

Preparation for Double Rep. Rxns - Solubility Chart

Sodium hydroxide + cobalt (II) chloride $\rightarrow$ sodium chloride + cobalt (II) hydroxide

barium chloride + sodium phosphate $\rightarrow$ barium phosphate + sodium chloride

Silver nitrate + zinc sulfate $\rightarrow$ silver sulfate + zinc nitrate

lead (II) nitrate + potassium hydroxide $\rightarrow$ lead (II) hydroxide + potassium nitrate

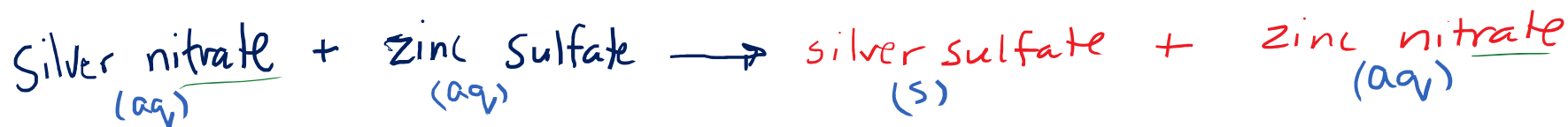
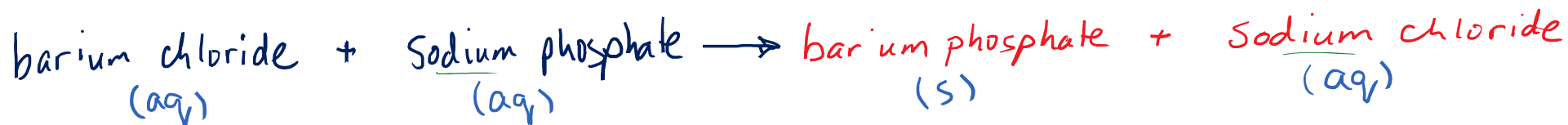
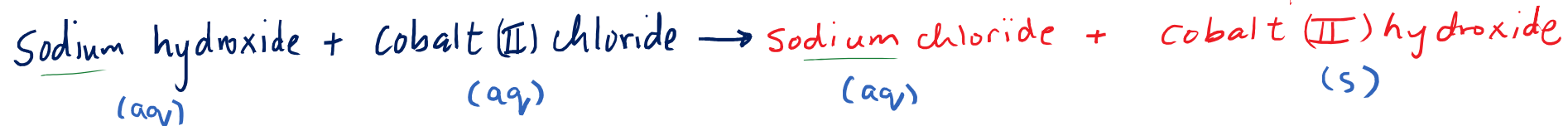
Preparation for Double Replacement Reactions - using the solubility chart

	hydroxide
Sodium	aq

Note: the following are always aqueous

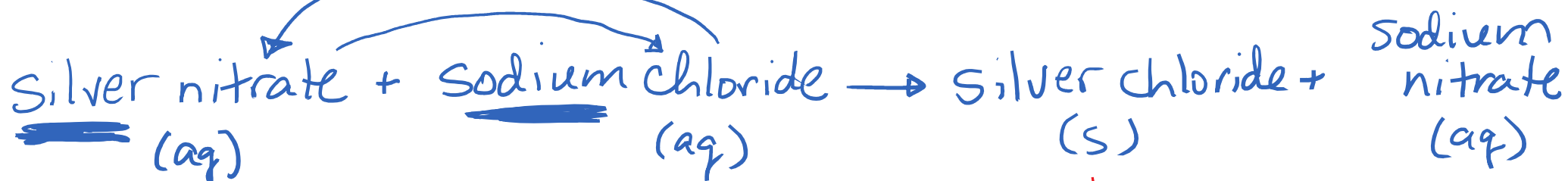
- Sodium
- potassium
- ammonium
- chlorates
- nitrates

	hydroxide
cobalt (II)	s



Double Replacement Reactions

→ ions switch places in two compounds



↓
aqueous -
dissolved in
water

↓
solid precipitate

* Often produce a new solid (precipitate) when two aqueous solutions react.

* Use the solubility chart to identify the solid.

