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## Objectives

After reading this chapter, you will understand:

- How to apply deductive reasoning to a series of analytical data.
- The limitations of presumptive (screening) tests.
- The relationship between the electromagnetic spectrum and spectroscopic analysis.
- The difference between qualitative and quantitative analysis.
- The dangers of using prescription drugs, controlled substances, over-the-counter medications, and alcohol.

You will be able to:

- Chemically identify illicit drug types.
- Classify the types of illicit drugs and their negative effects.
- Discuss the federal penalties for possession and use of controlled substances.
- Explain the need for confirmatory tests.
- Describe IR, UV-VIS spectroscopy, and GC-MS, and explain how they are used in forensic science.
- Present and interpret data with graphs.
- Use the *Physicians' Desk Reference (PDR)* to identify pills.



**“Every form of addiction is bad, no matter whether the narcotic be alcohol or morphine or idealism.”**

*—Carl Jung (1875–1961), Swiss psychiatrist*



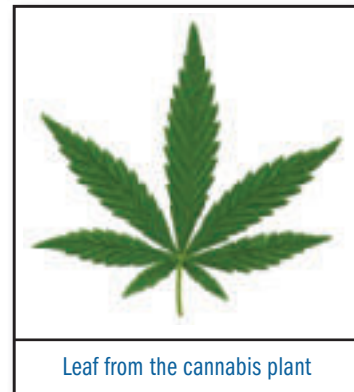


## Drugs and Crime

A drug is a natural or synthetic substance designed to affect humans (or other animals) psychologically or physiologically. A drug can affect the function or structure of living tissue through various chemical reactions. When any type of drug is taken in excessive amounts and causes illness or death, it shows toxic effects and is classified as a poison. Most drugs are legitimately manufactured by drug companies and are prescribed for particular medical problems.

Sometimes, however, they can have effects that people find pleasurable. When these substances, produced legally or illegally, are taken strictly for pleasure, they are considered “drugs of abuse” or “illicit drugs.”

In the United States, all drugs covered by law that are somehow restricted are called “controlled substances.” They are listed in a part of the Federal Code called the Controlled Substances Act (Act 21 U.S.C. 812).

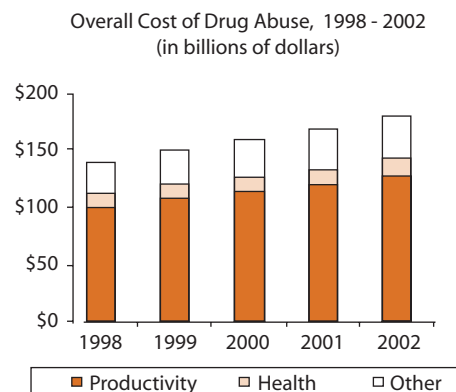


The Drug Enforcement Agency (DEA) is the lead agency for enforcement of federal drug laws.



Recreational use of drugs may cause dependency, where one's energy and time are spent getting the drugs needed to feel good. This can lead to theft to support such a habit. It also can lead to violence committed during robberies as well as for control of drug distribution and profits. Drug abuse can lead to extreme mental and physical health problems, creating a burden on the public health system.

The Office of National Drug Control Policy estimated that illegal drug use cost American society \$181 billion in 2002, and it appears the trend increases by about 5 percent a year. About 30 percent of prisoners committed their current offense while under the influence of drugs. In 2002, 4.7 percent of the 14,263 homicides in the United States were drug-related. One-quarter to one-half of all incidents of domestic violence are drug-related. Scary, isn't it?



And, of course, there is a social, moral, and religious stigma associated with drug abuse, as well as loss of productivity and pride among users.

In the United States, as much as 75 percent of the evidence being examined in forensic laboratories is considered drug-related, either the drugs themselves or evidence from drug-related crimes. Forensic chemists have many tests for identifying drugs and poisons. As soon as a pharmaceutical company produces a new drug, it sends a sample to the FBI Crime Lab so they can develop procedures to identify both large and minute quantities. These procedures are then filed for reference when unknown samples are collected at a crime scene.

There are five categories of **controlled drugs**, as defined by the Controlled Substances Act; see Appendix A at the end of the book. The legal penalties are listed in Appendix B for marijuana and in Appendix C for other controlled substances.

The following list is a widely used classification scheme for illicit drugs, based on their pharmacological effects:

- **Hallucinogens** are mostly naturally occurring substances that can change normal thought processes, perceptions, and moods. The most widely used hallucinogen in the United States is marijuana, a plant that has been used as medicine, to make rope (hemp), and to produce euphoria for thousands of years. The active ingredient in marijuana can range from 1 percent in low-grade material to 10 percent in sinsemilla, a cultivated female **cannabis**. A resinous oil (“hash oil”) extracted from the plant can contain as much as 50 percent of the active ingredient; more often, this oil is only partially extracted from the ground leaves and sold as hashish, or “hash.” Marijuana decreases the user’s ability to concentrate, slows reactions, and impairs coordination.  
LSD (*d*-lysergic acid diethylamide, or “acid”) is chemically derived from ergot alkaloids found in a fungus that grows on grain. As little as 25 µg can cause visual and auditory hallucinations. LSD is often sold soaked into blotter paper (“blotter acid”), as tiny

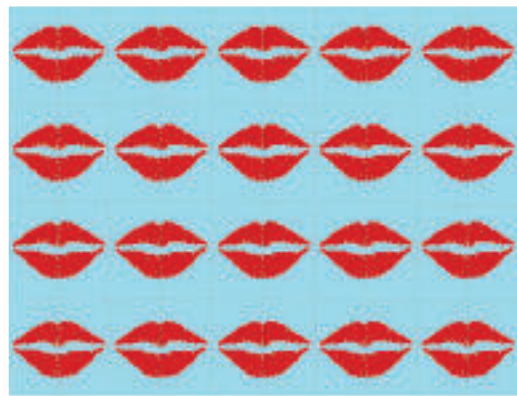
### Salem Witch Trials and LSD

Ergot is a fungus that infects the grain of rye plants in cold, wet weather. Ingested in large doses, it restricts blood flow, causing gangrene in the extremities. It was called St. Anthony’s Fire in the Middle Ages and killed tens of thousands in Europe. It has been suggested that the visions of the girls of Salem, Massachusetts, and the frenzied response of others in the famous witch hunts of 1692 were caused by ergot-tainted rye! Twenty innocent townsfolk were executed.

—[www.crimelibrary.com/notorious\\_murders/not\\_guilty/salem\\_witches/1.html](http://www.crimelibrary.com/notorious_murders/not_guilty/salem_witches/1.html)

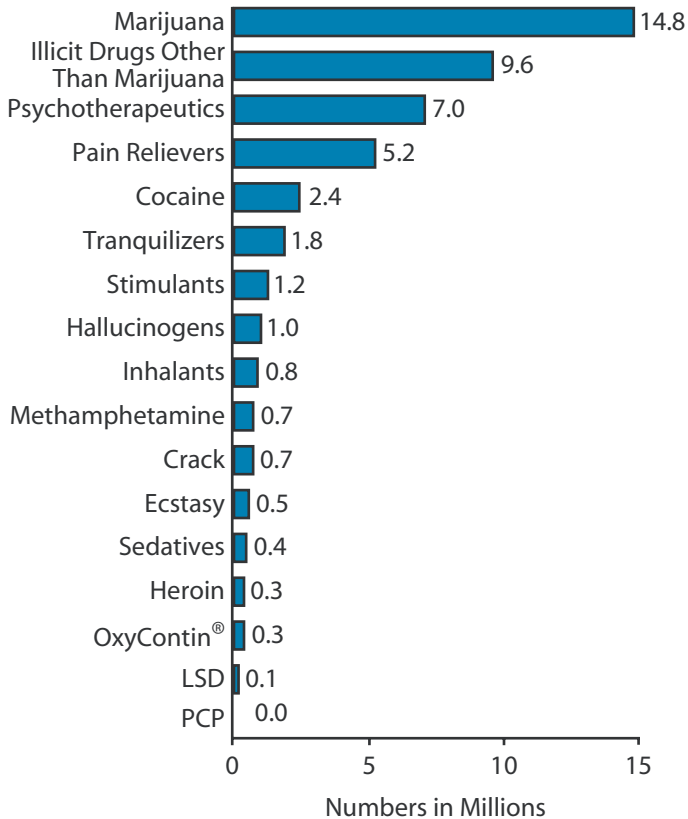
**controlled drugs:** material whose possession is controlled by law

**cannabis:** a tall, annual hemp plant native to central Asia from which marijuana is harvested; also refers to the harvested product



Blotter tabs of LSD

**teratogen:** an agent that can cause birth defects in an embryo or fetus. Two well-known examples are alcohol and thalidomide.



Typical illicit drug use, 2006.  
From <http://oas.samhsa.gov>.

colored tablets (“microdots”), or in small pieces of dried gelatin (“window panes”). It is an extremely dangerous drug that can cause psychosis, flashbacks (hallucinogenic episodes), and impaired memory and attention span. It is also a possible

**teratogen.**

PCP (phencyclidine) was sold as an intravenous anesthetic, but it was dropped because of its bizarre side effects, including insensitivity to pain, feelings of super strength, rage, memory loss, and paranoia. It is often mixed with other drugs, for example, with marijuana (“wobble-weed” or “Jim Jones”) or with LSD or amphetamines (“angel dust”).

MDMA (methylenedioxymethamphetamine, or “ecstasy”) is a synthetic drug that both stimulates the user and causes hallucinations. Many of its less desirable effects are similar to those of cocaine and amphetamines: psychological problems such as confusion, depression, addiction, severe anxiety, and paranoia—sometimes even weeks after use—and increased blood pressure and heart rate, which may cause death. Recent research links MDMA use to long-term damage to the parts of the brain that are important in thought and memory.

Ketamine (“Special K”), used legitimately as an animal tranquilizer, is a powerful hallucinogen. It is sometimes called a “club drug,” along with MDMA, GHB, GBL, rohypnol, LSD, PCP, and methamphetamine, because it is most commonly found at nightclubs and “rave” parties. Ketamine can cause delirium, amnesia, depression, and long-term memory difficulties.

- **Stimulants** act on the central nervous system to make the user feel better and increase his or her energy alertness while suppressing

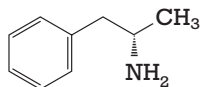
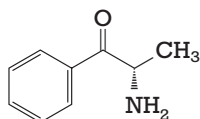
appetite and fatigue. The downside of these drugs is the accompanying restlessness and anxiety, and, after the drug wears off, depression. Stimulants vary from mild, such as caffeine, to strong, such as amphetamines (“uppers” or “bennies”) and methamphetamines (“speed” or “crank”). Cocaine is considered a stimulant. It is an **alkaloid** that comes from the leaf of the coca plant, grown almost

**alkaloid:** one of a class of bitter-tasting, basic organic compounds with nitrogen-containing rings, which are normally obtained from plants. Alkaloids often have powerful effects on living things. Examples are cocaine, nicotine, strychnine, caffeine, and morphine.

## In Yemen, drug dependency grows

**Chewing qat leaves is a national pastime, with few opposing its usage**

The primary active stimulant in qat is cathinone, which is structurally and chemically similar to



The structure of cathinone (top) and d-amphetamine

d-amphetamine. Fresh leaves are wadded in the cheek and chewed to release the stimulant, which produces euphoria.

It is estimated that 80 percent of the inhabitants of Yemen are stoned every day! This decreases productivity and diverts resources and income, deeply influencing Yemeni life and culture.



