

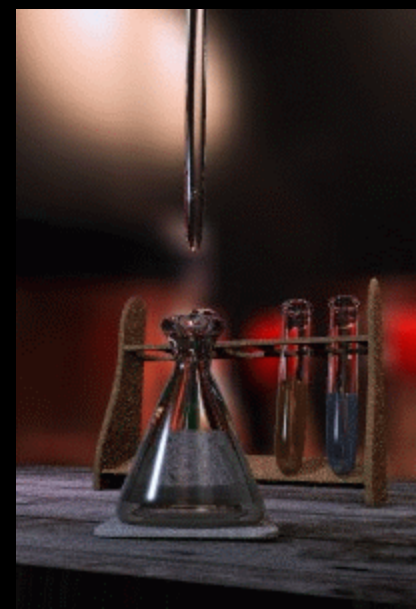
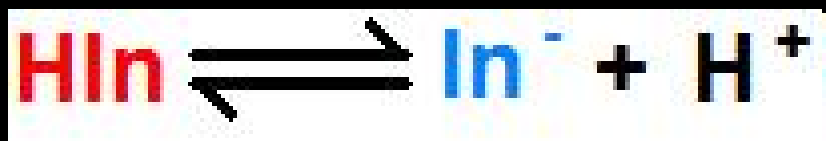
ACIDS AND BASES

Determining the
Acidity of a Solution



Measuring pH

- The traditional way of determining the pH of a solution is by using indicators → substances that exhibit different colors in acidic and basic solutions.
- We can represent a generic indicator as a weak acid HIn.
- Let's assume that for a particular indicator HIn is red and the anion In⁻ is blue.



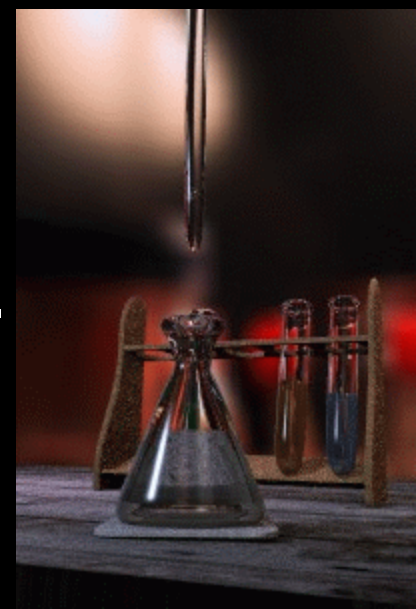
Measuring pH



- Thus, in a solution containing lots of H^+ , the indicator will be red (the HIn form).



- In a basic solution, the H^+ will be removed from HIn to form the blue In^- ion and the solution will be blue.



