

Look

@ the Atom the $\ominus$ Charge is on for the conjugate base

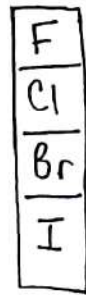
A

Same period (Row):



increasing stability (and electronegativity)

Same Group (column)



increasing stability
*larger atoms have better stability

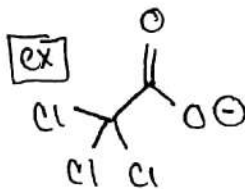
Resonance possibilities

- the more delocalized the better
- one Oxygen is better than many Carbon

R

Induction - "pulling" of e^- density by an electroneg. atom

*this helps stabilize the neg. charge



↑ good placement of Cl (in Goldilocks region)

Alpha effect: destabilizing $\cap$
a L.P. on the atom next to neg. charge

Alkyl groups (Carbon groups) destabilizing $\cap$
b/c they are e^- donating

Orbitals

Single bonds
 sp^3

double bonds
 sp^2

Triple Bonds
 sp

e^- 's

nucleus



increasing stability

exception:

$\equiv \ominus$ is more stable than $N \ominus$