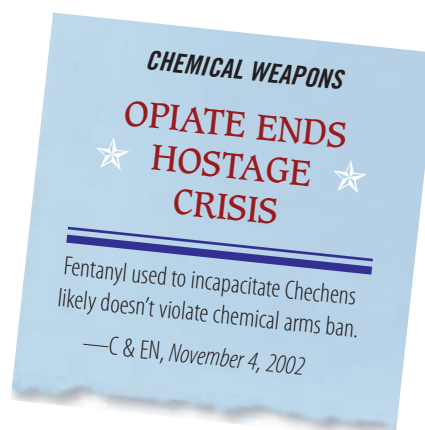
Qat (khat)

—from news article, 2001

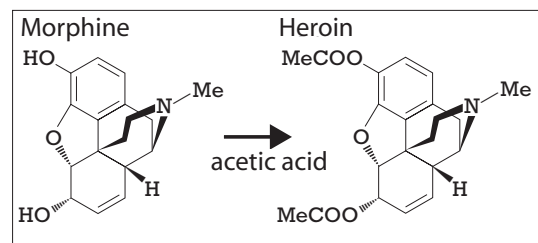
exclusively on the Amazon slopes of the Andes. It takes about 500 pounds of coca leaves to produce one pound of cocaine powder. Most cocaine used to be snorted as powder, but now smoking the recrystallized freebase (“crack”) is more popular. Crack is highly addictive. Regular use can cause anxiety, insomnia, and weight loss and leads to paranoia and mental deterioration.

Nicotine is also a habit-forming stimulant, but it is not a controlled substance.

- **Narcotics** are **analgesics**, that is, substances affecting the central nervous system to relieve pain. Mild analgesics are found in many over-the-counter (OTC) drugs such as aspirin, Tylenol, and Motrin. Illicit narcotics come from opium, harvested from a particular type of poppy flower grown mostly in Asia. Opium contains from 4 to 21 percent morphine, often used medicinally as a powerful painkiller. Heroin (“horse” or “smack”) is easily made from morphine and is highly addictive. Overdoses can cause death; serum hepatitis is common in users. Codeine is commercially prepared from morphine and is often found in OTC drugs in Canada.



**analgesics:** drugs that relieve pain



Opium comes from the milky fluid in the ripe seed bulb of the opium poppy. It is estimated that in 2007 Afghanistan supplied 92 percent of the world's illicit production.



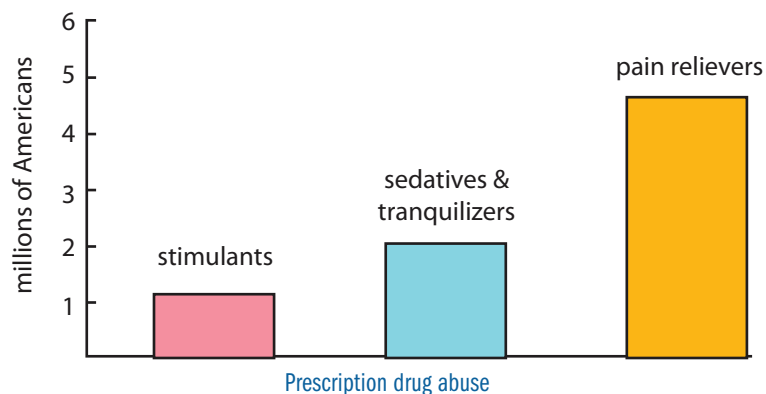
Opium poppy seed bulbs.

Synthetic narcotics, also classified as opiates, are often prescribed for pain but are also abused. Methadone is used in this country as a heroin substitute to wean addicts from heroin. Propoxyphene (Darvon) and oxycodon (Percocet) are two prescription drugs commonly abused. Fentanyl is an anesthetic and is 100 times stronger than morphine. It spawned a series of **designer drugs** known as “China White.”

**designer drugs:** drugs synthesized with particular pharmacological characteristics, designed for abuse and evasion of specific drug laws

- **Depressants:** Ethyl alcohol is a common depressant—alcohol is a \$40 billion industry in America. Alcohol, however, is not a controlled substance. The effects of overuse of alcohol will be covered in the next chapter.

In 2005, 6.4 million Americans used a prescription drug for nonmedical purposes in a period of one month (data from NSDUH).





Barbiturates (“downers”) are highly addictive; withdrawal is difficult and dangerous. Most (such as Phenobarbital, Nembutal, and Seconal) are prescribed to reduce anxiety and help the user sleep; they pose no problems in small amounts. Other synthetic depressants include meprobamate (Miltown), methaqualone (Quaaludes), and benzodiazepines (tranquilizers such as Librium and Valium).

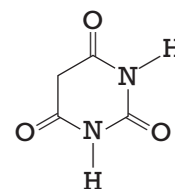


Many abused substances are controlled by prescription. Legitimately manufactured pills, tablets, and capsules can often be identified using the *Physicians' Desk Reference (PDR)*, which is an illustrated dictionary of medicines and drugs. A typical illustration and description—in this case, of oxycodone—is shown in Figure 7.1.

There are many different editions of this reference, which is updated each year; usually one can be found in a town or city library or a bookstore. The school nurse may have previous years' copies available.

Investigators test drugs and poisons with gas, paper, and thin-layer chromatography (TLC); UV and infrared spectrophotometry; mass spectrometry; spot tests; and qualitative analysis.

Some over-the-counter (OTC) analgesics can be harmful when taken in excess or combined with other medications or alcohol. For example, some people are allergic to aspirin, which can cause their throat to swell, even suffocating them. People with asthma and chronic sinusitis are more likely to have this extreme reaction. The same symptoms can occur with



Barbituric acid from which commercial barbiturates are derived. Sodium pentothal, the “truth serum,” is also a derivative.

## Percocet

*Pronounced: PERK-o-set*

*Generic ingredients: acetaminophen, oxycodone hydrochloride*

*Other brand names: Roxicet, Tylox*



Figure 7.1 PDR reference

NSAIDS (nonsteroidal anti-inflammatory drugs) such as Advil or Motrin. See Table 7.1 below for a list of some OTC analgesics.

**Table 7.1: Some OTC Analgesics and Cold Remedies**

Trade Name	Active Ingredients
Actifed	pseudoephedrine hydrochloride, triprolidine hydrochloride
Advil	ibuprofen
Aleve	naproxen sodium
Anacin	acetylsalicylic acid, caffeine
aspirin	acetylsalicylic acid
Benadryl	diphenhydramine hydrochloride
Contac	acetaminophen, pseudoephedrine hydrochloride, dextromethorphan hydrobromide, chlorpheniramine maleate
Excedrin	acetaminophen, acetylsalicylic acid, caffeine
Motrin	ibuprofen
No-Doz	caffeine
Sudafed	pseudoephedrine hydrochloride
Tylenol	acetaminophen
Tylenol PM	acetaminophen, diphenhydramine hydrochloride

See [www.nlm.nih.gov/medlineplus/druginformation.html](http://www.nlm.nih.gov/medlineplus/druginformation.html) for brand names, chemical names, proper use, precautions, side effects, and the like.

Anabolic steroid abuse has been associated with a wide range of adverse side effects ranging from some that are physically unattractive, such as acne and breast development in men, to others that are life-threatening, such as heart attacks and strokes, even in athletes younger than 30. Use of steroids can also cause increased irritability and aggression. Anabolic steroids are schedule III drugs (see Appendix C) and are illegal. Simple possession of illicitly obtained anabolic steroids carries a maximum penalty of one year in prison and a minimum \$1,000 fine for the first offense.



Testosterone

## Spot Tests

A spot test is a chemical reaction that occurs when a particular substance is added to an unknown. Color reactions for spot tests on drugs are a relatively easy and quick method used to detect the presence of certain substances. You can take this procedure into the field for preliminary testing. These tests are not conclusive, however, because a few substances may give *false-positive* results, indicating drug existence even though no drugs are present. If you get a positive result, you can use further tests, such as chromatography or infrared **spectrophotometry**, to confirm that the drug is found in the sample. A negative test result indicates that the drug in question is not present.

### Reminder

*false positive*: a test result that comes out positive when the material being tested for is not present

### spectrophotometry:

measurement of the absorption of electromagnetic radiation of a substance at different wavelengths of the spectrum

## Spot Test Lab

## Laboratory Activity 7.1

This activity will give you practice in identifying some over-the-counter drugs with spot tests.



Color spot tests

## Materials

For each lab group:

- aspirin
- Alka-Seltzer
- sodium bicarbonate
- acetaminophen
- Excedrin
- stereomicroscope
- distilled water
- Beral pipette
- universal indicator
- 0.6 M HCl
- 2 percent ferric chloride solution
- spot plates
- spatula
- toothpicks
- unknown powders #1 and #2
- mortar and pestle

If you do not use distilled water, test your water with the universal indicator.



### SAFETY ALERT! CHEMICALS USED

*Always wear goggles and an apron when working in the laboratory*

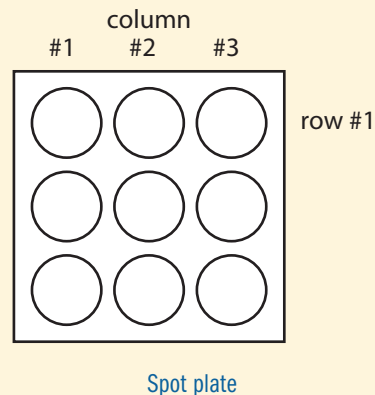


**SAFETY NOTE** *Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.*

## Procedure

Do not write in your textbook. Take notes in your science notebook.

1. In the first horizontal row of the spot plate, place a small amount of aspirin (no bigger than a grain of rice) into each of the three depressions. Label the row.
2. Repeat step 1 for each of the other five powders. Label each row.
3. Examine each of the powders, noting color and texture. Make a data table and record your observations on it.
4. Examine each of the powders under the microscope. Record your observations.
5. Add five drops of distilled water to each powder in column #1 of the spot plate. Record your observations.
6. Add one drop of universal indicator to each of the depressions in column #1. Use a different toothpick to stir each one. Record the color and pH of each powder. Note whether the substances are acidic or basic.





7. Add two drops of HCl to each of the powders in column #2. Record your observations.
8. Add two drops of the ferric chloride solution to each of the powders in column #3. Use a different toothpick to stir each one. Record your observations.
9. Carefully discard all solutions into the sink and rinse the plate with water.

### Analysis Questions

1. Aspirin (acetylsalicylic acid), the most widely used drug in the world, is usually taken as a pain reliever. An acidified solution of ferric chloride can be used to detect the presence of aspirin in an unknown powder. The aspirin hydrolyzes to form acetic acid, and the ferric ion reacts with the salicylic acid to form a compound of what color?
2. Acetaminophen, a widely used pain reliever, is not acidic; therefore, it is often taken by people who cannot tolerate aspirin. How can you tell that a sample contains acetaminophen?
3. Antacids are slightly basic compounds used to treat hyperacidity, too much hydrochloric acid in the stomach. Many of these products contain carbonates that react with or neutralize the acid in the stomach to produce a salt, water, and carbon dioxide gas. How can you tell that a sample contains an antacid (sodium bicarbonate)?
4. Alka-Seltzer contains sodium bicarbonate, citric acid, and aspirin; it reacts with water to produce carbon dioxide gas. How can you tell that Alka-Seltzer is present in a sample?
5. Excedrin is a mixture of aspirin, acetaminophen, and caffeine. What would be a good test for Excedrin? Would you need more than one test?
6. Use what you know about color spot tests to identify unknown powders #1 and #2.

