

$$\text{Equation: } \text{pH} + \text{pOH} = 14$$

$$14 - \text{pH} =$$

$$= 14 - \text{pOH}$$

pH

pOH

$$\text{Molarity} = \frac{\text{Mole}}{\text{Liter}}$$

*Strong Acid or Base

= Acid

Base =

$$-\log [\text{OH}^-] =$$

$$= 10^{-\text{pOH}} \quad \text{Second log then -pOH}$$

$$\text{Equation: } \text{pOH} = -\log [\text{OH}^-]$$

$$\text{Equation: } \text{pH} = -\log [\text{H}^+]$$

$$-\log [\text{H}^+] =$$

$$= 10^{-\text{pH}} \quad \text{Second log then -pH}$$

H⁺

OH⁻

$$1e^{-14} / [\text{H}^+] =$$

$$= 1e^{-14} / [\text{OH}^-]$$

$$\text{Equation: } [\text{H}^+] [\text{OH}^-] = 1e^{-14}$$