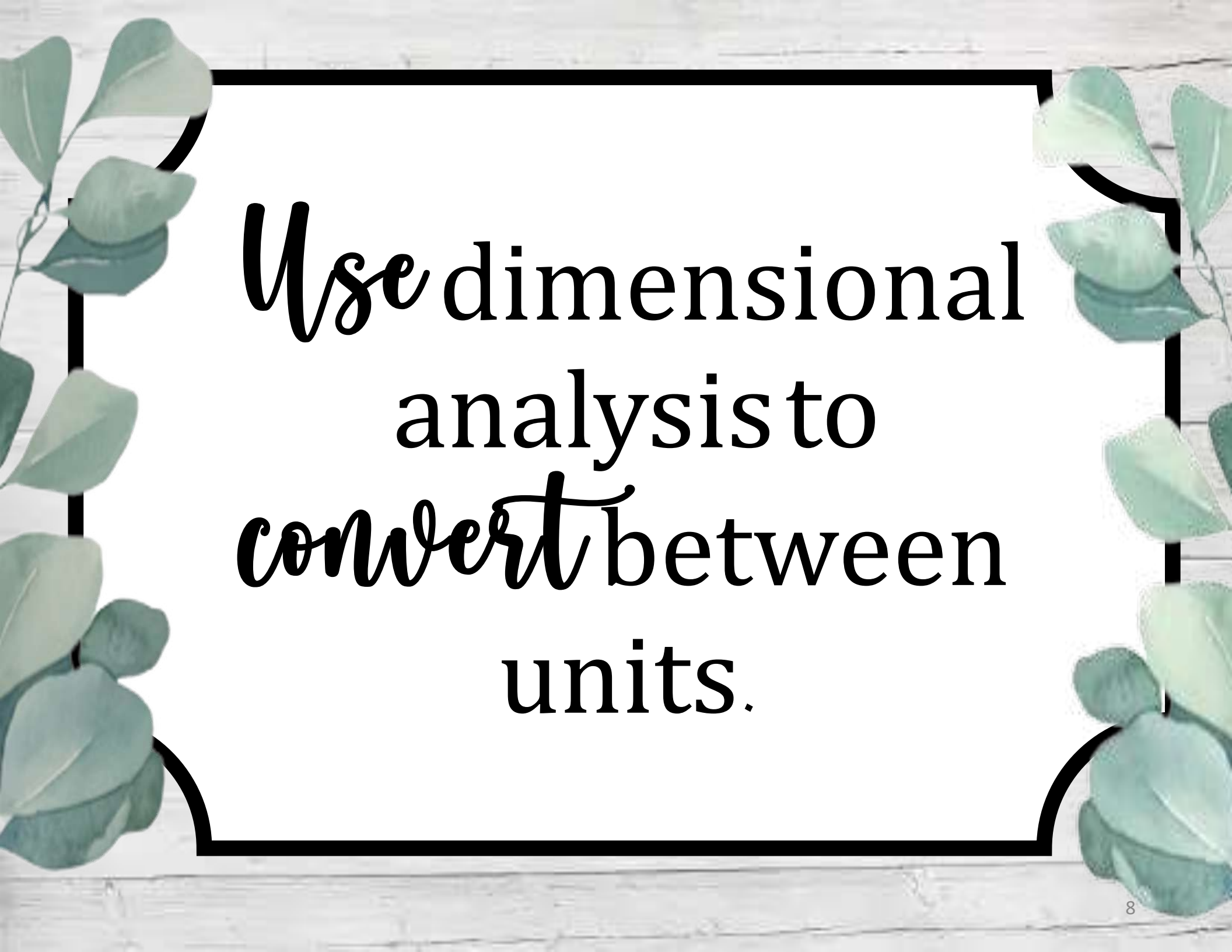


*Compare* the  
derived units  
for volume and  
density.



*Express*  
numbers in  
scientific  
notation.



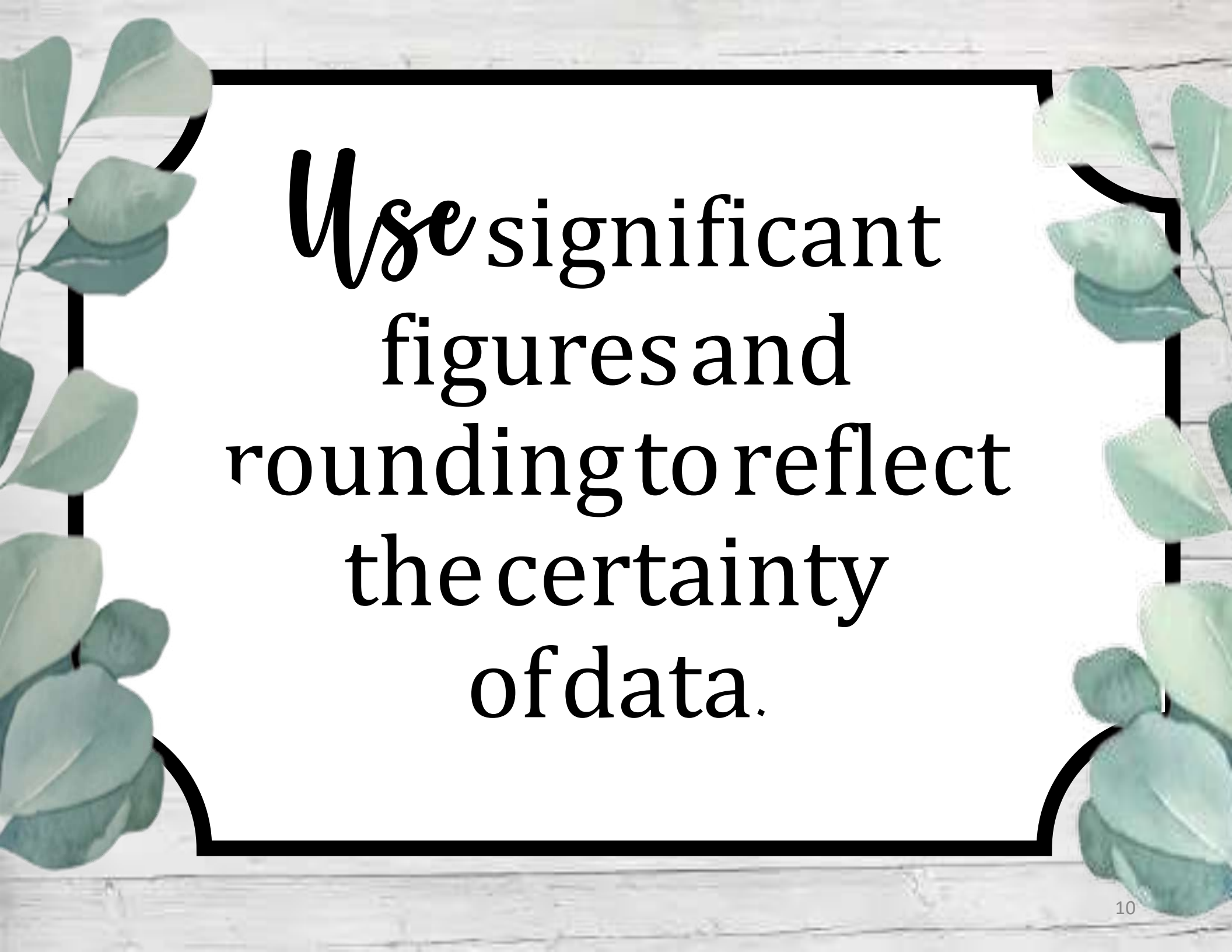


*Use* dimensional  
analysis to  
*convert* between  
units.

