Valley School District Chemistry - Organic: Unit 1 - Bond Line Drawings								
Unit #:	ACVSD-00067110	Date(s)	08-31-20	18 to 10-02-2018				
Team: Kristin Grade(s): 11 Subject(s): 5	Team: Kristin Hurrelbrink (Author)       Grade(s): 11, 12       Subject(s): Science							
Unit Focus								
This unit is fo	r learning how to "read" and under	stand bond-line drawing	gs that are use	d in organi	c chemistry.			
Prior Learnings/Connection								
Students should have learned about covalent bonding, valence electrons, charges of ions, and how to draw Lewis Structures and read molecular formulas.								
	Stage 1: Desired Results - Key Understandings							
Standard(s)			Transfer					
<b>Pennsylvania Assessment Anchors and Eligible Content</b> <i>Chemistry: 12</i>		What kinds of long-term, independent accomplishments are desired? Students will be able to independently use their learning to Meaning						
Predict     ion bas	t characteristics of an atom or an	Underst	anding(s)		Essential Question(s)			
table ( potenti <i>CHEM</i> • Recogr of mod the bor compo and-sti solid-s skeleta <i>CHEM</i> • Utilize structu	e.g., number of valence electrons, al types of bonds, reactivity). A.2.2.2 hize and describe different types els that can be used to illustrate hds that hold atoms together in a und (e.g., computer models, ball- ck models, graphical models, bhere models, structural formulas, l formulas, Lewis dot structures). B.1.4.1 Lewis dot structures to predict the re and bonding in simple	What specifically do yo understand? What infer make? Students will und U1 Students will und Organic chr system of o called bond Bond-line d carbon bac because ca bonds Bond-line d	<i>y want studen</i> rences should derstand that emists use a d drawing molect d-line drawings frawings focus kbone of mole arbon only mak	ifferent ules on the cules ses 4 t show	<ul> <li>What thought-provoking questions will foster inquiry, meaning making, and transfer?</li> <li>Students will keep considering</li> <li>Q1 <ul> <li>How can different models/structures be used to represent the same molecule?</li> <li>Why do organic chemists have their own way of drawing molecules?</li> </ul> </li> </ul>			

company and CUEMD112	by draman attached to correct	
Compounds. <i>CHEM.B.1.4.2</i>		
	molecules (but must snow	
	hydrogen atoms attached to	
	atoms other than carbon)	
	<ul> <li>Bond-line drawings are based off</li> </ul>	
	of the Lewis Structures of	
	molecules	
	<ul> <li>The carbon backbones are drawn</li> </ul>	
	using zig-zag lines where the end	
	of each line represents a carbon	
	atom, double bonds have an	
	additional line between carbons	
	and triple bonds have 3 lines	
	between the carbon atoms	
	Triple bonds are drawn in a linear	
	fashion due to the molecular	
	geometry of the atoms	
	Bonds are usually drawn with the	
	atoms as far apart as possible and	
	may be drawn in any direction	
	Terms like "eliminated added or	
	substituted" can be used to	
	describe the changes taking place	
	from one molecule to the other	
	Floctrons can be found in two	
	Electrons can be found in two	
	locations within an atom, either	
	bonds of lone pairs	
	<ul> <li>Neutral atoms are surrounded by the same number of clostness of</li> </ul>	
	the same number of electrons as	
	we would predict for them as	
	valence electrons from the	
	periodic table	
	Atoms with a positive formal	
	charge are surrounded by less	
	electrons than their predicted	
	number of electrons according to	
	the periodic table	
	<ul> <li>Atoms with a negative formal</li> </ul>	
	charge are surrounded by more	
	electrons than their predicted	
	number of electrons according to	
	the periodic table	
	1	

Acquisition of Kn	Acquisition of Knowledge and Skill	
Knowledge	Skill(s)	
<ul> <li>What facts and basic concepts should students know and be able to recall? Students will know</li> <li>K1 During the unit the students will be able to define and use the following scientific terms: <ul> <li>Bond-Line Drawing</li> <li>Molecular Formula</li> <li>Condensed (Structural) Formula</li> <li>Lewis Dot Structure</li> <li>Orbitals</li> <li>Valence electrons</li> <li>Lone pairs</li> <li>Formal Charge</li> <li>Negative Charge</li> <li>Molecular Geometry</li> <li>Reaction terms: <ul> <li>added</li> <li>eliminated</li> <li>substituted</li> </ul> </li> </ul></li></ul>	<ul> <li>What discrete skills and processes should students be able to use? Students will be skilled at</li> <li>S1 Students will be able to</li> <li>Define what a bond-line structure is.</li> <li>Read bond-line drawings by writing the molecular formula for them.</li> <li>Draw bond-line drawings when given a Lewis Structure and/or a condensed structural formula.</li> <li>Analyze a bond-line structure for the number of carbon and hydrogen atoms it contains.</li> <li>Explain the changes taking place in a reaction when given the bond- line drawings of the reactants and products.</li> <li>Explain how/why a formal charge is created.</li> <li>Assign formal charges to atoms in molecules that are surrounded by a different number of electrons than would be predicted from the periodic table.</li> <li>Determine how many (if any) lone pairs are present on various atoms, including atoms with charges.</li> </ul>	
Stage 2: Assessment Evidence		

Alignment	Code	Assessment Evidence					
	PT1	Untitled Performance Task Formative: • student note packet highlighting and practice problems • worksheet: organic chemistry w/ Klein #1 • Worksheet: drawing formal charges #2 • Practice with building organic molecules from kits and bond-line drawings • Lab - polyurethane hands Summative: • Quiz on sections 1-4 • Chapter 1 test • Lab questions					
	Stage 3: Learning Plan						
Alignment	Code	Learning Activities					
	LA1	Untitled Learning Activity Day 1: Complete notes and practice for section 1.1 (how to read B-L drawings) Day 2: Complete notes and practice for section 1.2 (how to draw B-L drawings) Day 3: Molecule building - students will build the various complex molecules with the kits. They will have them checked as they complete them. Day 4: Complete notes and practice for section 1.3 (Mistakes to avoid) *additional class practice on the board with sections 1.1-1.3 Day 5: Complete notes and practice for section 1.4 (More exercises) Students identify what has changed from one side of an equation to the other. Day 6: Molecule building - students will again build the various complex molecules with the kits. They will have them checked as they complete them. Day 7: Review and take quiz on sections 1.1- 1.4. Day 8: Go over quiz and begin section 1.5 (Identifying formal charges)					

Day 9: Continue with identifying formal charges.         Day 10: Complete section 1.6 (finding lone pairs that are not drawn)         Day 11: Students will complete the second set of practice problems (worksheet) that covers sections 1.5 and 1.6.         Day 12: Videos discussing formal charges/lone pairs/bond line drawings from Crash Course Chemistry and/or Professor Dave Explains. Time permitting: additional group practice problems depending on what students n practice with.         Day 13: Polyurethane Lab experiment and concept questions.         Day 14: Review for all types of problems on the test. (go through and complete group practice problems of e type)         Day 15: Chapter 1 Test.					
Additional Information					
Resources					
RES	RES	RES			
RES1	Organic Chemistry I: As a Second Language By: David R. Klein - This is the textbook we use for notes and many practice problems.				
RES2	Video - Professor Dave Explains - Lewis Structures	https://www.youtube.com/watch? v=Sk7W2VgbhOg			