Drug possession defendant Christopher Jansen, on trial in Pontiac, Michigan, said he had been searched without a warrant. The prosecutor responded that the officer didn't need a warrant because a bulge in Jansen's jacket could have been a gun. Nonsense, said Jansen, who happened to be wearing the same jacket that day in court. He handed it over so the judge could see it. The judge discovered a packet of cocaine in the pocket and laughed so hard he required a five-minute recess to compose himself.

## Is It Ibuprofen?

At the international arrivals area of the Detroit airport, a random inspection of a passenger's suitcase reveals a film canister containing a white powder. The passenger indignantly says it is ibuprofen, powdered for quick dissolving as a headache remedy, and that he never travels without it. Customs agents need to have a quick spot test done because they don't know whether to call the DEA or let the man go. You run a quick test. What do you advise them to do?



### Materials

For each lab group:

- reagents and materials from the previous lab
- unknown white powder
- ibuprofen standard



#### **SAFETY ALERT! CHEMICALS USED**

*Always wear goggles and an apron when working in the laboratory*



**SAFETY NOTE** Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.

### Procedure

Discuss with your teacher how you would solve this problem. Write a procedure and run the tests. Remember, you must support your conclusions.

## Thin-Layer Chromatography

Forensic scientists use several types of confirmatory tests based on chromatography.

Chromatography is a way to separate the components of a mixture. There are many types of chromatography: gas chromatography (GC), liquid chromatography (LC) and high-performance liquid chromatography (HPLC), paper chromatography, and thin-layer chromatography (TLC).

In all types of chromatography, there is contact between a stationary phase and a mobile phase, which cannot be mixed. The sample is carried through the chromatographic system by the mobile phase and continuously interacts with the stationary phase. The interactions of each component in the sample are based on their physical and chemical properties.

Separation takes place as each component is repeatedly adsorbed on or desorbed from the stationary phase at different rates. Think about an oily mixture of spheres and cubes poured onto a slanted plane. Which will land at the bottom first, the spheres or the cubes? Why? The components of this mixture were separated because of their shape. What is the mobile phase? What is the stationary phase?

In the stationary phase of TLC, a thin layer of adsorbent particles is attached to an inert substrate. In the following activity, you will use silica ( $\text{SiO}_2$ ) on plastic, aluminum, or glass. A small amount of a sample is applied ("spotted") near the bottom of the plate, and the plate is placed in a beaker so that the mobile phase can be drawn up the plate by capillary action. Each part of the sample that is soluble in the mobile phase will interact differently with the stationary phase, creating individual areas or spots on the plate (see Figure 7.2). Often the spots are colorless, and you must develop them to make them visible, as with latent fingerprints.

The retention factor,  $R_f$ , of each component is an individual characteristic used to



Spotting a TLC plate

*Chromatography* comes from the Greek *chroma*, color, and *graphe*, writing. A chromatogram is the record of the separation.

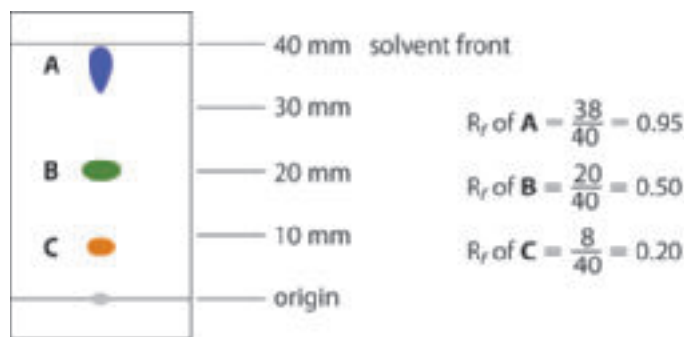


Figure 7.2 Chromatogram with retention factors

compare the components in the various samples. You can find the  $R_f$  values for various substances in reference manuals. In your lab, you simply want to compare the  $R_f$  values of the suspected powders with the known standards:

$$R_f = \frac{\text{distance from origin to spot}}{\text{distance from origin to solvent front}}$$

### 7.1: "The Drugs Made Me Do It"

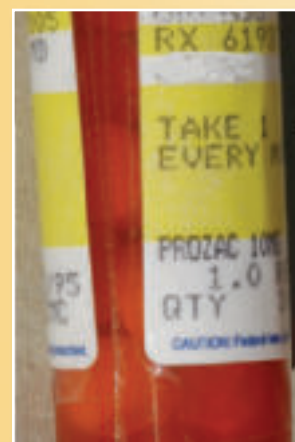
What are the legal aspects of violent behavior associated with the ingestion of prescribed drugs? Let's use Prozac as an example.

Prozac was introduced in 1987 by Eli Lilly & Company as an antidepressant. It is one of a type of antidepressants known as SSRIs (selective serotonin reuptake inhibitors), which includes Paxil, Zoloft, Luvox, Celexa, and Effexor. Prozac became so popular that 4.5 million Americans had taken it by 1992.

In a small percentage of the population, side effects included suicide, violence, and other criminal acts. Sometimes the behavior can be quite bizarre. A man taking Prozac as well as the tranquilizer Xanax robbed the bank where his wife worked, disguised with only a fake mustache. He drove off from the front of the bank in his easily identifiable vintage automobile in a hail of bullets. At trial, the judge attributed this irrational behavior to the drugs and found him not guilty.

Indeed, a claim of involuntary intoxication can be admitted under the law in many states as a mitigating factor in a criminal case when there is no history of drug abuse or prior violence. This does not excuse the offense, but it allows the judge to reduce the sentence, advise psychiatric help, or even entertain a plea of temporary insanity.

### CASE STUDY



'PROZAC MADE  
OUR FATHER  
★ KILL OUR ★  
MOTHER THEN  
HIMSELF'

COURT FINDS  
★ PROZAC ★  
AND XANAX  
CAUSE CRIMINAL  
CONDUCT

PROZAC'S  
★ ROLE IN ★  
MAUI DEATHS  
GOING TO COURT

PROZAC  
BLAMED FOR  
★ YOUTH'S ★  
MURDER SPREE



While working quietly at his desk, N. Ron Leigh, an executive of a multinational energy company, was stricken with what appeared to be a severe allergic reaction. His throat swelled to such an extent that he could hardly breathe. Rapid and effective response to a 911 call saved his life. Both the paramedics and the police wondered what might have brought on such a sudden attack. A bottle of Tylenol was found on Mr. Leigh's desk. It was well known that he suffered from asthma, bouts of sinusitis, and, lately, tension headaches.

Several caplets from the Tylenol bottle have been submitted to your lab for analysis with the thought that perhaps they were not Tylenol but that someone had substituted an aspirin or ibuprofen product, hoping to cause a severe allergic reaction.

### Materials

For each lab group:

- aspirin
- Tylenol
- Motrin
- No-Doz
- Excedrin
- test tubes
- capillary tubes
- methanol
- iodine crystals in a covered beaker
- TLC plates
- UV light
- eluting solvent
- beaker large enough to hold TLC plate
- 2 percent  $\text{FeCl}_3$  solution
- metric rulers
- mortar and pestle



### SAFETY ALERT! CHEMICALS USED

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**SAFETY NOTE** Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.

### Procedure

Do not write in your textbook. Take notes in your science notebook.

1. Pulverize (pound into powder) a tablet from each of the following known standards:

aspirin	No-Doz
Tylenol	Excedrin
Motrin	

### Reminder

The known sample originates from a verifiable or known source; it can also be called the exemplar in forensic science.

The questioned, or unknown, sample has been collected from a known location but is of unknown origin.

### Reminder

*Elute* means to separate one material from another with a solvent.

### Reminder

A fluorescent material absorbs light of a shorter wavelength (often in the ultraviolet) and gives off light of a longer wavelength (often in the visible part of the spectrum).

## Laboratory Activity 7.3, *continued*

2. Pulverize a tablet from the Tylenol bottle on Ron's desk, the questioned sample.
  3. Label six 10 × 75-mm test tubes.
  4. Add a small amount, equivalent to a grain of rice, of each sample to a test tube.
  5. Add five drops of methanol to each sample.
  6. Add five more drops of methanol to the known aspirin and Tylenol samples.
  7. Agitate each test tube for a minute or so. *Hint:* To agitate a small test tube without putting a stopper in it, simply hold its neck between your thumb and forefinger and flick it repeatedly with your middle finger. Be careful not to splash out the contents. The methanol will dissolve the active ingredients and leave behind the starch, cellulose, and other products used as "fillers" to form the pills.
  8. Prepare a 10 × 10-cm TLC sheet by lightly drawing a pencil line across it 1 cm from the bottom. Lightly mark six tick marks equally spaced along the line for each sample, and label.
  9. Dip an open-ended capillary tube into each of the sample test tubes.
  10. Gently spot each sample on the appropriately labeled mark. Be careful not to scratch the silica coating.
  11. After each spot dries, repeat the spotting two or three times. You can make the spots dry faster by placing the TLC sheet on a hot plate set on "low."
  12. The eluting solvent has already been prepared. Pour up to ½ cm into an 800- to 1,000-ml beaker. Do not put enough solvent in the beaker to cover any of the spots. You can line the beaker with filter paper, as in Figure 7.3. Separation is quicker in the solvent atmosphere provided by the soaked filter paper.
- Two TLC plates, hinged together at the top with tape, form a stable tent to place in the beaker. Cover the beaker and allow the solvent to rise to within 1 cm of the top. This should take about 15 to 20 minutes. Then remove the sheets, mark the solvent front, and let the chromatograms dry.
13. Look at the chromatogram under short-wave ultraviolet light. Outline each spot that you see. The TLC plate being used in this experiment has a fluorescent substance added to the silica. The spots on the chromatogram block the fluorescence. *Caution: Do not look directly into the light; it can damage your eyes.*
  14. To develop the spots so they are visible, put each plate in a beaker with a few crystals of iodine. Cover the beaker and watch to see what happens.



**Figure 7.3** Lining a beaker with filter paper

Note which spots become totally cloudy. *Caution: Iodine vapor is poisonous; be careful.* After using iodine to make the spots visible, blot each spot with a paper towel wet with  $\text{FeCl}_3$  solution. Write down all your observations in your lab book and draw the chromatograms. Calculate the  $R_f$  values and present your data in a table.

### Analysis Questions

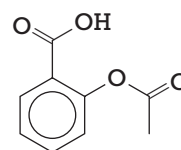
1. What can you say about the identity of the questioned material? Support your conclusions.
2. What is the mobile phase in this experiment?
3. What is the stationary phase in this experiment?
4. What is meant by an eluting solvent?

## ADVANCED TOPICS in the Analysis of Drugs

A **plasma** sample taken from N. Ron Leigh has been submitted to the lab for analysis. Did Mr. Leigh actually take aspirin, and if so, how much?

The salicylate ion that comes from aspirin forms a violet or slightly red complex with the  $\text{Fe(III)}$  ion. The intensity of the color is directly related to the concentration of salicylate, and thus aspirin; it can be measured using a UV-VIS spectrophotometer (Spec 20 or equivalent).

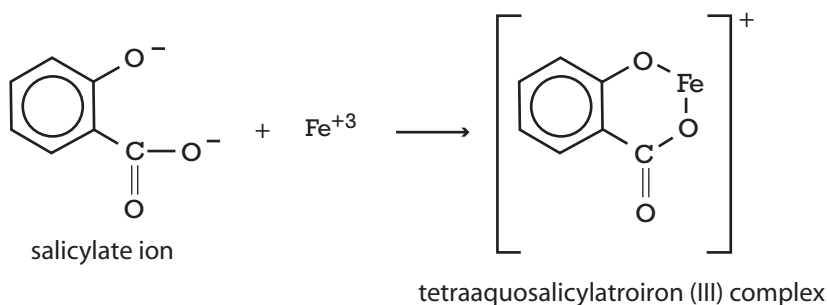
Thus, with the proper information, you can determine the number of aspirin tablets taken by Mr. Leigh.