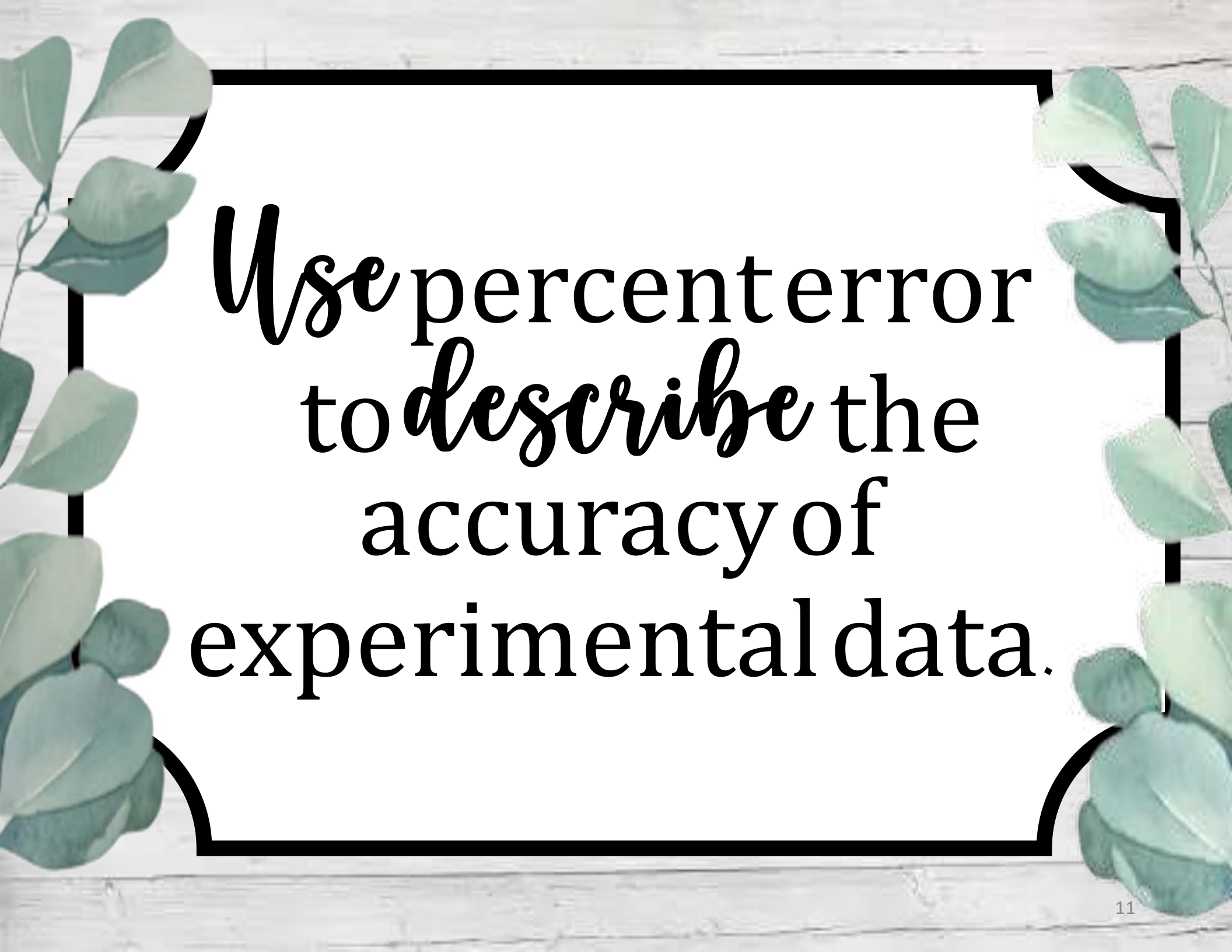




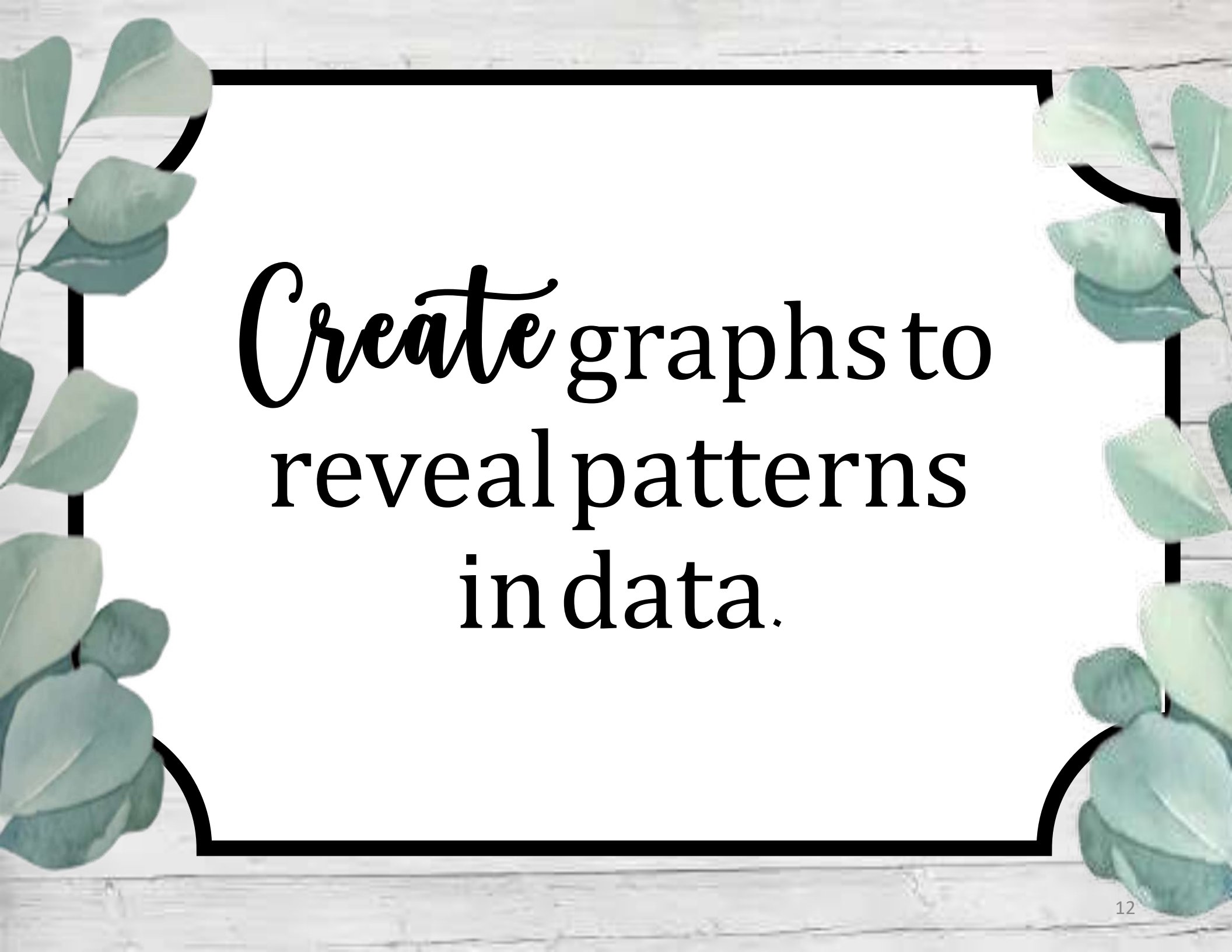
*Define* and  
*compare* accuracy  
and precision.



