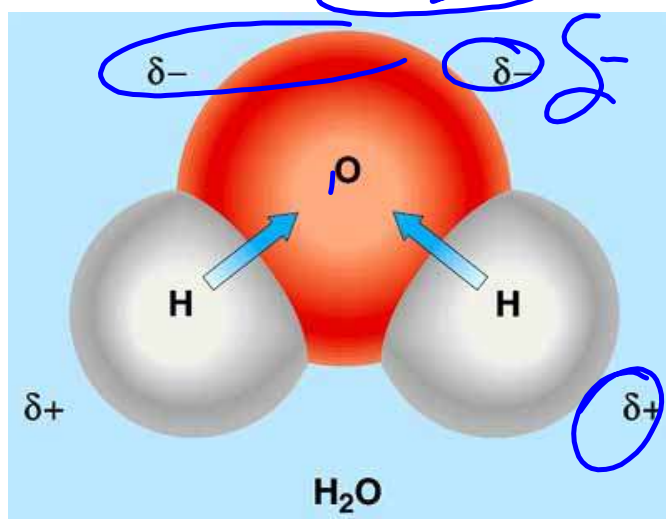
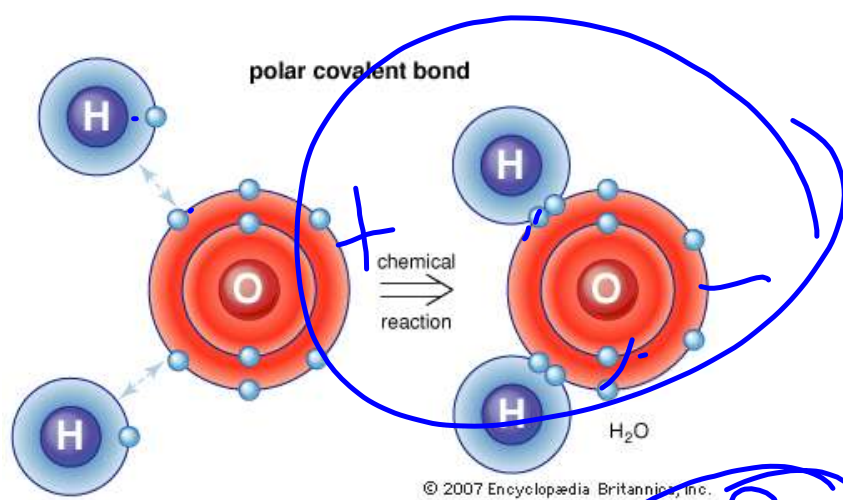


## Naming Molecules:

- Electrons are shared
- Many different combinations of the same 2 atoms are possible: N,O

NO, N<sub>2</sub>O, NO<sub>2</sub>, N<sub>2</sub>O<sub>3</sub>, N<sub>2</sub>O<sub>5</sub>

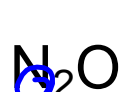
- Can use stock system to name cpds, but since e<sup>-</sup> are shared, no ions formed
- Often, however, e<sup>-</sup> spend more time nearer to one atom making it more "electronegative" than the other atom
- This "apparent charge" = oxidation #

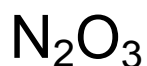


Since oxygen has an oxidation # of  $-2$ , in virtually all its cpds, we can name the N, O cpds:

Stock system

Molec form.	Charge 1 O	Charge all O	Charge all N	Charge 1N	Name
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 $-2$ 
 $-2$ 
 $+2$  nitrosyl oxide

 $-2$ 
 $-2$ 
 $+2$ 
 $+2$ 

 $-2$ 
 $-6$ 
 $+6$ 
 $+3$ 


Can also use Traditional Naming

Use prefixes

$\text{CCl}_4$   $\begin{matrix} +4 & -4 \end{matrix}$  carbon(IV) chloride  
carbon tetrachloride

NO

$\text{N}_2\text{O}_5$