

R. Chem Worksheet 1 (Back To!)

Straight naming - no Roman num.

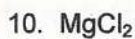
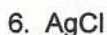
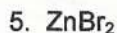
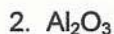
Name: Key

Hour: _____

Date: _____

Chemistry: Ionic Binary Compounds: Single-Charge Cations

Write the name of each of the following compounds.



1. sodium sulfide

2. aluminum oxide

3. sodium chloride

4. rubidium iodide

5. zinc bromide

6. silver chloride

7. boron nitride

8. barium fluoride

9. strontium nitride

10. magnesium chloride

Write the chemical formula for each of the given names.

11. magnesium nitride ^{+2 -3} ← Found on P.T.

12. calcium oxide

13. silver fluoride

14. beryllium chloride

15. potassium iodide

16. aluminum chloride

17. zinc oxide

18. barium bromide

19. lithium nitride

20. potassium sulfide

11. Mg_3N_2

12. CaO

13. AgF

14. BeCl_2

15. KI

16. AlCl_3

17. ZnO

18. BaBr_2

19. Li_3N

20. K_2S

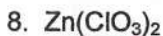
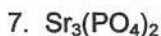
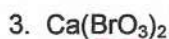
neg ion element
ending to -ide

USE Table E!

Name: _____
Hour: _____ Date: _____

Chemistry: Ionic Compounds: Polyatomic Ions

Write the name of each of the following compounds.



1. Ammonium chloride

2. hydrogen chlorate

3. calcium bromate

4. beryllium sulfate

5. ammonium nitrate

6. ammonium nitride

7. strontium phosphate

8. zinc chlorate

9. silver iodate

10. potassium dichromate

Write the chemical formula for each of the given names.

11. sodium chromate

12. barium nitrate

13. ammonium sulfate

14. aluminum hydroxide *exception*

15. calcium phosphate

16. cesium cyanide *exception*

17. sodium nitrite

18. calcium acetate

19. beryllium chlorite

20. rubidium sulfite

11. Na_2CrO_4

12. $\text{Ba}(\text{NO}_3)_2$

13. $(\text{NH}_4)_2\text{SO}_4$

14. $\text{Al}(\text{OH})_3$

15. $\text{Ca}_3(\text{PO}_4)_2$

16. CsCN

17. NaNO_2

18. $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$

19. $\text{Be}(\text{ClO}_2)_2$

20. Rb_2S