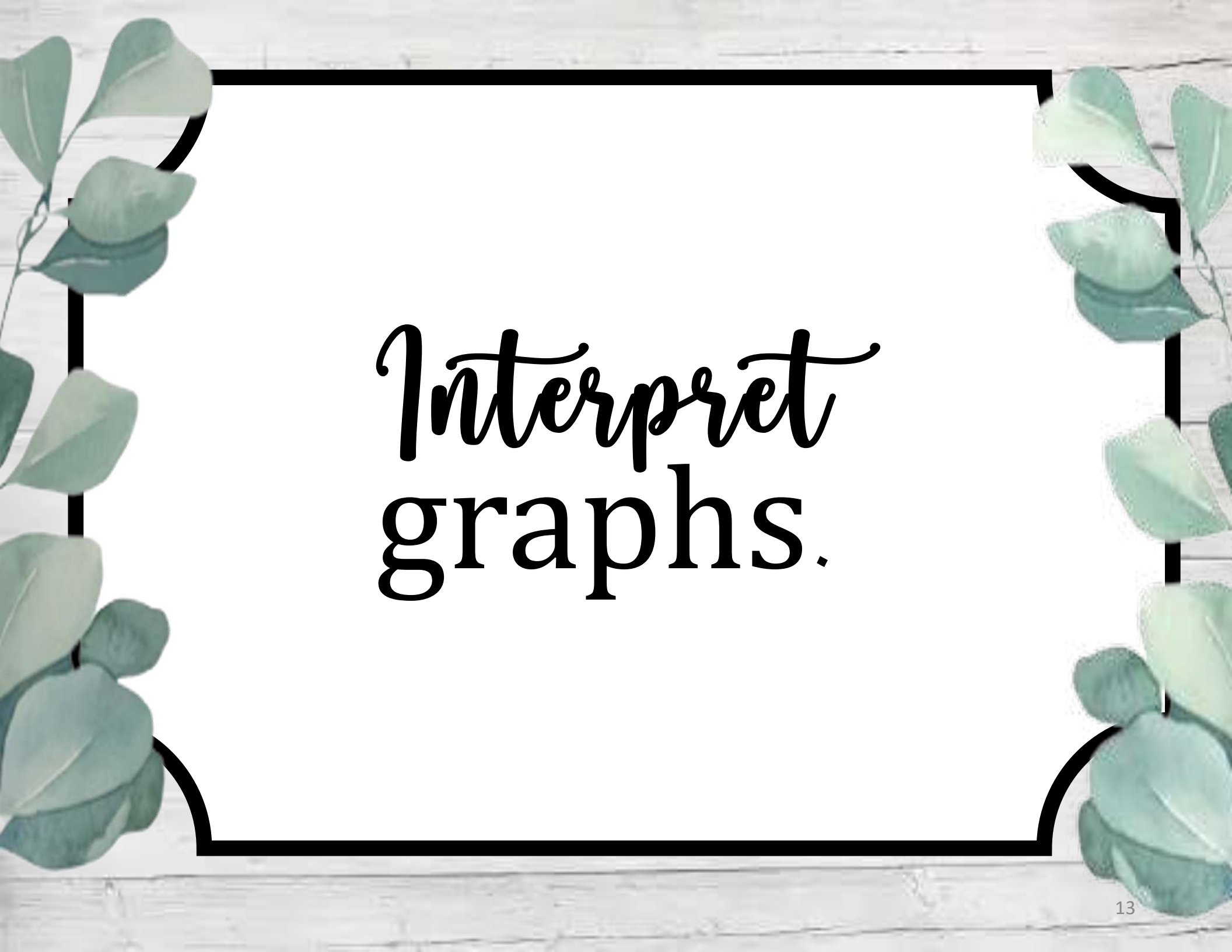
*Use* significant  
figures and  
rounding to reflect  
the certainty  
of data.



Use percent error  
to *describe* the  
accuracy of  
experimental data.

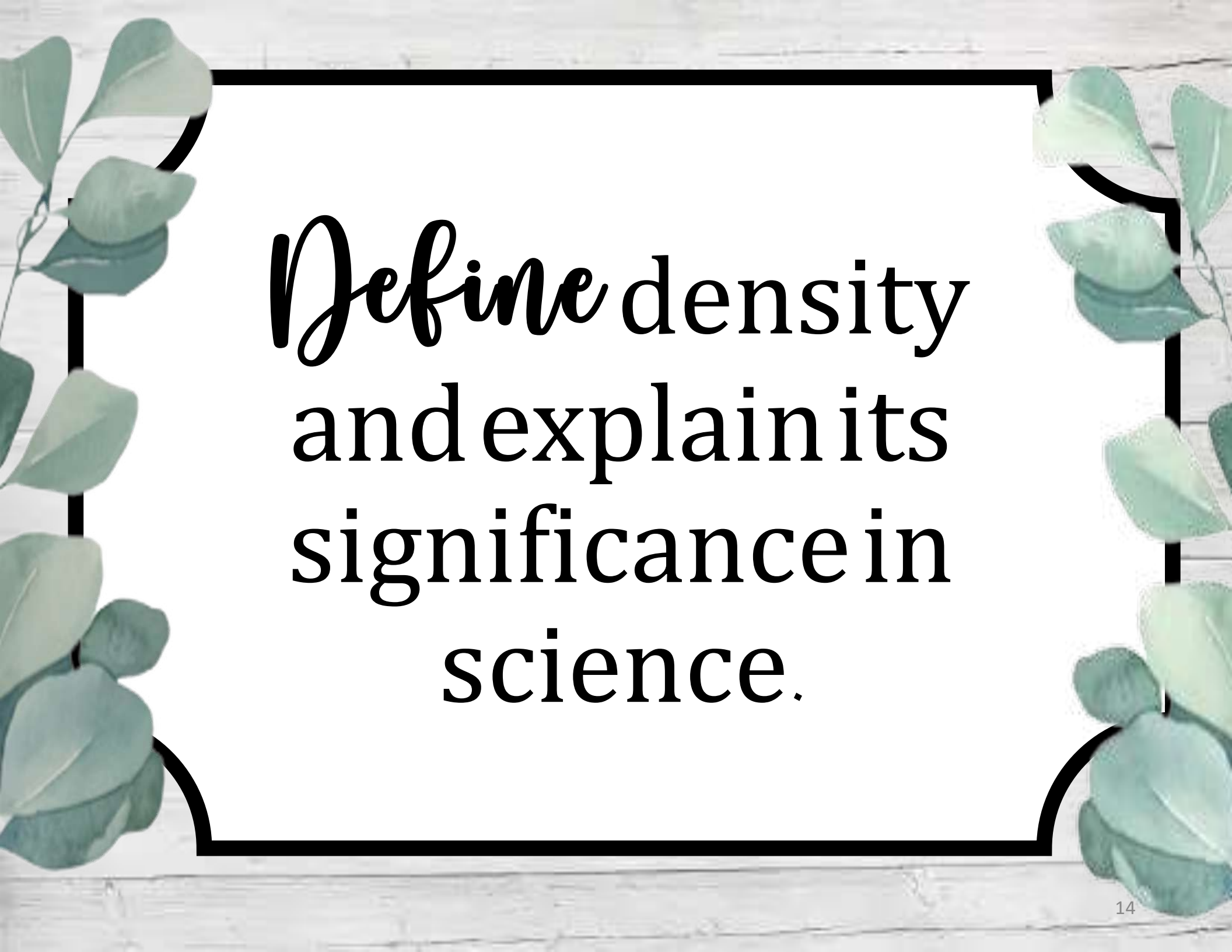
A decorative border featuring green leaves and a black frame surrounds the central text. The leaves are arranged in a symmetrical pattern on the left and right sides, with some leaves overlapping the black frame. The black frame is a simple rectangular border with rounded corners.