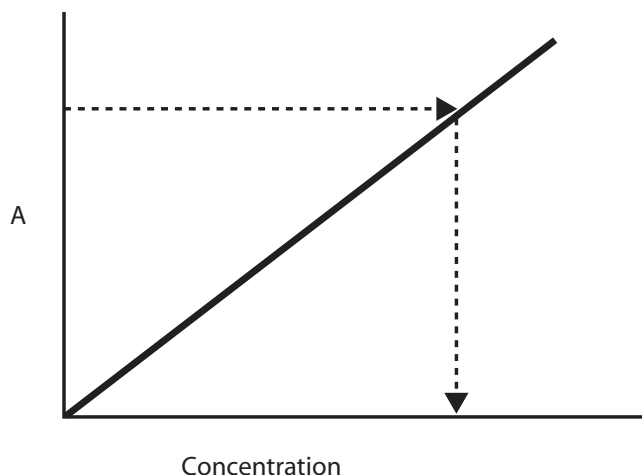


Acetylsalicylic acid  
(aspirin)

**plasma:** the fluid portion of blood, obtained by centrifuging a whole blood sample

The charge on an ion can be expressed in two ways:  $\text{Fe(III)}$  or  $\text{Fe}^{+3}$





**Figure 7.4** A calibration curve

Beer's law describes the linear relationship between absorbance and concentration, discovered by August Beer in 1852:

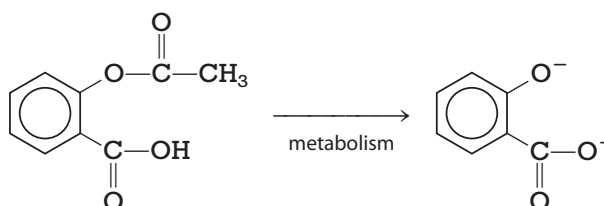
$$A = abc = \log [1/T]$$

where  $A$  = absorbance,  $T$  = transmittance,  $a$  = a constant dependent on the dissolved substance (solute),  $b$  = the path length of light (width of cuvette), and  $c$  = concentration.

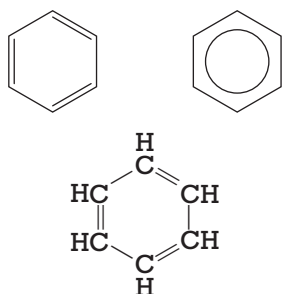
A linear calibration curve is also called a Beer's law plot.

In the following activity, you will prepare a set of standards, that is, solutions of different known salicylate ion concentrations. Each of these will be complexed with an Fe(III) solution, and the amount of light absorbed at a fixed wavelength will be measured. You can then draw a calibration curve (which is more often than not a straight line) relating the intensity of color, that is, how much light the solution absorbs (absorbance =  $A$ ) to the concentration of salicylate (see Figure 7.4).

You can complex the salicylate in a sample of blood plasma taken from Mr. Leigh at some time after the incident and measure the absorbance. You can then read the concentration from the calibration curve. Acetylsalicylic acid (aspirin) is metabolized in the body to salicylate ion at a known rate:



There are several ways to write structural formulas; three examples show the molecule benzene:



Thus, with the proper information, you can determine the number of aspirin tablets taken by Mr. Leigh.



### Materials

For each lab group:

- Spec 20 or equivalent spectrophotometer
- 25-ml beakers or flasks
- cuvettes or test tubes
- sodium salicylate
- test tube rack
- $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$
- sodium salicylate standards
- iron(III) stock solution ( $\text{FeCl}_3$ )
- 5-ml pipette
- tissues
- graph paper
- Mr. Leigh's plasma sample (the unknown)



### SAFETY ALERT! CHEMICALS USED

*Always wear goggles and an apron when working in the laboratory*



**SAFETY NOTE** *Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.*

### Procedure

Do not write in your textbook. Take notes in your science notebook.

Standard solutions have been prepared from sodium salicylate that will yield the following concentrations of salicylate ion when mixed with exactly 5 ml of Fe(III) stock solution:

100 ppm

80

60

40

20

0 "Blank"

unknown (Mr. Leigh's plasma sample)

1. Pipette 5.00 ml of Fe(III) stock solution into each labeled container of salicylate solution. Swirl to mix and develop the color.
2. Read the directions on use of the UV-VIS spectrophotometer.

Try different wavelengths to determine the best absorption wavelength for your sample. The salicylate-iron complex looks violet-red because it absorbs all wavelengths of the visible spectrum except those in the violet and red. These colors are transmitted and reflected, and that is what you

100 parts per million (ppm) = 0.01% = 1 part in 10,000

### Laboratory Activity 7.4, *continued*

see. To choose the best wavelength for absorption by the complex, you will have to scan one sample at different wavelengths, looking for the particular wavelength that gives the maximum absorbance.

Remember that the spectrophotometer responds differently when the wavelength changes, so you must reset the display to 0.00A with the blank solution every time you change the wavelength.

3. Choose a wavelength of 450 nm. Fill a cuvette (test tube specifically for use with the spectrophotometer) half full with the blank solution. Wipe the test tube with a tissue to remove liquid, dust, and fingerprints.
4. Put the test tube in the sample compartment, lining up the guide mark on the tube with the guide mark at the front of the sample compartment.
5. Close the lid of the sample compartment.
6. Adjust the display to 0.00A by turning the %T/A selector knob.

7. Take the blank out of the sample compartment and save it.

8. Fill another test tube with the 60-ppm solution, wipe it off, and put it in the sample compartment, lining up the guide marks.

9. Close the lid of the sample compartment.

10. Record the absorbance,  $A$ , from the display.

11. Take the test tube out of the compartment and save it for subsequent measurements.

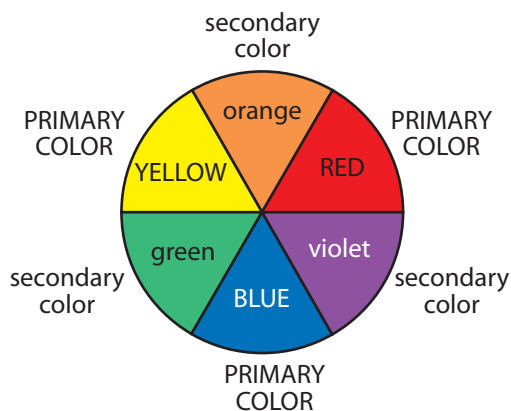
12. Choose 470 nm as the next wavelength, and repeat steps 3 through 7.

13. Put the 60-ppm solution in the sample compartment again and line up the guide marks. Repeat step 9 and so on every 20 nm to 550 nm.

14. Plot  $A$  versus wavelength ( $\lambda$ ) for an absorption spectrum. Estimate the optimum wavelength for measuring the absorption of the salicylate-iron complex. Now you are ready to prepare a calibration curve.

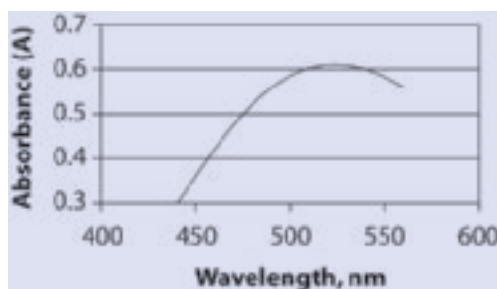
Look at the color wheel; if the solution appears violet-red, then the light of blue, green, yellow, and orange wavelengths is absorbed. The absorption maximum will be in the complementary green-yellow part of the visible spectrum, barring other absorbers.

See <http://science.csustan.edu/tutorial/color/index.htm> about colors, absorption, and chromatography.



15. Set the wavelength on the spectrophotometer to the best one found in the previous exercise.
16. Put the blank you used previously into the sample compartment, lining up the guide mark on the tube with the guide mark at the front of the sample compartment.
17. Close the lid of the sample compartment.
18. Adjust the display by turning the %T/A selector knob.
19. Take the blank out of the sample compartment and save it.
20. Fill a test tube with the 100-ppm salicylate standard, wipe it, and put it in the sample compartment, lining up the guide marks.
21. Close the lid of the sample compartment.
22. Record the absorbance,  $A$ , from the display.
23. Take the test tube out of the compartment and put in the one containing the 80-ppm salicylate standard. (Because the wavelength is fixed, you don't need to run a blank each time.)
24. Repeat the above steps until all the standards have been measured.
25. Plot absorbance ( $A$ ) of salicylate versus concentration of salicylate, in ppm (see Figure 7.4). This is the calibration curve you will use to determine the concentration of salicylate in the plasma sample taken from Mr. Leigh three hours after he allegedly took the aspirin in the Tylenol bottle.
26. Measure the absorbance of Mr. Leigh's plasma sample.
27. Using your calibration curve, determine the concentration of salicylate in the plasma sample. Remember that this sample was diluted by a factor of two when the Fe(III) solution was added.

The absorption spectrum for the complex should look something like this:



## The Metabolism of Aspirin

The **metabolism** of acetylsalicylic acid takes place rather rapidly in the body; you can follow this process by sampling urine or blood plasma at different times after ingestion. Figure 7.5 shows how the relative concentration of salicylate ion changes over time in an average adult male.

**metabolism:** the sum of all chemical processes occurring in an organism. Metabolites are organic molecules involved in the process of metabolism; they can form more complex molecules, or they may result from the degradation of compounds.

You must adjust the concentration of salicylate you found in Mr. Leigh's plasma to correct for the time because some of it has been metabolized to products other than salicylate. Use the graph in Figure 7.5 to determine the maximum amount of salicylate that was in Mr. Leigh's plasma. Record your result.

The number of aspirin tablets that would have produced the maximum amount of salicylate found in the plasma sample can be estimated from Figure 7.6. Calculate how many aspirin tablets Mr. Leigh took. Does this number make sense?

Write a report analyzing the incident with N. Ron Leigh. Your report should include a summary of what happened as well as a brief description of the laboratory tests and the results and your opinion about those results. This report may be used as evidence if this case comes to trial, so it must be defensible in court.

