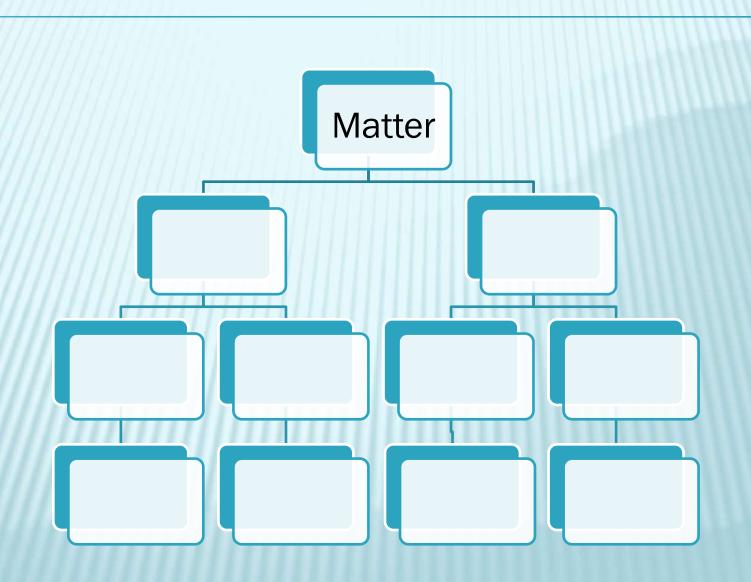
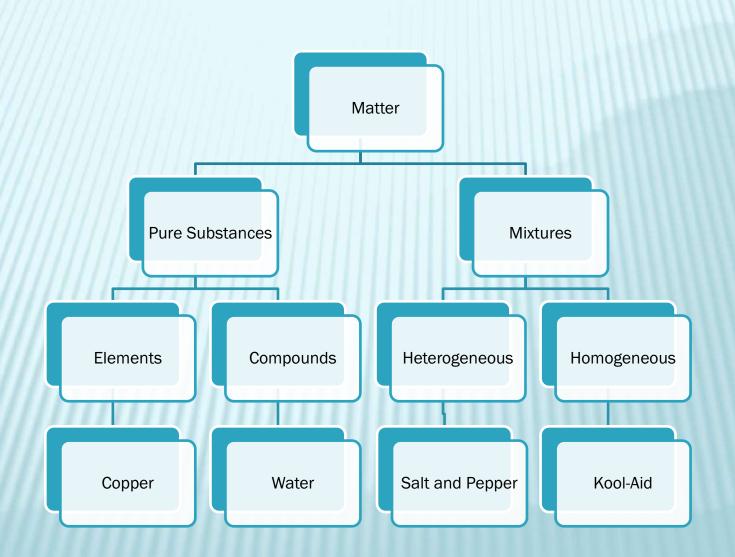
Chapter 5

ELEMENTS, COMPOUNDS, AND MIXTURES





PHYSICAL CHANGE

Does not make a different substance

Ex:

- Melting
- Freezing
- Boiling/evaporation
- **X** Condensation
- **X** Sublimation

- **X** Dissolving
- **×** Bending
- **X** Crushing
- Breaking
- Chopping
- **×** Filtration
- **x** distillation

CHEMICAL CHANGE

Produces new substances

Examples of Chemical Change		
Chemical Change	Description	Example
Combustion	Rapid combination of a fuel with oxygen; produces heat, light, and new substances	Gas, oil, or coal burning in a furnace
Electrolysis	Use of electricity to break a compound into elements or simpler compounds	Breaking down water into hydrogen and oxygen
Oxidation	Slow combination of a substance with oxygen	Rusting of an iron fence
Tarnishing	Slow combination of a bright metal with sulfur or another substance, producing a dark coating on the metal	Tarnishing of brass

