Introduction to Atoms

What is the smallest particle into which an element can be divided and still be the same substance?

- A. electron
- B. neutron
- C. proton
- D. atom

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How did the Greek Democritus describe atoms?

- A. large, soft particles
- B. small, hard particles that were dividable.
- C. small, hard particles-like a tiny marble
- D. large, hard particles

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If an isotope of uranium, uranium-235, has 92 protons, how many protons does uranium-238 have?

- **A. 92**
- **B. 95**
- **C. 143**
- **D. 146**

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Which of the following is the smallest subatomic particle and has the least mass in the atom??

- A. nucleus
- B. proton
- C. neutron
- D. electron

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An atom of gold with 79 protons, 79 electrons, and 118 neutrons would have a mass number of

- A. 39
- B. 158
- C. 197
- D. 276

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In Thomson's "plum pudding" model of the atom, the plums represent

- A. atoms
- B. protons
- C. neutrons
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What did Democritus, Dalton, Thomson, Rutherford, and Bohr all have in common?

- A. They each identified new elements.
- B. They each identified new isotopes of atoms.
- C. They each discovered something about what we know about atoms today.
- D. They each were born in Greece..

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The sum of protons and neutrons in an atom

- A. Atomic number
- B. Mass number
- C. Atomic mass unit (amu)
- D. Electron cloud

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- B. Atomic number
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Particle that cannot be cut according to Democritus.

- A. atom
- B. nucleus
- C. isotope
- D. proton

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Region where electrons are likely to be found

- A. electron cloud
- B. nucleus
- C. isotope
- D. proton

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Atoms that have the same number of protons but different numbers of neutrons

- A. proton
- B. electron
- C. isotope
- D. neutron

Atoms that have the same number of protons but different numbers of neutrons

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- B. electron
- C. isotope
- D. neutron

Subatomic particle that has a positive charge

- A. neutron
- B. electron
- C. isotope
- D. proton

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Negatively charged particle discovered by J J Thomson

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- B. electron
- C. proton
- D. atom

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- C. proton
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Particle in the center of an atom that has no charge

- A. neutron
- B. nucleus
- C. isotope
- D. electron

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A unit of mass that describes the mass of an atom or molecule.

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Central region of the atom

- A. isotope
- B. proton
- C. nucleus
- D. electron

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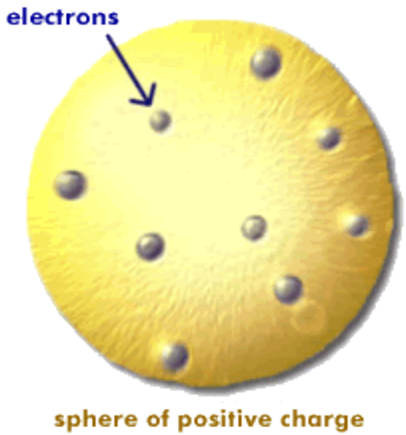
Who proposed this model of an atom?



A. Bohr
B. Thomson
C. Rutherford
D. Democritus

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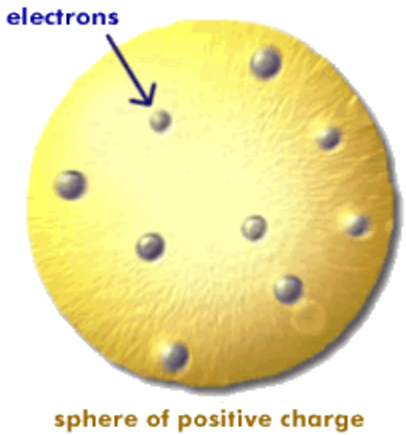

A. Bohr
B. Thomson
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Nice to have met you…