*Create* graphs to  
reveal patterns  
in data.


A decorative border featuring green leaves and a black frame surrounds the central text. The leaves are arranged in a symmetrical pattern on the left and right sides, with some leaves overlapping the black frame. The black frame is a simple rectangular border with rounded corners.

*Interpret*  
graphs.



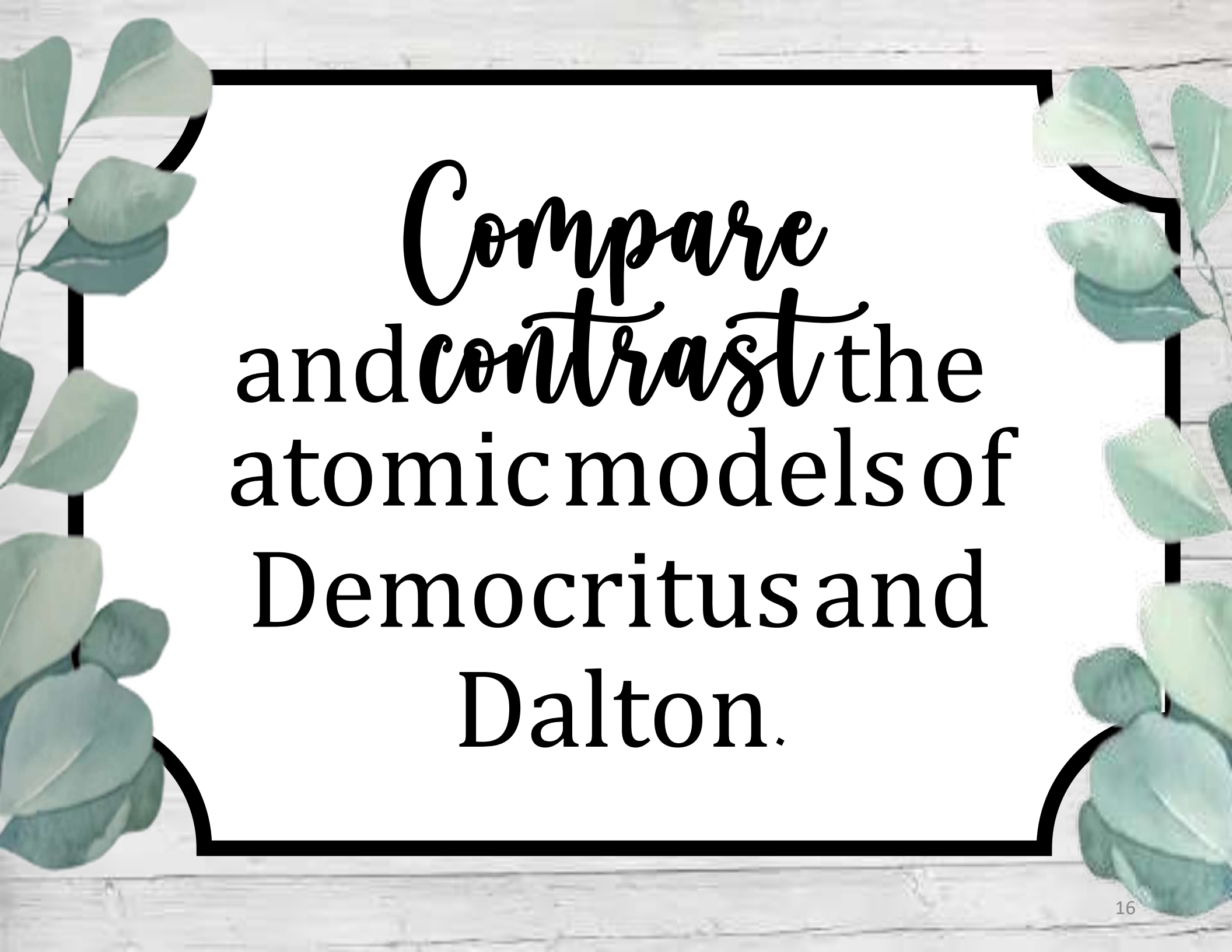


*Define* density  
and explain its  
significance in  
science.



