Name _____ Block _____

Balancing Chemical Equations – the Law of Conservation of Mass (Matter)

1. _____HCl + _____$NaOH$ → _____H_2O + _____$NaCl$

2. _____Fe + _____HCl → _____$FeCl_2$ + _____H_2

3. _____N_2 + _____O_2 → _____N_2O_5

4. _____H_2CO_3 → _____H_2O + _____CO_2

5. _____CuO + _____H_2SO_4 → _____$CuSO_4$ + _____H_2O

6. _____H_2 + _____O_2 → _____H_2O

7. _____CH_4 + _____O_2 → _____CO_2 + _____H_2O

8. _____Ca + _____H_2O → _____$Ca(OH)_2$+ _____H_2

9. _____HCl + _____$Pb(NO_3)_2$ → _____$PbCl_2$ + _____HNO_3

10. _____HCl + _____$NaOH$ → _____H_2O + _____$NaCl$

11. _____NH_4OH + _____$HC_2H_3O_2$ → _____$NH_4C_2H_3O_2$ + _____H_2O

12. _____Fe_2CO_3 + _____C → _____Fe + _____CO_2

13. _____SO_2 + _____O_2 → _____SO_3

14. _____$(NH_4)_2CO_3$ + _____$FeBr_3$ → _____$Fe_2(CO_3)_3$ + _____NH_4Br

15. _____PbO_2 → _____PbO + _____O_2