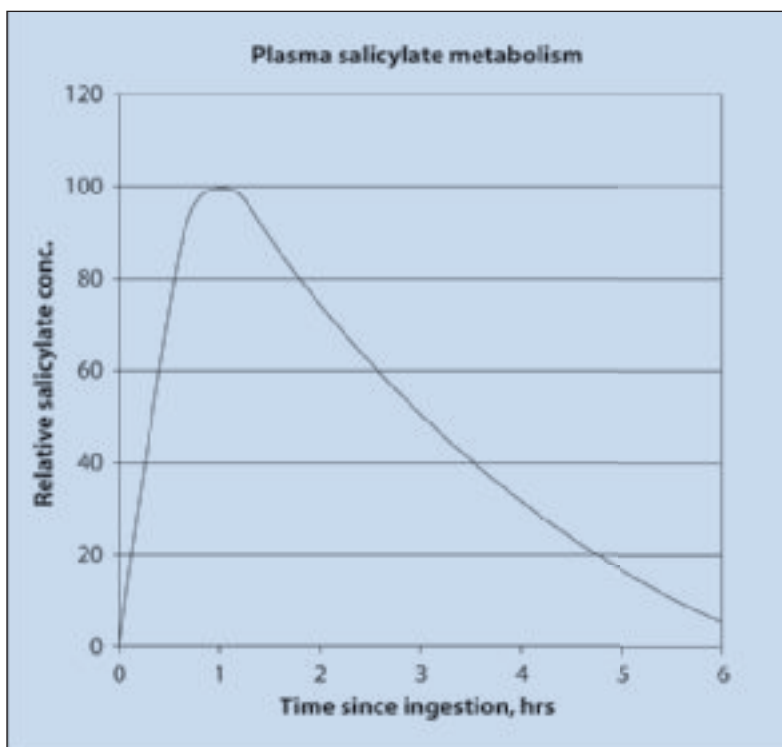
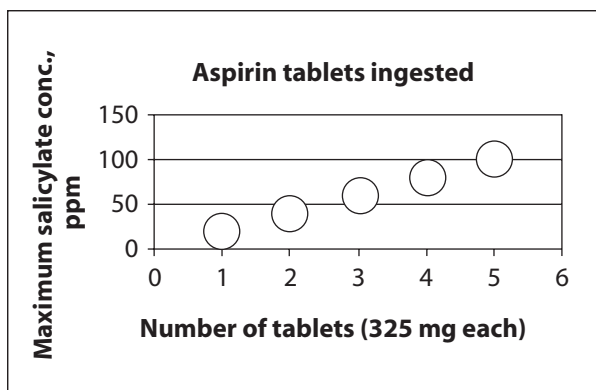


Figure 7.5 Changes in plasma salicylate concentration with time



**Figure 7.6** Relationship between plasma salicylate and ingested acetylsalicylic acid

## Testing for Marijuana

A **presumptive test**, or screening test, for illicit drugs is based on spot tests that produce specific colors for specific drugs. The most commonly used test for marijuana is the Duquenois-Levine test, developed in 1941, which creates a purple color when the active ingredients of marijuana are present in a sample.

**presumptive test:** test that presumes the presence of the questioned substance; also called a screening test. A blue-purple color in the test for marijuana can also be caused by other substances. In such a case, it is a false positive. A negative test result, however, confirms that there is no marijuana in the sample.

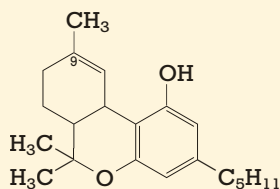
## Detecting Marijuana

### Laboratory Activity 7.5

A police officer pulls a car over for a minor traffic violation. The officer thinks she smells marijuana in the car. A search uncovers a plastic bag containing plantlike material stuffed under the front seat. The occupants of the car insist it is “stuff used to make incense, like oregano and cloves” and is strictly innocent.

A sample has been submitted to your laboratory for preliminary analysis. Does it contain marijuana?

The primary active ingredient of the cannabis plant is  $\Delta^9$ -tetrahydrocannabinol, more easily referred to as THC.





### Materials

- test tubes
- hexane
- oregano
- cloves
- alleged marijuana
- marijuana standard
- “Duquenois reagent” (simulated)
- $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$
- potassium iodide (KI)
- tea



#### SAFETY ALERT! CHEMICALS USED

*Always wear goggles and an apron when working in the laboratory*



**SAFETY NOTE** *Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.*

### Procedure

Do not write in your textbook. Take notes in your science notebook.

1. Place about 2 to 3 mm of the suspect material in a 10 × 75-mm test tube. Add just enough “Duquenois reagent” to cover the material, then add about 2 to 3 mm more. Note the color.
2. Add a volume of hexane equal to one-half the volume of the Duquenois reagent in the test tube, and shake it for a minute. A red-purple color in the aqueous phase as well as in the **immiscible** (hexane) layer at the top of the test tube indicates the presence of THC, the active ingredient of marijuana.

You should also test a control sample or standard of known marijuana (simulated in this case) exactly the same way, as well as each of the alleged ingredients of the “incense,” oregano and cloves.

3. Prepare a brief report of your results and justify your conclusion.

Spot tests like this are used just for screening. Negative results mean that the drug you are testing for is *not* present. There are, however, substances that can cause a false positive, so you should do another test to confirm so it will be conclusive as evidence. The presence of marijuana can be confirmed with a microscopic examination, if possible, or thin-layer chromatography (TLC).

#### **immiscible:**

describes materials that do not mix. When shaken, two immiscible liquids will separate to their original volumes.

#### **Dumb Crook**

Forty-five-year-old Amy Brasher was arrested in San Antonio, Texas, after a mechanic reported to police that 18 packages of marijuana were packed in the engine compartment of the car that she had brought in for an oil change. According to police, Brasher later said that she didn't realize that the mechanic would have to raise the hood to change the oil.

## Drug Tests That Use Color

As with the Duquenois-Levine test for marijuana, you can use a number of spot tests to determine the presence or absence of a particular drug or type of drug. A positive result implies the presence of the drug, but these tests are not conclusive, and they must be confirmed by a more specific test to rule out false positives.

Table 7.2 below summarizes the more common spot tests used by law enforcement agencies.

**Table 7.2: Color Tests for Drugs**

Drug	Reagent	Color
marijuana	Duquenois-Levine (D-L)	blue-violet
LSD	Erllich/Van Urk (ERL)	blue-violet
amphetamines	Marquis (MARQ)	red-orange → brown
cocaine	cobalt thiocyanate (CO)	blue flaky precipitate
heroin	Marquis (MARQ)	purple
barbiturates	Dille-Kopanyi (D-K)	violet

## Presumptive Color Tests for Drugs

### Laboratory Activity 7.6

This activity will give you some practice with testing samples for the presence of drugs.

### Materials

- spot plates
- spatula
- test reagents (ERL, MARQ, CO, D-K)
- concentrated sulfuric acid
- drug standards
- unknown drug
- toothpicks
- safety goggles
- Beral pipettes
- 2 percent  $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$
- cobalt thiocyanate
- potassium iodide
- 1-naphthol
- ethanol
- potassium ferrocyanide
- aspirin
- powdered sugar
- Benadryl
- mortar and pestle



### SAFETY ALERT! CHEMICALS USED

*Always wear goggles and an apron when working in the laboratory*



**SAFETY NOTE** Also wear disposable lab gloves. Avoid inhalation, ingestion, and skin contact with chemicals.

## Procedure

Do not write in your textbook. Take notes in your science notebook.

1. Put a small amount (the size of a grain of rice—no more!) of each simulated drug in the cavity of a white ceramic or plastic spot plate, according to the diagram shown in Figure 7.7. The LSD sample is the only one impregnated on paper, so some of the tests are not required, as indicated by the X in the circle.
2. Add *one* drop from the labeled bottle of reagent to the edge of each sample. Mix if necessary. Note and record any changes, initially and over a period of 15 minutes. The D-K test for barbiturates is a two-step process: Add the D-K reagent; then add *one* drop of concentrated sulfuric acid.

**Caution:** Concentrated  $H_2SO_4$  is a strong acid; it is severely corrosive to skin, eyes, and other tissue. Wear safety goggles!

### Reminder

A *metabolite* is a by-product of a chemical reaction in the body. For example, over 31 metabolites of marijuana can be detected in urine.

**enzyme:** a protein that acts as a catalyst in a living organism

**antibody:** a protein produced by a body's immune system that tags a molecule in order to destroy it

**EMIT:** an enzyme immunoassay test to detect particular types of metabolites in body fluids, commonly used for screening drug use

### Reminder

A *false positive* is a test result that shows a substance is present, when it is not. That is why confirmatory tests are required.

ERL	MARQ	CO	D-K	
<input type="radio"/>	<input checked="" type="radio"/>	<input checked="" type="radio"/>	<input checked="" type="radio"/>	LSD
<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	amphetamines
<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	cocaine
<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	heroin
<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	<input type="radio"/>	barbiturates

Figure 7.7 Layout for spot plate in presumptive color tests

There are a number of new preliminary or screening tests now being used to detect the use of major classes of illicit drugs. Most of these use urine samples. They are based on competition between a drug metabolite, an **enzyme**-labeled metabolite, and a specific **antibody**. These tests result in a color reaction that is proportional to the amount of drug present. The Enzyme Multiple Immunoassay Test (**EMIT**) is simple to administer and inexpensive. It is even sold for home use; however, it is subject to many false positives.

## Confirmatory Tests

In court, forensic scientists are often required to explain the testing method to the jury, so now you will learn about the more sophisticated analytical techniques for testing drugs.

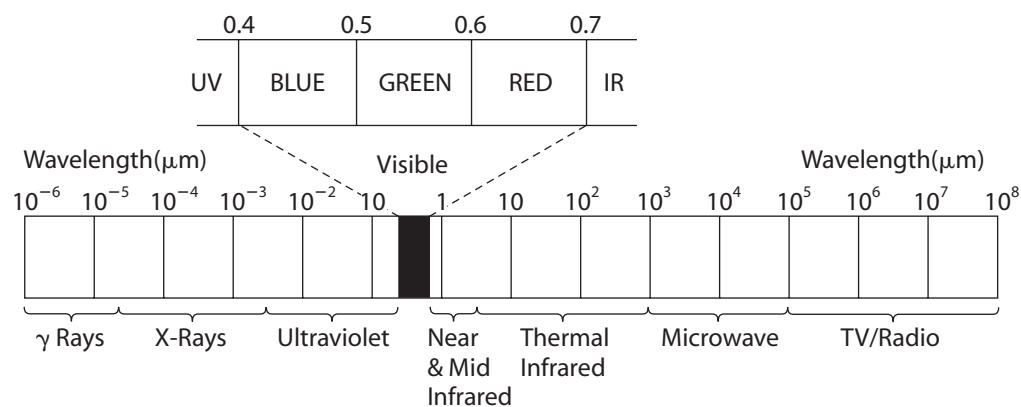
Most **confirmatory tests** for drugs are based on one of two analytical techniques, infrared spectroscopy and mass spectrometry.

**Spectroscopy** is the interaction of electromagnetic radiation with matter. In infrared (IR) spectroscopy, material absorbs energy in the near-IR region of the electromagnetic spectrum (Figure 7.8).

**confirmatory tests:** tests that specifically identify one substance

**spectroscopy:** the branch of science that involves the study of electromagnetic radiation and its interaction with matter

This range or band of energies is just enough to make the bonds that hold the atoms of a molecule together bend and stretch. Each bond, or group of bonded atoms, has a characteristic excitation energy. So, when a band of near-IR light passes through a substance, certain energies, frequencies, or wavelengths are absorbed by the many different molecular bonds. Comparing the IR light beam before and after it passes through a transparent sample results in a transmission or an absorption spectrum. The IR spectrum of each substance is a combination of the many different bonds that make up that substance, so it's not surprising to find that its IR spectrum gives a unique view of the substance, much as a fingerprint is unique to a single person.



**Figure 7.8** The electromagnetic spectrum

The equations  $E = h\nu$  and  $\nu = hc/\lambda$  describe the relationships between energy ( $E$ ), frequency ( $\nu$ ), and wavelength ( $\lambda$ ).  $h$  is a proportionality constant and  $c$  = speed of light. IR spectroscopists measure the energy absorbed by molecular vibrations as  $1/\lambda$ , called wave numbers and measured in  $\text{cm}^{-1}$ .