# Chemical Foundations

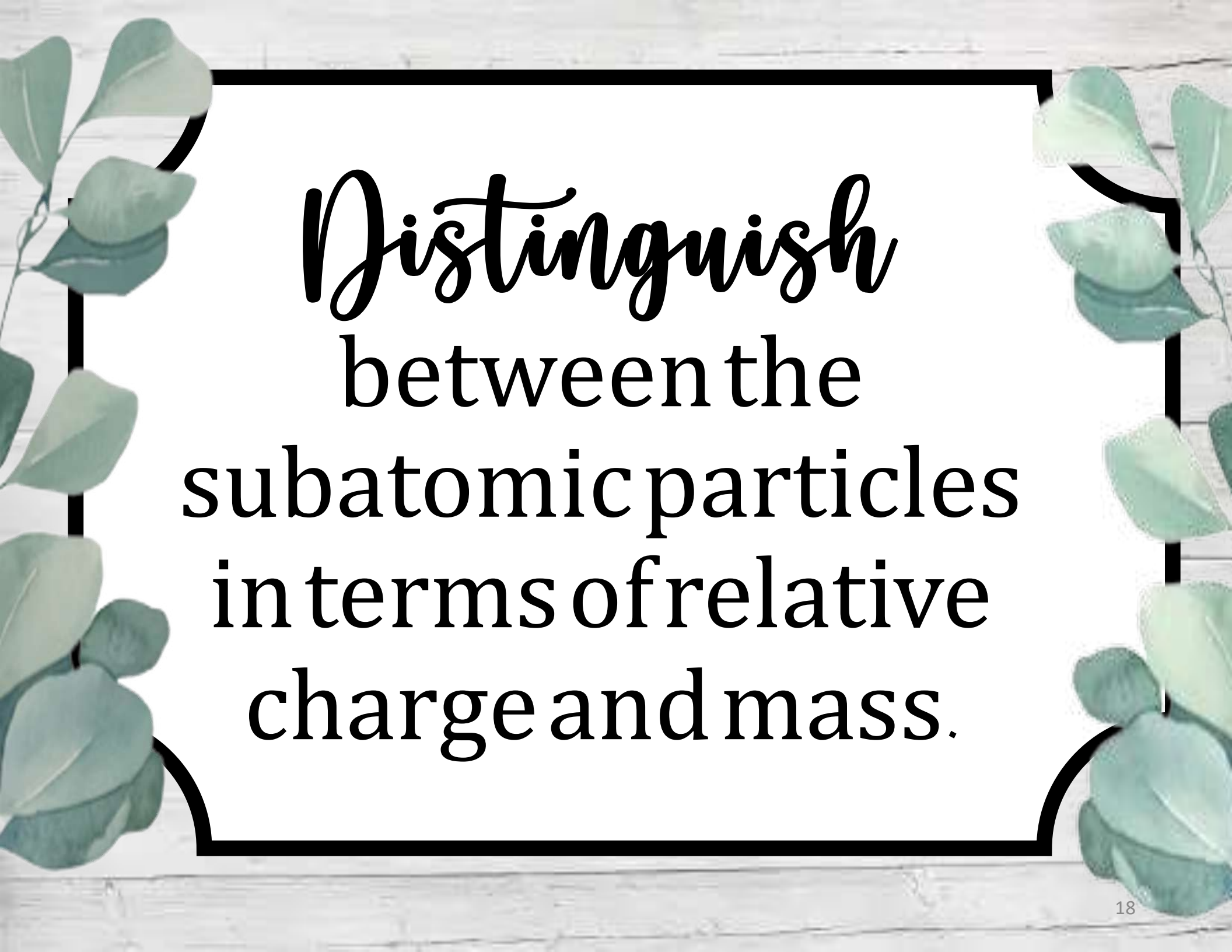




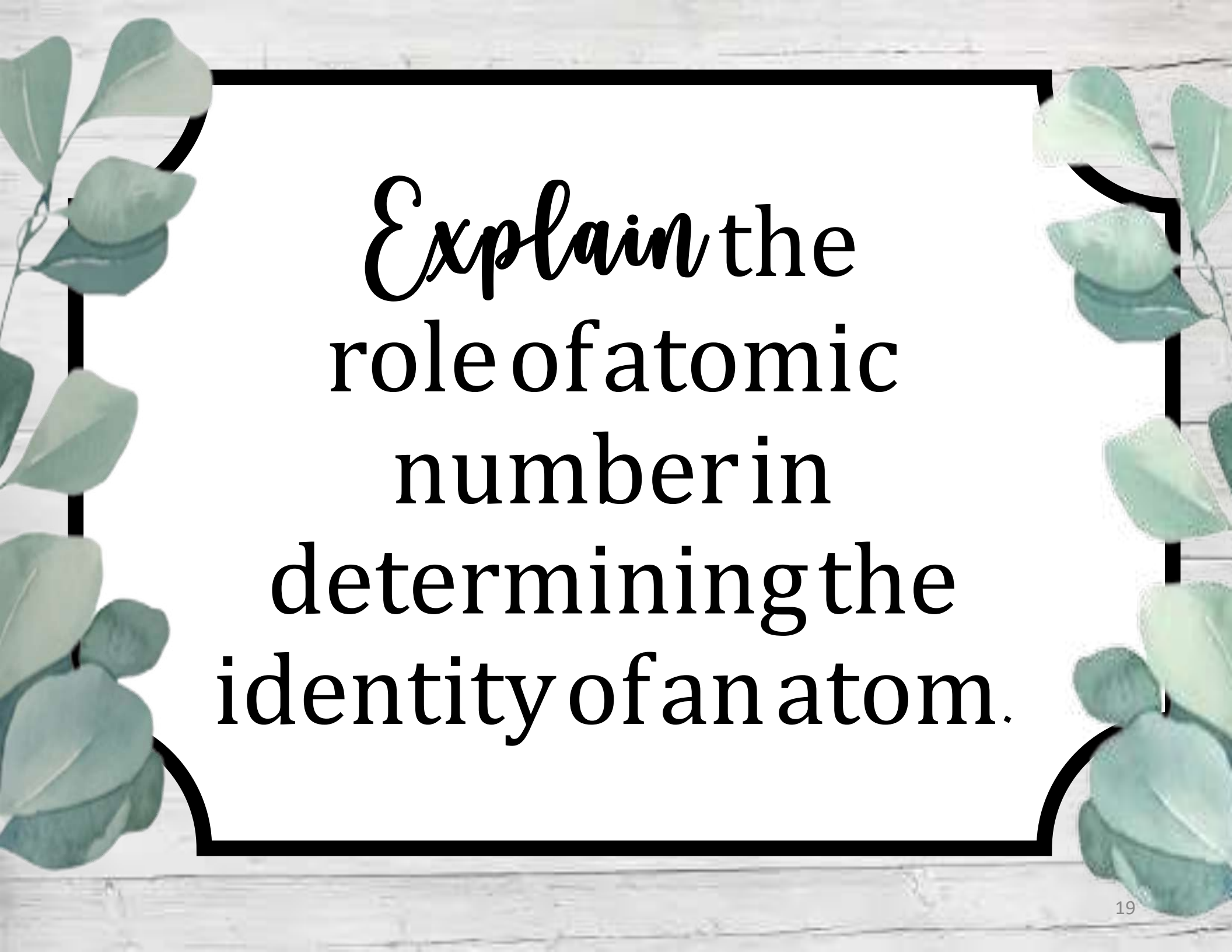
*Compare*  
and *contrast* the  
atomic models of  
Democritus and  
Dalton.

A decorative border featuring green leaves and a black frame surrounds the central text. The leaves are arranged in a corner pattern at the top-left, top-right, and bottom-right, with a black line connecting them to form a frame around the text.

