Name\_\_\_\_\_

A <u>chemical bond</u> - forms	when 2 or more atoms	rearrange
	to increase _	······································
i <u>onic bond</u> - forms when from one ato	valence om to another	are
cation – atom	electrons to b	ecome charged
anion - atom	electrons to be	come charged
In ionic compounds	the ions are arranged i	in a
forc	es hold the ions togeth	er.
Properties of ionic	compounds:	
•	and	points
•	not easily	
•	electricity when _	or
		s are free to
covalent bond	are	, forming
Covalent compounds t	s have ogether.	forces holding the
Properties of coval	ent compounds:	
<ul> <li>lower</li> </ul>	and	points
<ul> <li>Many cove</li> </ul>	alent compounds are	liquids or gases.
•	easier to	
<ul> <li>are not</li> </ul>	of	electricity
	rty that tells how stron	ng an atom's
		gativity than hydrogen, oxygen
nolds onto shared electro negative charge and the h		ring the oxygen a
regurive churge and the r		-
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polar covalent bonds:

 electrons are shared \_\_\_\_\_\_, creating partially charged ends or \_\_\_\_\_.

nonpolar covalent bonds:

electrons are shared \_\_\_\_\_\_ because atoms have the same electronegativities

Electonegativity difference:	Type of Bond
greater than or equal to 1.7	
between 1.7 and 0.3	
less than or equal to 0.3	
Examples: Mg and F?	S and O?

Program 501, problem set 1:

metallic bond -	electrons are (creates a " o				
properties of metals:					
1.					
2.					
3.					
4.					
The Chemistry Quiz					
CR1	CR2	1	2		
	3	4	5		
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