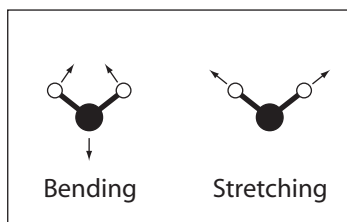
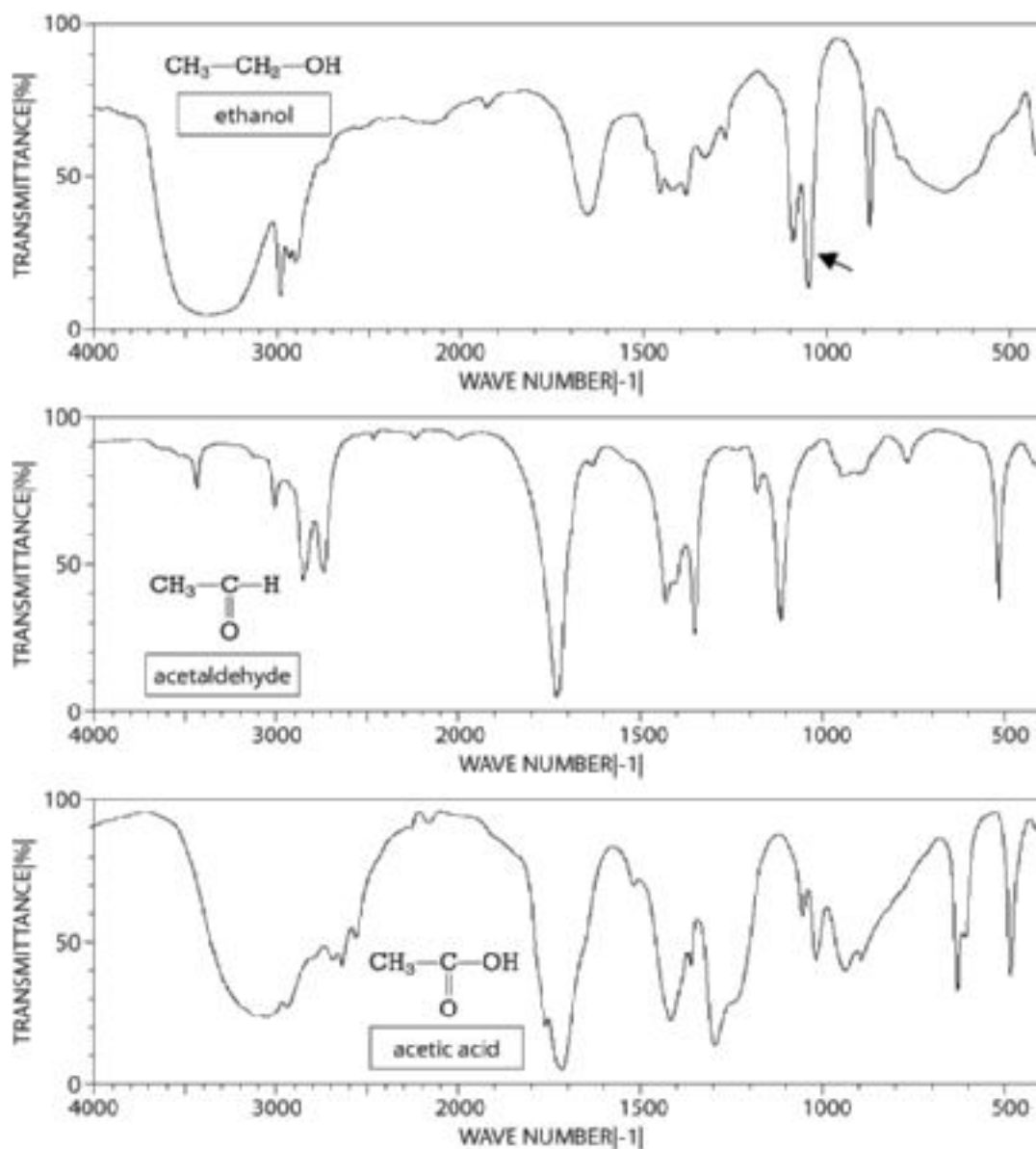


Figure 7.9 shows IR transmission spectra of ethyl alcohol and several other two-carbon molecules. Note the difference in their spectra.

The abscissa ( $x$ -axis) in the IR diagrams is a function of the energy of the IR radiation, measured in wave numbers.

The ordinate ( $y$ -axis) measures the amount of IR energy transmitted through the sample, that is, energy that is not absorbed. The molecular structure determines the position of the peak. For example, the “peak” at  $1,050\text{ cm}^{-1}$  in ethanol (ethyl alcohol) is caused by absorption of energy from stretching of the carbon–oxygen–hydrogen bond in the molecule. You can confirm identification of a substance by comparing the IR spectrum of an unknown sample with that of a known compound.

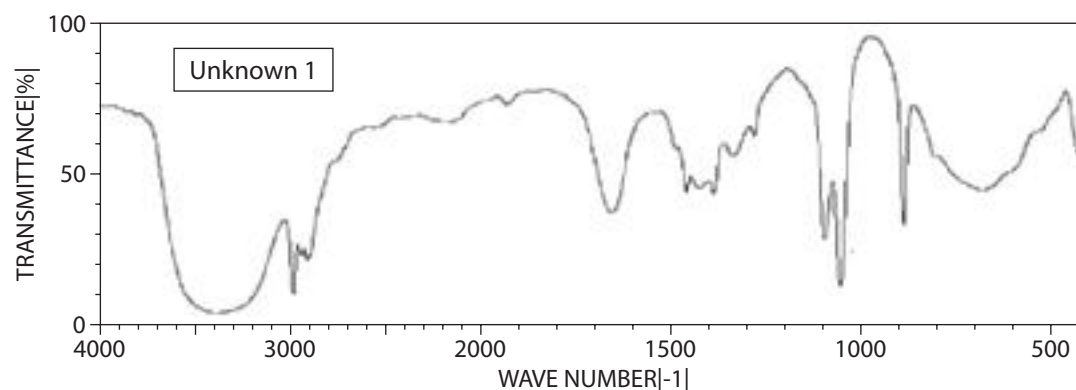


Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

**Figure 7.9** IR spectra of  $\text{C}_2$  compounds

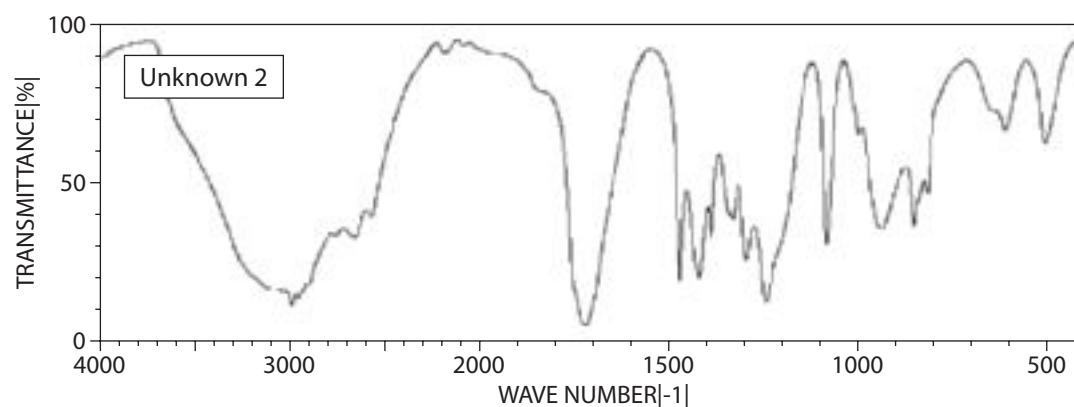


The IR spectrum of unknown liquid 1 is shown below. Is it ethanol, acetaldehyde, or acetic acid?



Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

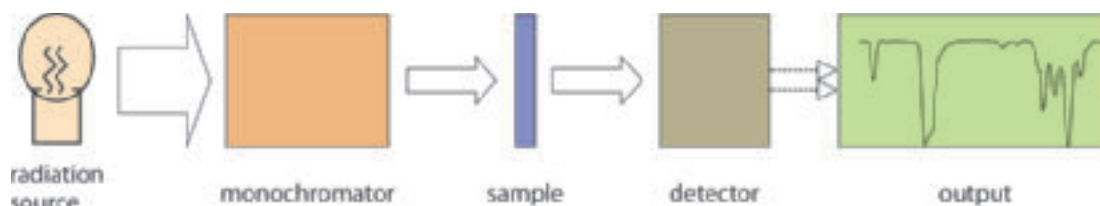
What is the nature of unknown liquid 2, whose IR spectrum is shown below? Is it an alcohol, an aldehyde, or an acid? Justify your answer.



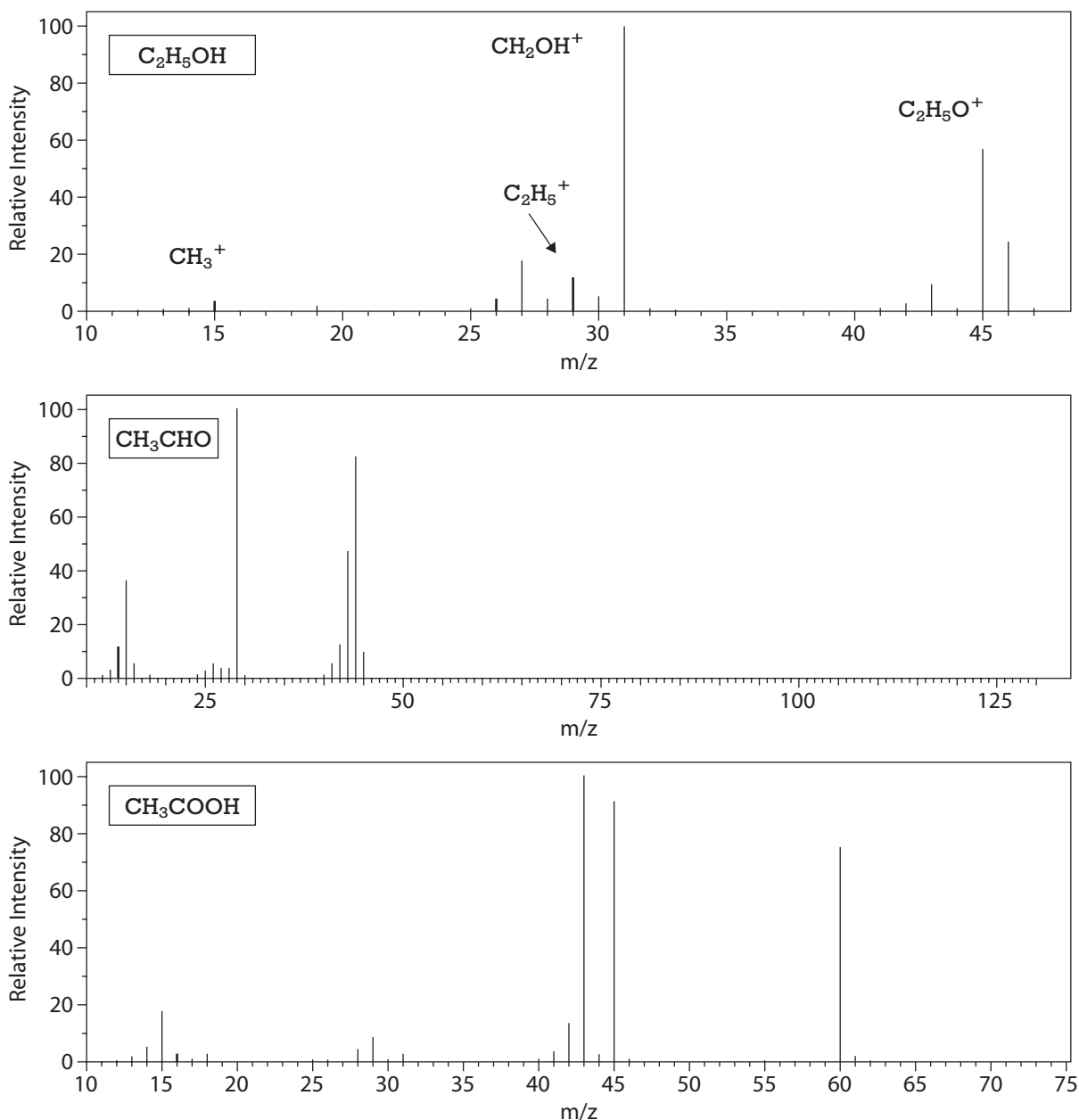
Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

The basic instrument now in use for taking IR spectra is the Fourier Transform Infrared (FTIR) spectrophotometer. A good one costs about \$25,000. A simple schematic of an IR spectrophotometer is shown in Figure 7.10.