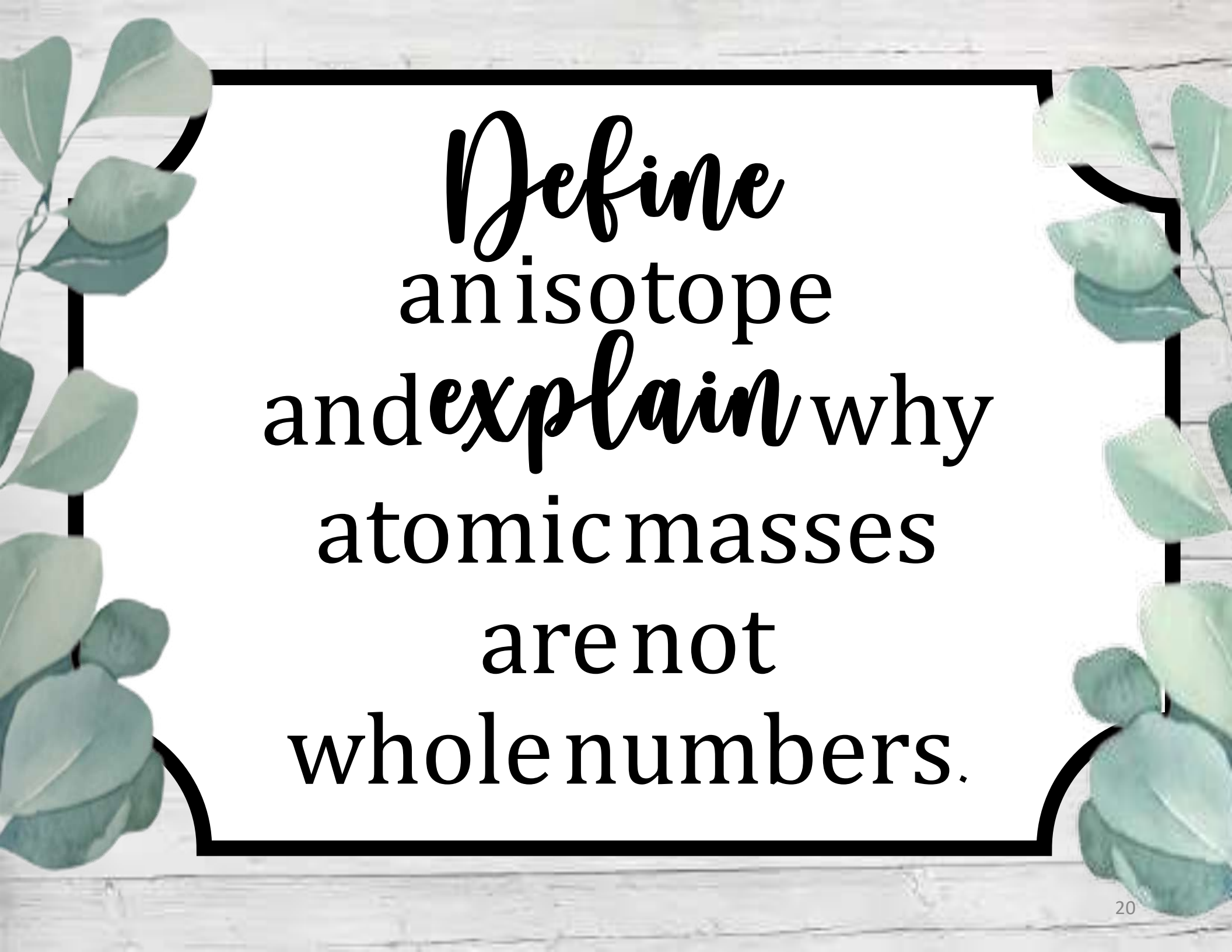
*Define*  
an atom.



*Distinguish*  
between the  
subatomic particles  
in terms of relative  
charge and mass.

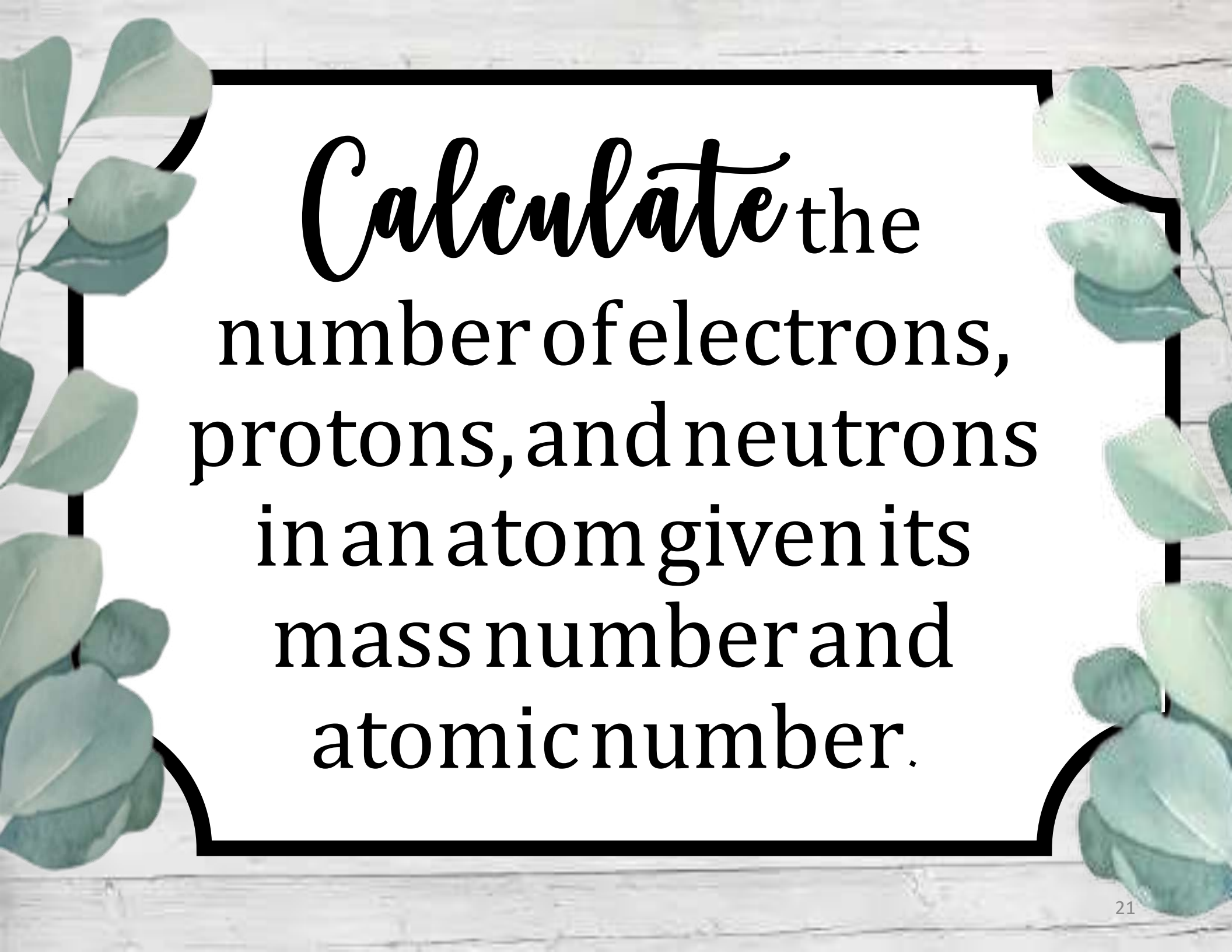


*Explain* the  
role of atomic  
number in  
determining the  
identity of an atom.