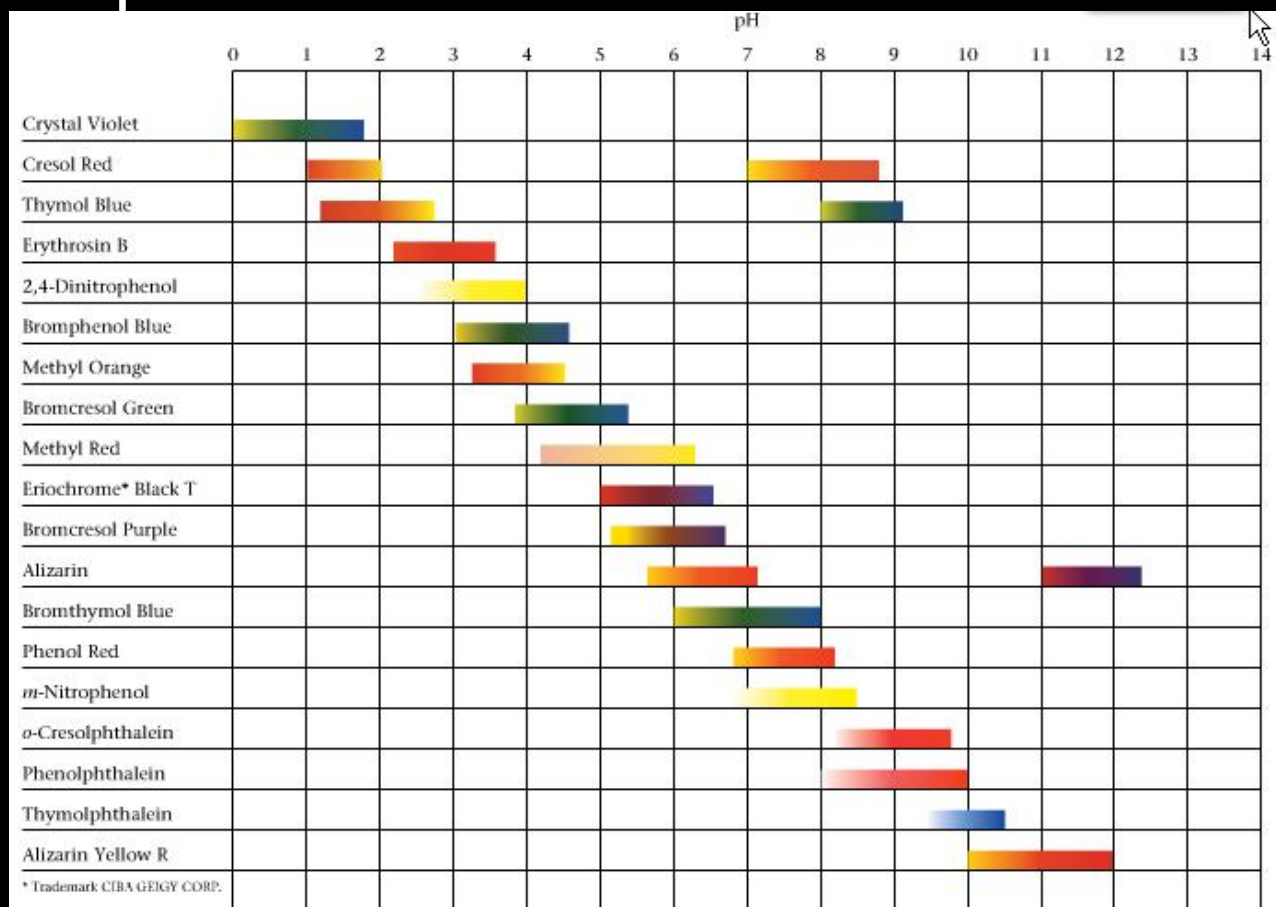
Measuring pH

- In intermediate pH regions, significant amounts of both HIn and In^- will be present and the solution will be some shade of purple (red plus blue).
- Many substances exist that turn different color when H^+ is present (HIn) than when H^+ is absent (In^-).
 - Some of these indicators are shown in the picture, along with the colors of their HIn and In^- forms.



Measuring pH

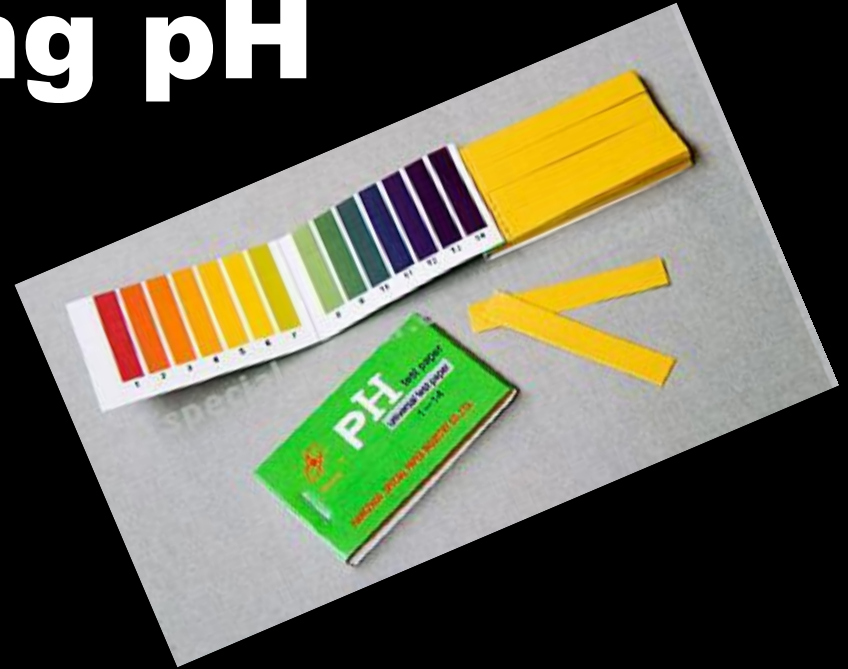
- Different indicators change at different pH values depending on the acid strength of HIn for a particular indicator.



The pH ranges shown are approximate. Specific transition ranges depend on the indicator solvent chosen.

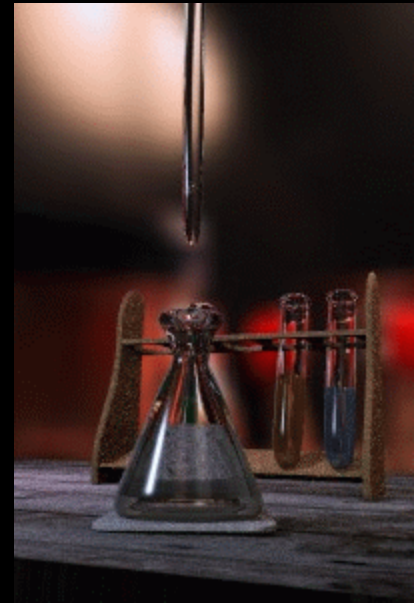
Measuring pH

- A convenient way to measure the approximate pH of a solution is by using indicator paper.
- Indicator paper is a strip of paper coated with a combination of indicators.
- Indicator paper turns a specific color for each pH value.



Measuring pH

- The pH value of a solution can be measured electronically using a pH meter.
- A pH meter contains a probe that is very sensitive to the $[H^+]$ in a solution.
 - When the probe is inserted into a solution, the $[H^+]$ in the solution produces a voltage that appears as a pH reading on the meter.
 - A pH meter can determine the pH value of a solution quickly and accurately.



- The End