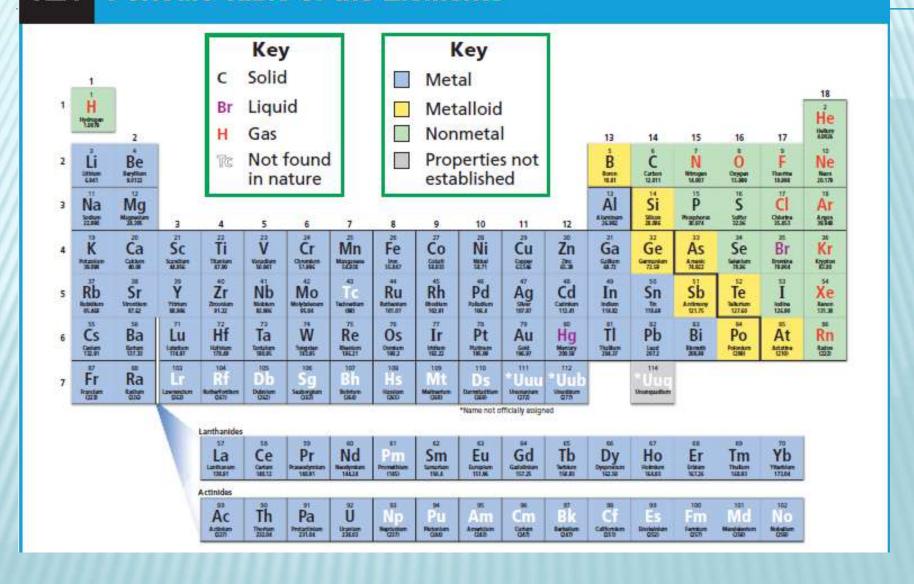
PURE SUBSTANCES

- *A single type of matter with a specific composition and a specific set of properties
- Includes elements and compounds

ELEMENTS

- **X**Pure substance
- **X**Simplest substances
- Cannot be broken down into simpler substances by physical or chemical means
- Made up of only one type of atom
- Have unique physical and chemical properties
- Examples: gold, silver, carbon, helium, calcium, etc. (over 100)

Periodic Table of the Elements

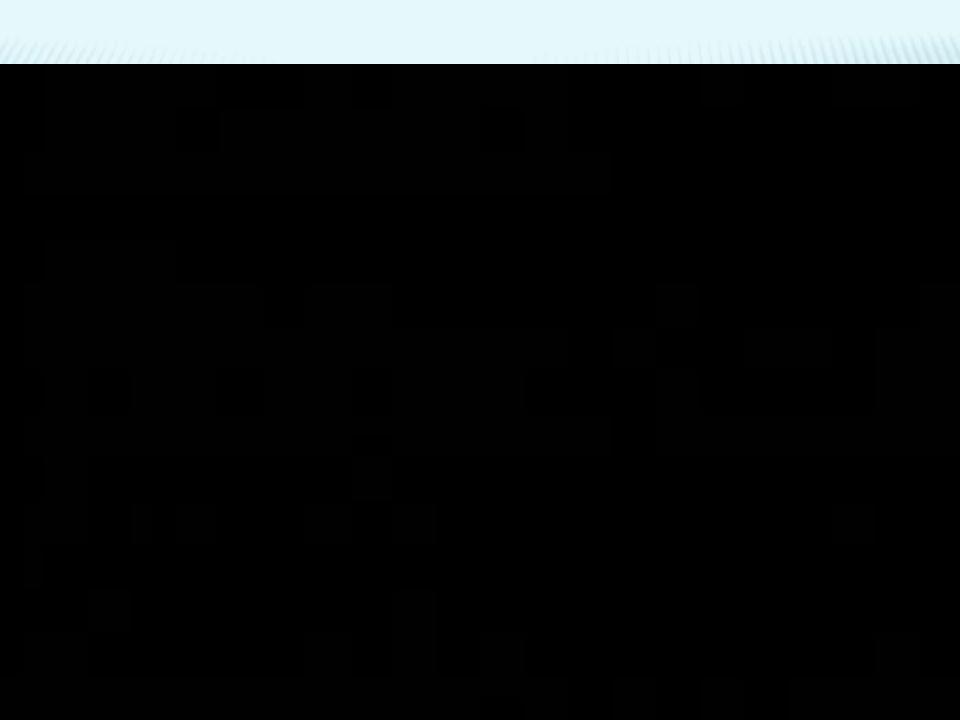


COMPOUNDS

- ➤ Pure substance
- Made up of two or more elements that are chemically combined
- Can be broken down chemically but not physically
- Have own set of physical properties that may be very different from their original parts.
- Combine in definite ratios
- **X** Examples: H₂O, NaCl, CO₂, C₆H₁₂O₆

CHEMICAL FORMULAS

- \times H₂0 = water
- ➤ NaCl = table salt
- \times CO₂ = carbon dioxide
- \times C₆H₁₂O₆ = sugar (glucose)
- ****the number of atoms for each element is determined by the number beside it. For example, water has 2 hydrogens and one oxygen.



