Chemical Reactions

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B. Food being digested
C. A match burning
D. Ice melting

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B. four
C. five
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- B. Increases the energy given off by a reaction
- C. Lowers the activation energy of a reaction
- D. Changes permanently when used in a reaction

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Match the labels to the equation.

A.
B.
C.
D.
E.

- yield signs
- subscript
- product
- coefficient
- reactants



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- B. subscript
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- D. product
- E. reactants

Match the correct description with the correct term.

- 1. $6CO_2 + 6H_2O + energy \longrightarrow C_6H_{12}O_6 + 6O_2$ • 2. $H_2CO_3 \longrightarrow H_2O + CO_2$ • 3. $2Na + CI_2 \longrightarrow 2NaCI + energy$
- 4. $Zn + 2HCI \longrightarrow ZnCl_2 + H_2$
- A. Exothermic reaction
- B. Decomposition
- C. Endothermic reaction
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A. decreasing the concentration of the reactants

B. grinding a solid reactant into a powder
C. lowering the temperature of the reactants

• D. raising the temperature of the products

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Why does the law of conservation of mass dictate that chemical equations be balanced?

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- C. because chemical formulas are written from left to right
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In which kind of chemical reaction do two or more substances combine to form one new compound?

A. synthesis reaction

- B. decomposition reaction
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Which of the following examples shows where a coefficient is used correctly to balance the equation?

• A. Na + $Cl_2 \longrightarrow 2NaCl$ • B. Na + $2Cl_2 \longrightarrow 2NaCl$ • C. $2Na + Cl_2 \longrightarrow 2NaCl_2$ • D. $2Na + Cl_2 \longrightarrow 2NaCl_2$ Which of the following examples shows where a coefficient is used correctly to balance the equation?

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