Name:	Date:	Per.

## **READING GUIDE: Chapter 2.3– Carbon Compounds (pages 45-49)**

The Chemistry of Carbon	6	12.0107
How many valence electrons does carbon have?	ľ	C
2. What types of bonds can carbon form?		CARBON
3. Examine Figure 2-12, how many covalent bonds are there between the two carbon ato	oms in acetylen	e?
Macromolecules  4. Look in the margin at Build Vocabulary (page 46). What does monomer mean?		
What does polymer mean?		<del></del>
5. What are the 4 major groups of macromolecules found in living things?		
6. How do living things use carbohydrates? a	HO H-C-OH HO-C-H	CH <sub>2</sub> OH C=O HO-C-H
b	н-¢-он н-с-он	H-Ç-OH H-Ç-OH
7. What is another name for simple sugars?	CH <sub>2</sub> OH Glucose	с́н₂он Fructose
8. Name 3 monosaccharides:		
What is the complex polysaccharide that animals store excess sugar called?		
10. What are the 2 polysaccharides found in plants? &		
11. Identify the 3 common categories of lipids		

For questions 13-15, use the table below.

12. Compare saturated and unsaturated fats.

Effect of Carbon Bonds on Melting Points					
Fatty Acid	Number of Carbons	Number of Double Bonds	Melting Points (°C)		
Stearic Acid	18	0	69.6		
Oleic Acid	18	1	14		
Linoleic Acid	18	2	-5		
Linolenic Acid	18	3	-11		

13. a	Which of the four fatty acids is satura	ted?
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b. Which are unsaturated?

4. How does the melting point change as the number of carbon-carbon double bonds increase?	
5. If room temperature is 25 C, which fatty acid is a solid at room temperature?	-
6. What elements do nucleic acids contain?	
7. What is the monomer of nucleic acids called?	
18. Using Figure 2-16 as a reference, draw a picture of a nucleotide and label the 3 parts that make up the monomer.	
9. What are the functions of nucleic acids?	
20. Identify the two types of nucleic acids.	
21. What are amino acids?	
22. Identify the functions of a protein.	Ž.
23. How do proteins differ from each other?	a a
24. Describe the 4 levels of structure in proteins.  Primary:	Qualemery etructure
Secondary:	
Tertiary:	
Quarternary:	