Finish each word equation and use your solubility chart to identify the precipitate.

iron(III) chloride + ammonium hydroxide $\rightarrow$

Copper(II) sulfate + sodium carbonate $\rightarrow$

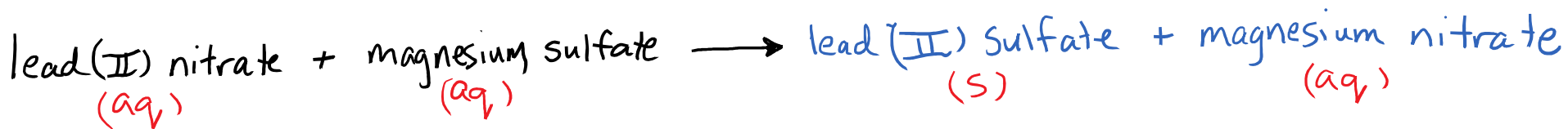
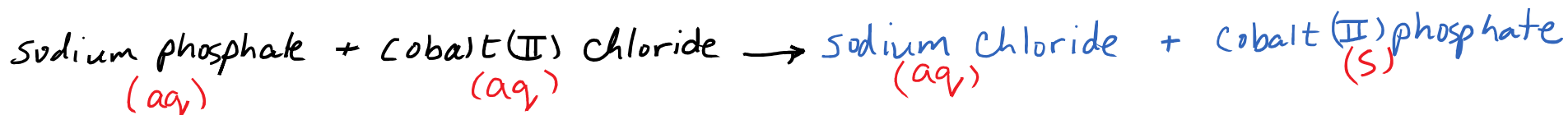
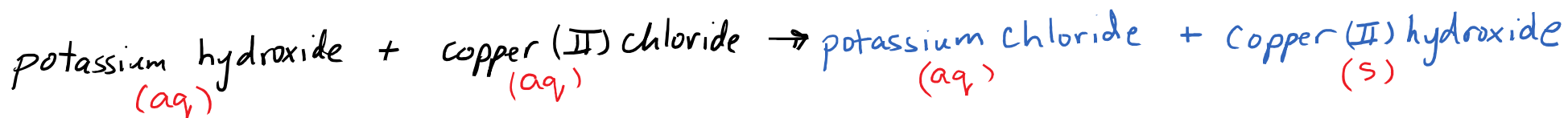
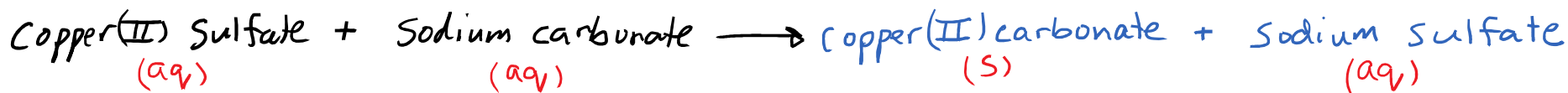
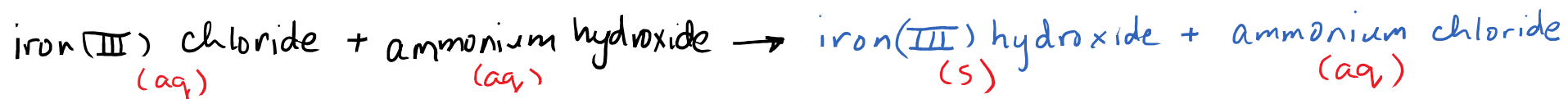
potassium hydroxide + copper(II) chloride $\rightarrow$

sodium phosphate + cobalt(II) chloride $\rightarrow$

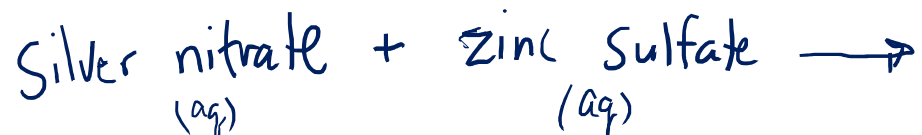
lead(II) nitrate + magnesium sulfate $\rightarrow$

Answers

(s) is the precipitate



You Try



You Try