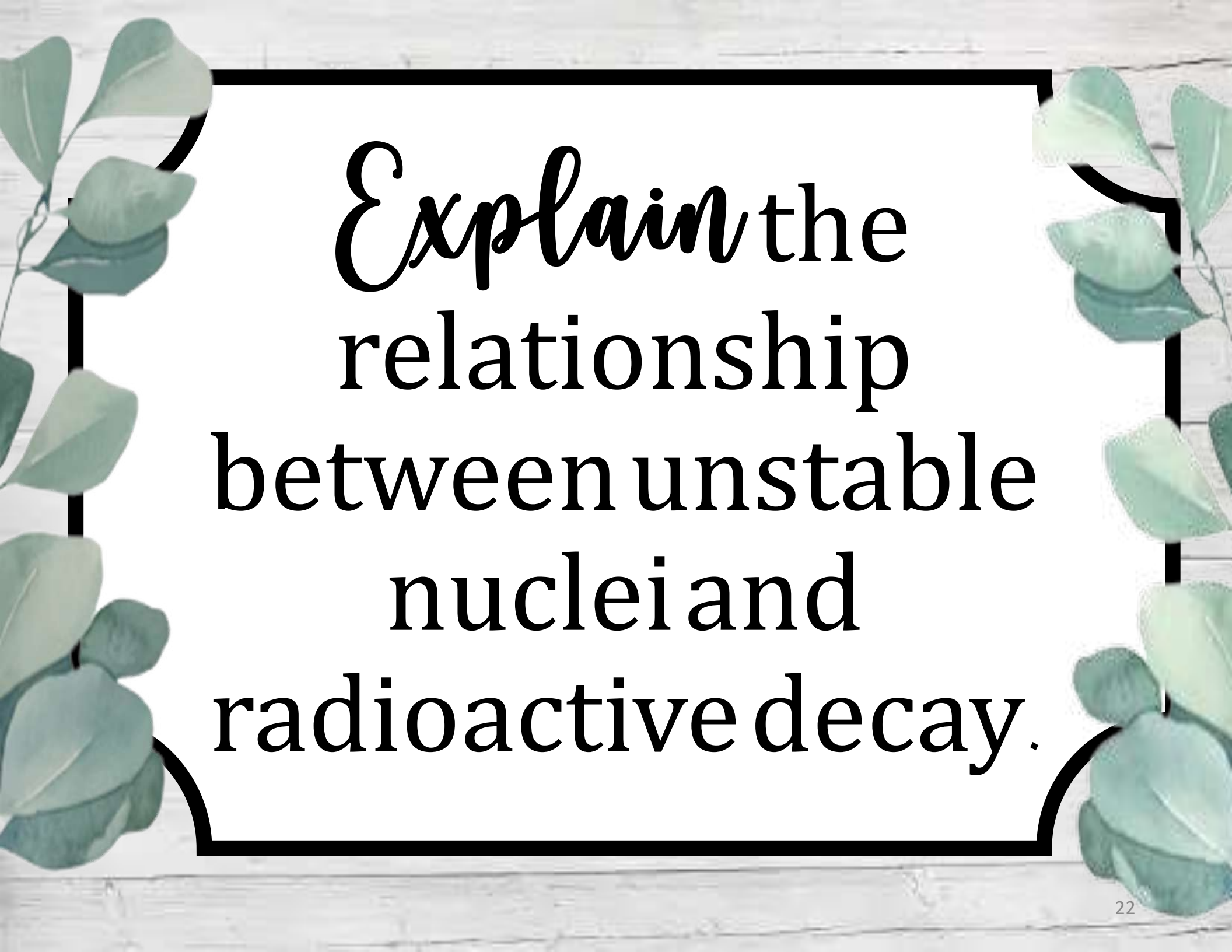


Define  
an isotope  
and *explain* why  
atomic masses  
are not  
whole numbers.



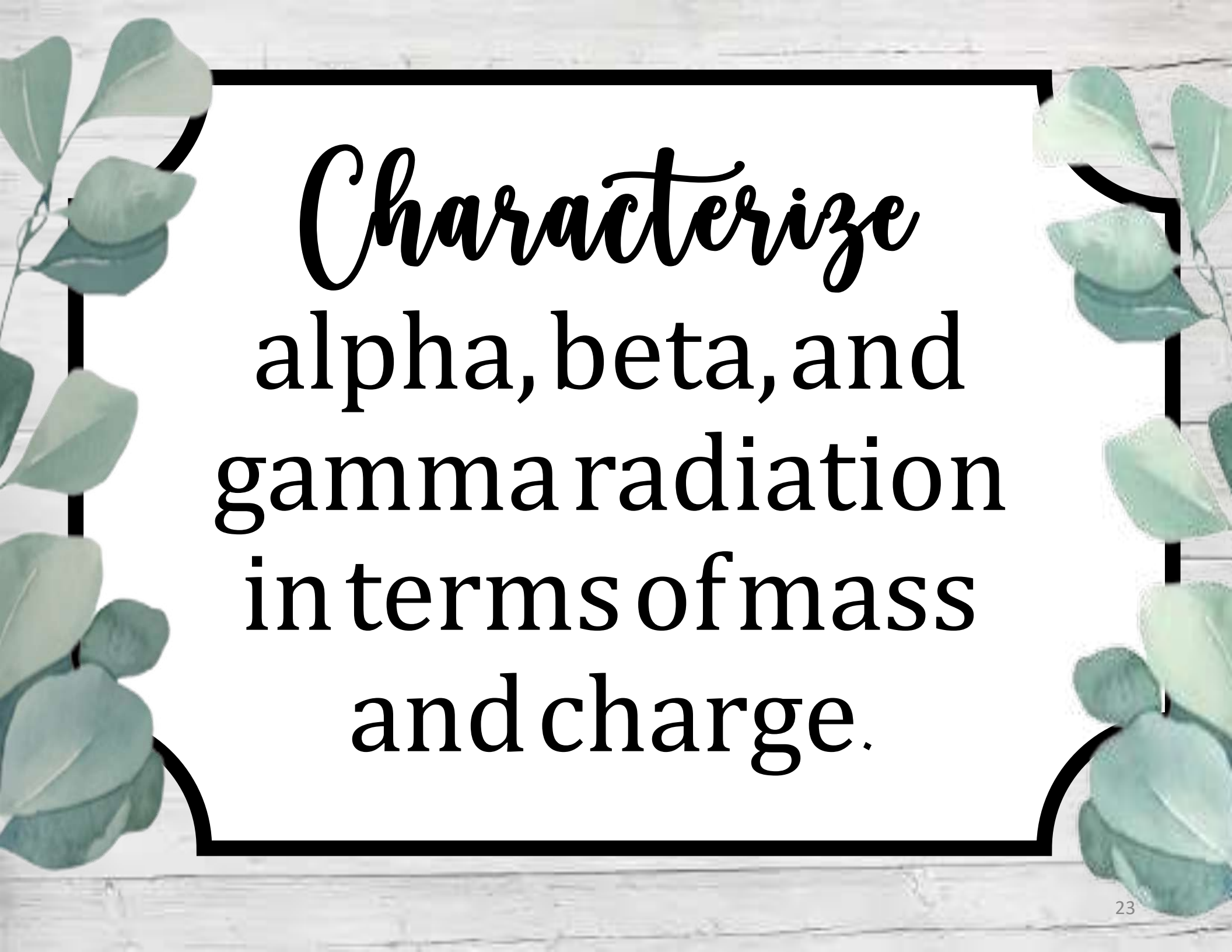


*Calculate* the  
number of electrons,  
protons, and neutrons  
in an atom given its  
mass number and  
atomic number.



*Explain* the  
relationship  
between unstable  
nuclei and  
radioactive decay.