Molecules can be identified through mass spectrometry by breaking them apart into pieces and then measuring the quantity of the different pieces. Each molecular compound has its own unique fragmentation pattern or spectrum. You can identify this pattern by comparing it to a catalog of known spectra. The mass spectra of ethanol and similar two-carbon molecules are shown in Figure 7.11. Note the differences in their spectra.



**Figure 7.10** Simple schematic of an infrared spectrophotometer



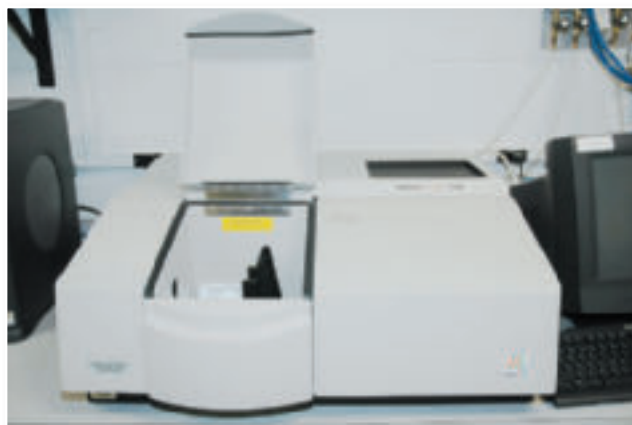
Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

**Figure 7.11** Mass spectra of C<sub>2</sub> compounds

The mass spectrometer works by bombarding the vaporized sample with electrons, which have enough energy to break a molecule

apart. Those fragments that have become positive **ions**

**ions:** atoms or molecules that have lost or gained one or more electrons and, thus, have a net positive or negative charge



FTIR spectrophotometer

(that is, the fragments that lose an electron and, therefore, have a net positive charge) are accelerated by a negative electric field through either a magnetic or oscillating electric field that sorts them according to their mass-to-charge ratio ( $m/e$  or  $m/z$ ;  $z$  or  $e$  is usually 1).

The resulting signal is a mass spectrum. Each molecular species has its own unique mass spectrum; therefore, a catalog can be created, as with IR spectra.

Three equations describing fragmentation of ethanol are:



Atomic Mass Units:

H = 1; C = 12; O = 16

To find the molecular mass, or molecular weight, of  $\text{CH}_3\text{CH}_2\text{OH}$ :

$$2 \text{ C} \times 12 = 24$$

$$6 \text{ H} \times 1 = 6$$

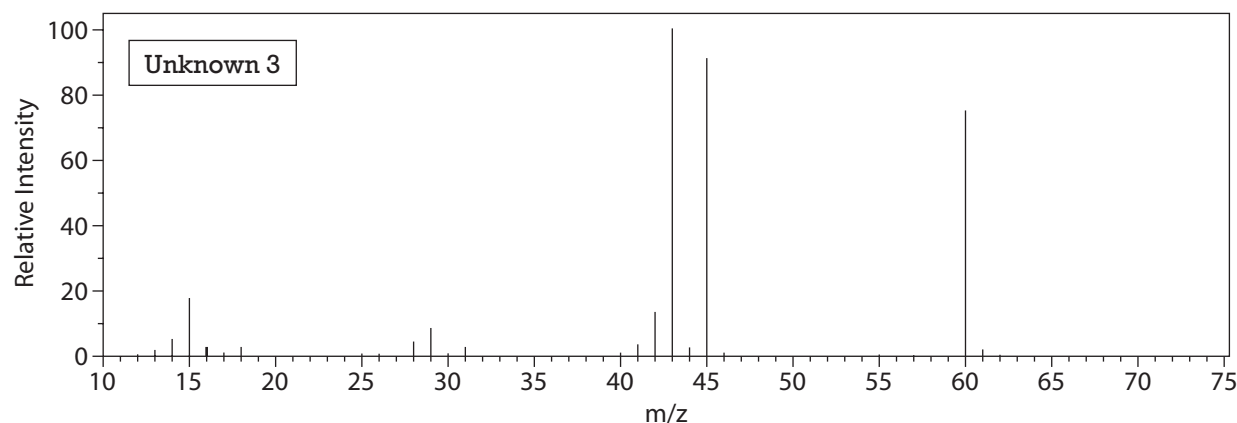
$$1 \text{ O} \times 16 = 16$$

46 units

Note which peak is related to which ion in the spectrum.

What other equations can be written to account for the peaks seen on the mass spectrum of ethanol? Account for the major peaks in the other two spectra.

What is the substance whose mass spectrum appears below?

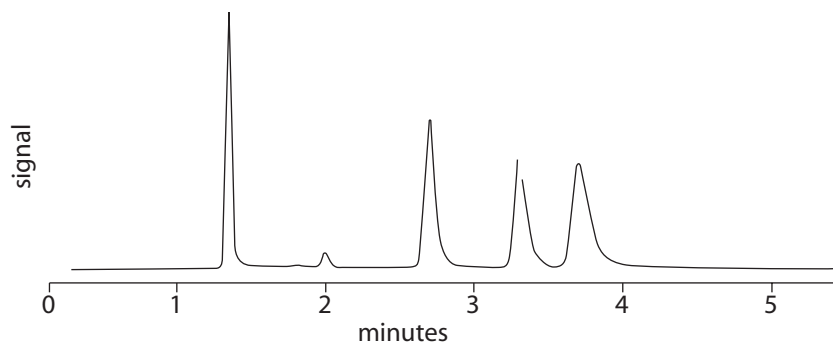


Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

Mass spectrometry works fine in identifying pure substances, just like IR spectrophotometry. However, mixtures are difficult to identify in both techniques because their spectra become superimposed.

Chromatography can separate a mixture into its component parts, but the technique does not specifically identify what the component is. In a gas chromatograph-mass spectrometer (GC-MS), a sample is vaporized and swept through a column containing some sort of adsorbent by using a “carrier gas” (usually helium), and each separated component

Illicit drugs are often diluted with other products to cut the purity, making them safer to take as well as increasing the profit for the dealer; examples of materials used to dilute these drugs include quinine, starch, lactose, and lidocaine. Dilution of an illicit drug complicates penalties based on weight of the substance (see Appendix C, for example).

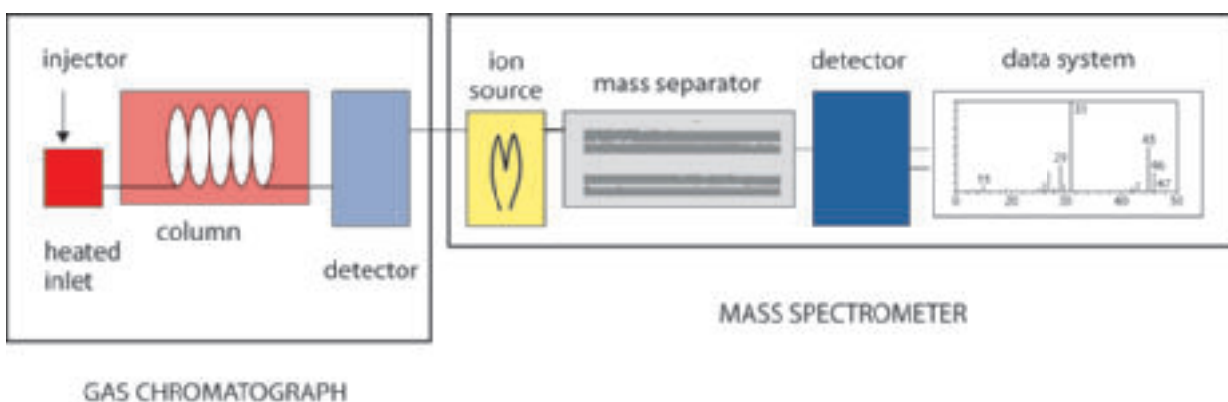


**Figure 7.12** A gas chromatogram

**analyte:** a substance being analyzed

(or **analyte**) is detected as an electronic signal when it emerges from the chromatographic column. An example of output from a gas chromatograph is shown in Figure 7.12.

Each peak represents a component of the mixture, but the analyte cannot be identified directly. You can identify an analyte by introducing each component coming from the chromatograph into the mass spectrometer. A schematic for a GC-MS is represented in Figure 7.13. A good GC-MS costs about \$60,000.



**Figure 7.13** Schematic of a GC-MS



Cocaine



Crack cocaine

## Checkpoint Questions

Answer the following questions. Keep the answers in your notebook, to be turned in to your teacher at the end of the unit.

1. What are the differences among a controlled substance, an illicit drug, a prescription drug, and an OTC drug?
2. What is a presumptive test? When can it be useful? What are its limitations?
3. How are illicit drugs classified?
4. A sample of light brown powder found in the kitchen of an alleged drug house gives a blue precipitate with cobalt thiocyanate. What is it? Is there enough evidence to prosecute?
5. What is the electromagnetic spectrum (EMS)? Which end of the EMS is the high-energy part? How is the EMS used in analytical analyses?
6. What is the difference between absorbance and transmittance in spectroscopy? What is their relationship to analyte concentration?
7. What are the types of chromatography?
8. A bag of pills was confiscated from a student. They were small, white, and heart-shaped, possibly with a numeral on one side. Should the student be arrested?
9. What are Xanax, Zithromax, and Zocor?
10. Why is Ritalin prescribed? Can it be addictive? What common side effects may occur with its prescribed use? What are some symptoms of Ritalin overdose?

### METH BILL RESTRICTING ★ COLD PILLS ★ ADVANCES IN LEGISLATURE

Sales of cold medicines used to make methamphetamine would be restricted under proposed legislation.

The new law would require stores to sell Sudafed, Nyquil, and other medicines containing pseudoephedrine only from behind the pharmacy counter.