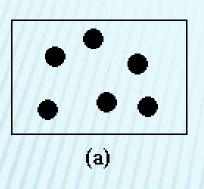
MIXTURES

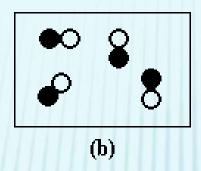
- Two or more substances that are <u>not</u> chemically combined
- Substances in a mixture keep their own individual properties
- × Parts of a mixture are not in set ratios
- Can be physically separated
- Examples: Salad, soil, Kool-Aid, salt water, air, brass, salt and pepper

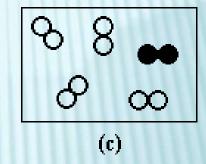
SEPARATION TECHNIQUES

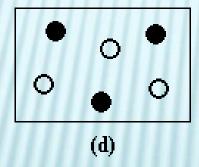
- Melting
- Boiling/Distilling
- **X** Magnets
- **×** Filtration
- **×** Evaporation
- Chromatography

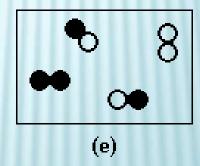
IDENTIFY THE FOLLOWING AS A ELEMENT, COMPOUND, OR MIXTURE











TYPES OF MIXTURES

- ★1. Heterogeneous
- **×**2. Homogeneous

HETEROGENEOUS MIXTURES

- Mixtures that are "different" throughout
- Not evenly mixed; you can see the different parts
- *Examples: salt and pepper, soil, salad

HOMOGENEOUS

- Mixtures that are the "same" throughout
- Evenly mixed; you cannot see the different parts
- **X**Also called solutions
- Examples: salt water, Kool-aid, air, brass

SOLUTIONS

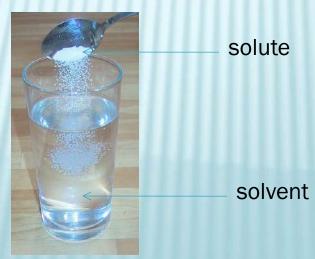
- Mixtures of two or more substances in which one or more of them seem to disappear in the other
- *Another name for a homogeneous mixture
- Can be made up of solids, liquids, or gases.

TYPES OF SOLUTIONS

- XLiquid: Kool-aid, salt water
- **XGas:** Atmosphere
- **Solid:** Brass, stainless steel (alloys)

HOW DO SOLUTIONS FORM?

- By a process in which a substance breaks up into atoms, ions, and molecules.
- Have two parts:
 - 1. Solute: disappears or dissolves
 - 2. Solvent: dissolves the solute



EXAMPLE

- **×** Kool-Aid
- **X**Solute: powder mix
- **X**Solvent: water
- **usually more solvent
- ** because the solvent is usually water it is called the universal solvent
- **solutions in which water is the solvent are called **aqueous** solutions

DETERMINING SOLUBILITY

- **Solubility:** how much of a solute dissolves in a given solvent at a specific temperature
- If a solute can be dissolved it is said to be soluble
- If a solute cannot be dissolved it is said to be insoluble
- **Saturated**: solution that contains all of the solute it can
- **Unsaturated**: solution that <u>does not</u> hold all of the solute it can

FACTORS THAT AFFECT SOLUBILITY

- *1. Temperature: an increase in temperature causes an increase in solubility.
- ***2.** Pressure: an increase in pressure causes an increase in solubility

DETERMINING CONCENTRATION

- Concentration is how much solute there is compared to the amount of solvent
- **Concentrated**: large amount of solute
- **Dilute**: small amount of solute



PRECIPITATE

- New solute that falls out of a solution by chemical means.
- Examples: soap scum, stalactites/stalagmites