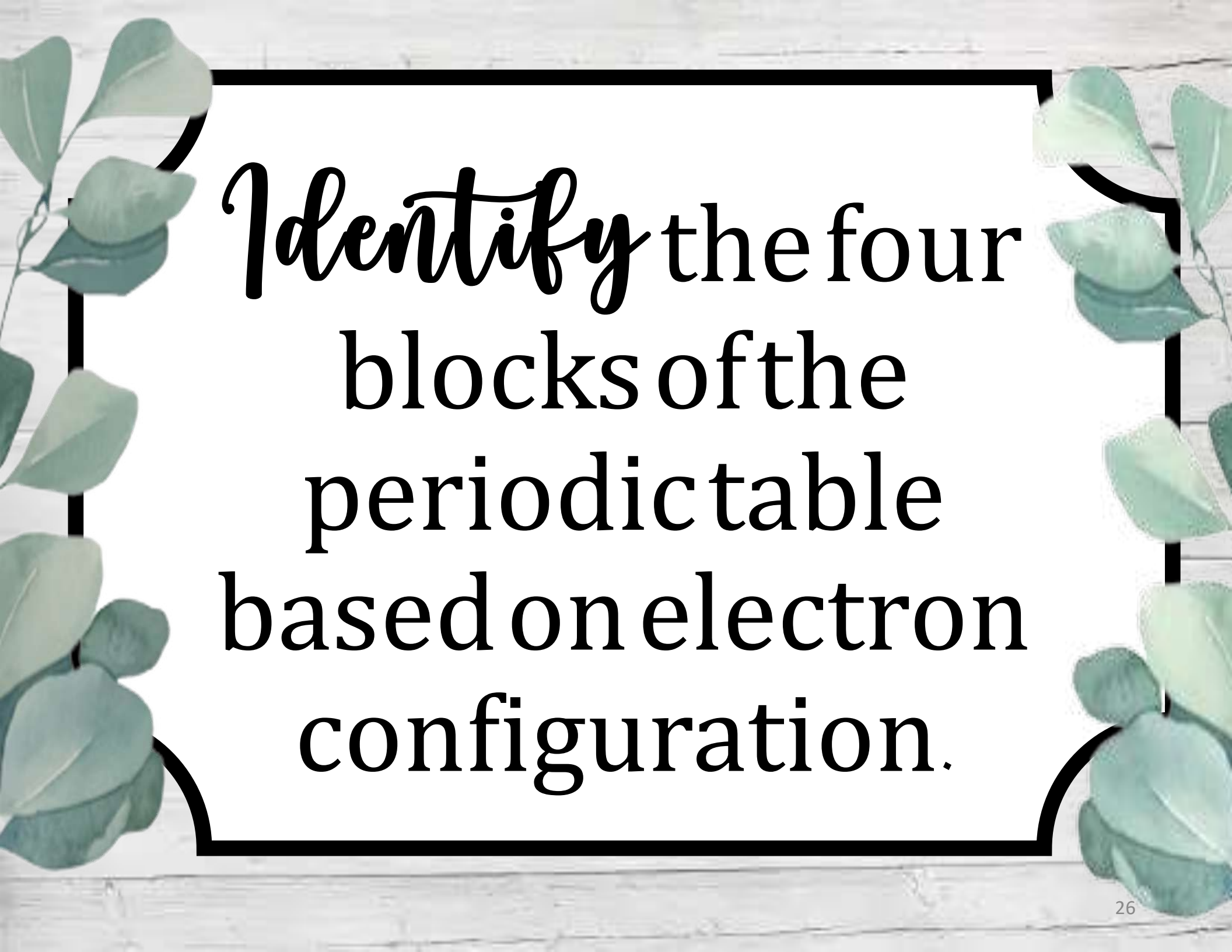
*Characterize*  
alpha, beta, and  
gamma radiation  
in terms of mass  
and charge.



*Trace* the  
development and  
*identify* key  
features of the  
periodic table.



*Explain* why  
elements in the  
same group have  
similar properties.

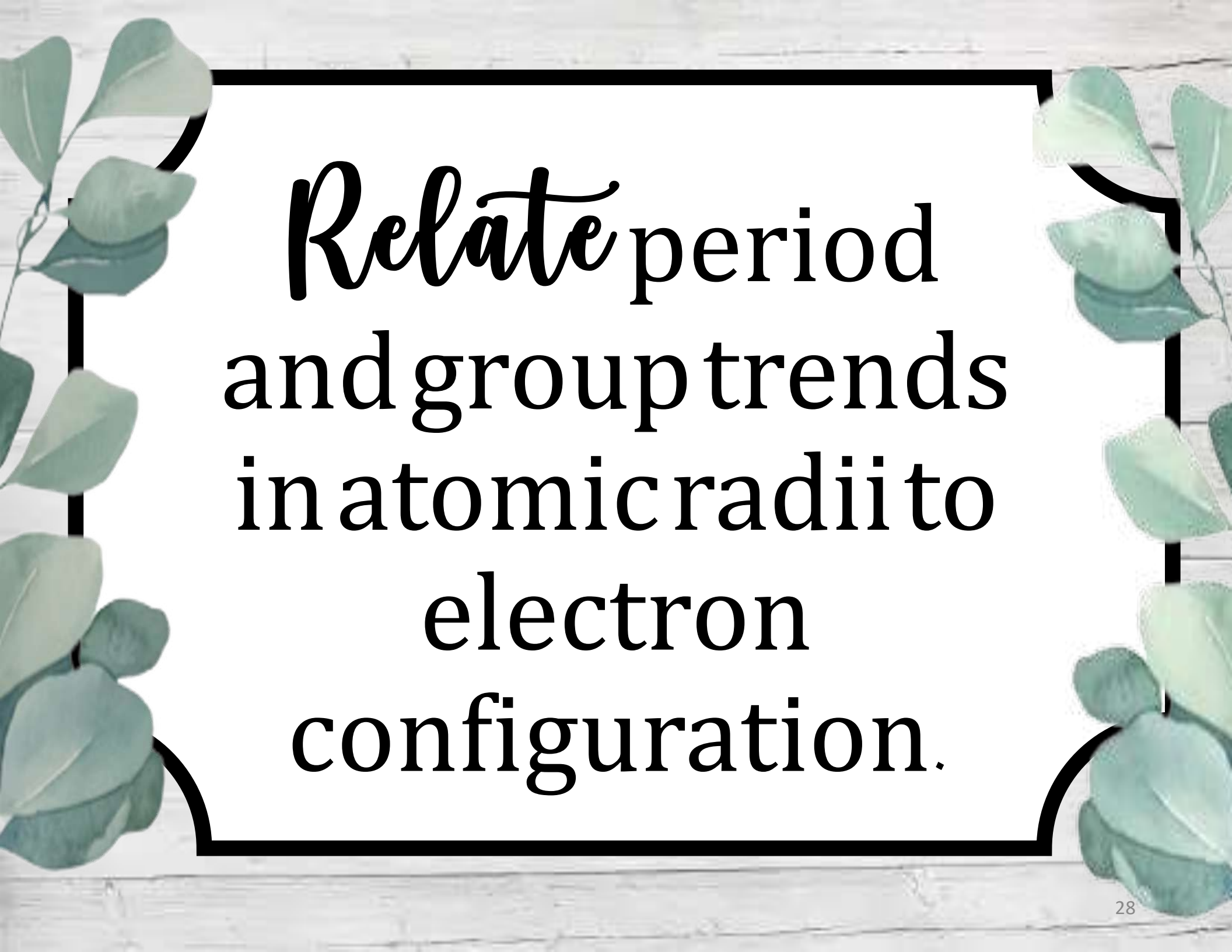


*Identify* the four  
blocks of the  
periodic table  
based on electron  
configuration.



*Compare* period  
and group trends  
of several  
properties.





*Relate* period  
and group trends  
in atomic radii to  
electron  
configuration.