—abstracted from the Boston Globe,  
July 29, 2005



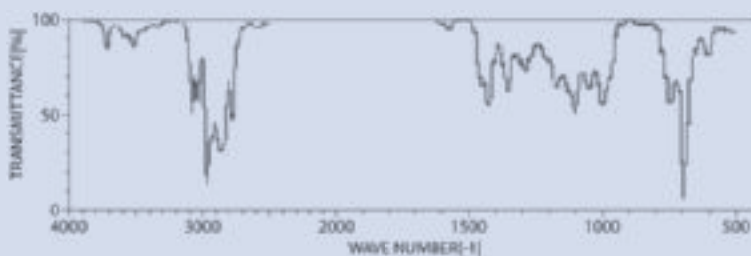
**11. Define each of the following abbreviations:**

- |        |          |           |
|--------|----------|-----------|
| a. PDR | f. MARQ  | k. PCP    |
| b. IR  | g. GC-MS | l. TLC    |
| c. OTC | h. LSD   | m. NSAID  |
| d. DEA | i. ppm   | n. $R_f$  |
| e. THC | j. AMU   | o. Co(II) |

**12. Define the following terms, with examples:**

- |                      |                        |
|----------------------|------------------------|
| a. quantitative      | k. Motrin              |
| b. presumptive       | l. molecular structure |
| c. spectrum          | m. chromatography      |
| d. metabolism        | n. confirmatory test   |
| e. absorbance        | o. analyte             |
| f. spectrophotometer | p. questioned sample   |
| g. a standard        | q. narcotic            |
| h. a stock solution  | r. allergic reaction   |
| i. an ion            | s. barbiturate         |
| j. wavelength        | t. false positive      |

**13. You have tested a sample of a white powder with cobalt thiocyanate reagent and obtained a blue precipitate. You think you have \_\_\_\_\_, but you ask the spectroscopy lab for a confirmation. They provide the following spectrum:**

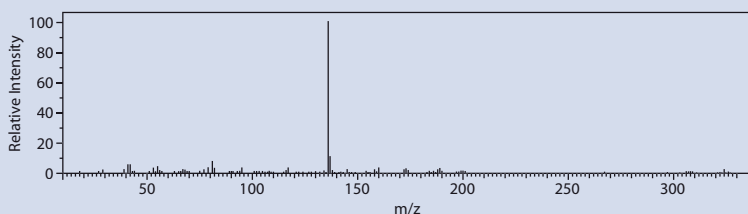
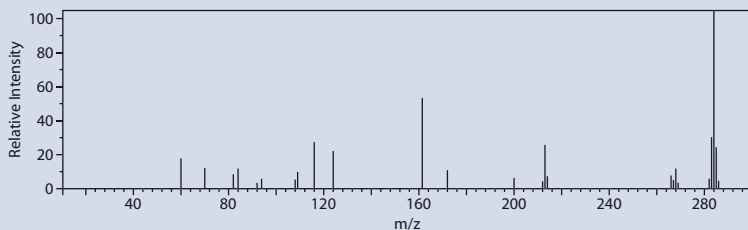


Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

**What is your sample? (Compare with the spectra in Appendix D at the end of this book.)**

**14. You get a violet-blue color with Marquis reagent on a sample submitted to your lab. You ask the mass spectroscopy lab to check it for you, and they**

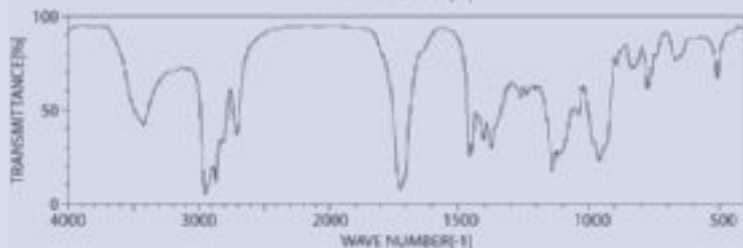
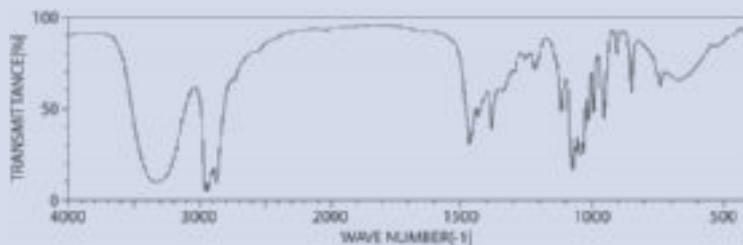
submit the following mass spectra (there were two components):



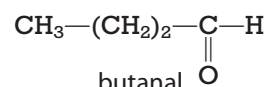
Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

What is your sample? (Compare with the spectra in Appendix E.)

15. Explain the basis of GC-MS to a jury (generally considered to be, on the average, at seventh- to ninth-grade level), and how ethyl alcohol can be differentiated from propyl alcohol.
16. Infrared spectra of butanol (butyl alcohol) and butanal (butyraldehyde) are given below. Which do you think is the aldehyde? Why?



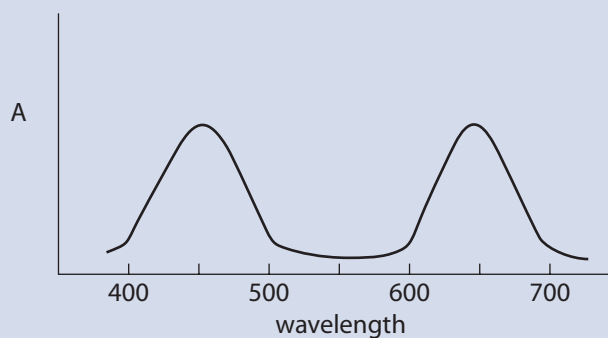
butanol



butanal

Source: SDBSWeb: <http://riodb01.ibase.aist.go.jp/sdbs/> (National Institute of Advanced Industrial Science and Technology, 2005)

17. In the “incense” incident described on page 189, about two pounds of the material was retrieved from the car. What would be the maximum sentence if the driver were convicted as a first offense? (See Appendix B.)
18. If, in the incident at the Detroit airport, the film canister contained 50 grams of cocaine and 10 more grams were found in the man’s camera bag, what could be the sentence if he were convicted as a first offense? (See Appendix C.)
19. What color is a solution of a compound that has an absorption spectrum like the one below?



20. If you were asked in court to assess the accuracy of your analytical results, how would you respond?

## Project: Both Sides of the Issue; Legalization of Drugs

Write a paper analyzing the arguments about legalizing drugs in the United States. Billions of dollars have been spent on the “war on drugs.” Seventy-five percent of people in prison are there on drug or drug-related crime charges. Do we need more severe drug laws? Would legalization benefit society? Which drugs should be legal? How would the cost of drugs be affected? How would the cost of drugs affect violent crimes? Should the state protect people from harm to themselves? What has history taught us about government control? Can the success or failure of drug laws in other countries help us decide?

So you can gain an understanding of both sides of the issue and get experience in identifying and defending the side of the issue that you disagree with, structure your paper as follows:

**TITLE:** Would the Legalization of Drugs Benefit American Society?

**AUTHOR:** Your name

**INTRODUCTION:** Write one or two paragraphs briefly explaining what specific laws would be changed and the controversy surrounding the issue.

**PRO SIDE:** Write one sentence saying that legalization of drugs in the United States would benefit society.

**SUPPORT:** Write a short statement explaining why society would benefit. Write at least three paragraphs supporting the statement, using at least three different sources.

**CON SIDE:** Write one sentence saying that legalization of drugs in the United States would not benefit society.

**SUPPORT:** Write a short statement explaining why legalization of drugs in the United States would not benefit society. Write at least three paragraphs supporting the statement, using three different sources.

**PERSONAL OPINION:** Write your views and conclusions based on the arguments you have already presented. You must support one side or the other.

**WORKS CITED:** List references for all sources that you have used.

## Schools add tobacco to list for student drug screening

### Critics say tests are an invasion of kids' privacy

Some schools are administering urine tests to teenagers to determine if they have been using tobacco. Random drug testing of students involved in extracurricular activities has been affirmed by the U.S. Supreme Court, but opponents say such tests violate the students' rights.

This article appeared in a local newspaper in October 2002. Do you think it is an invasion of individual privacy to be subjected to random drug testing in school? How about for participating in after-school events? What about for attending after-school events? Do you think that tobacco products should be included in the drug-testing programs? Is tobacco a drug? What constitutional amendment pertains to individual privacy?

## Additional Projects

1. How is IR absorption related to the greenhouse effect?
2. Research where each area of the EMS originates and how that frequency of radiation can be used by humankind. For example: Ultraviolet light originates from the sun. It causes sunburn and skin cancer. A black light can cause fluorescence in many substances and is useful in forensic investigation.
3. Research the history of opium—its use and the regulations governing it. The history of other naturally occurring drugs such as marijuana, mescaline (from peyote cactus), and psilocybin mushrooms is also interesting.
4. What incident brought the designer drug fentanyl into the news in 2002? What went wrong?
5. Is the illicit drug trade the fault of the supplier or the user? Support your opinion.
6. Analyze different brands of aspirin tablets to determine (a) the amount of acetylsalicylic acid in each and (b) the type of filler. (See “The Chemistry of Over-the-Counter Drugs,” Flinn Scientific, [www.flinnsci.com](http://www.flinnsci.com).)
7. What is meant by the “doping” of racehorses? Is it a problem?



# References

## Books and Articles

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- Cormier, R. A., W. B. Hudson, and J. A. Siegel. "Thin-Layer Chromatographic Separation of Common Analgesics—A Consumer Experiment," *J. Chem. Ed.* Vol. 56 (March 1979), p. 180.
- Elder, J. W. "Thin-Layer Chromatography of Analgesics—An Update," *J. Chem. Ed.* Vol. 72 (November 1995), p. 1049.
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- Meloan, C. E., R. E. James, and J. R. Saferstein. *Criminalistics: An Introduction to Forensic Science, Lab Manual* (6th ed.). Upper Saddle River, NJ: Prentice Hall, 1998.
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- Ragle, Larry. *Crime Scene*. New York: Avon Books, 2002.
- "Thin-Layer Chromatography," Cat. No. AP9095. Batavia, IL: Flinn Scientific, 1996.

## Websites

- [www.drugabuse.gov/NIDAHome.html](http://www.drugabuse.gov/NIDAHome.html); source of information on drugs
- [www.usdoj.gov/dea/concern/concern.htm](http://www.usdoj.gov/dea/concern/concern.htm); information on many drugs
- [www.ojp.usdoj.gov/bjs/dcf/contents.htm](http://www.ojp.usdoj.gov/bjs/dcf/contents.htm); drugs and crime statistics
- [www.drugabuse.gov/DrugPages](http://www.drugabuse.gov/DrugPages); source of information on drugs
- [www.whitehousedrugpolicy.gov/streetterms/Default.asp](http://www.whitehousedrugpolicy.gov/streetterms/Default.asp); index of street names of drugs
- <http://science.csustan.edu/tutorial/ir/index.htm>; good college-level tutorial with problems on IR spectroscopy
- [www.chem.arizona.edu/massspec/intro\\_html/intro.html](http://www.chem.arizona.edu/massspec/intro_html/intro.html); a good tutorial on mass spectrometry
- [www.chemistry.nmsu.edu/Instrumentation/GC.html](http://www.chemistry.nmsu.edu/Instrumentation/GC.html); explanation of gas chromatography
- [www.chem.vt.edu/chem-ed/sep/gc/gc.html](http://www.chem.vt.edu/chem-ed/sep/gc/gc.html); good tutorial on gas chromatography
- [webbook.nist.gov/chemistry/name-ser.html](http://webbook.nist.gov/chemistry/name-ser.html); database for structures, spectra
- [http://riodb01.ibase.aist.go.jp/sdbs/cgi-bin/cre\\_index.cgi?lang=eng](http://riodb01.ibase.aist.go.jp/sdbs/cgi-bin/cre_index.cgi?lang=eng); spectral database for organic compounds
- [www.aist.go.jp/RIODB/SDBS/cgi-bin/cre\\_index.cgi](http://www.aist.go.jp/RIODB/SDBS/cgi-bin/cre_index.cgi); database for structures, spectra
- [www.rxlist.com/drugs/alpha\\_a.htm](http://www.rxlist.com/drugs/alpha_a.htm); index of drugs
- [www.rxlist.com/cgi/rxlist.cgi](http://www.rxlist.com/cgi/rxlist.cgi); prescription drug identification