Reynolds School District | **Reynolds Jr.-Sr. High School**
Weekly Lesson Plan/Assignments/Homework | **Week of: December 04, 2017**
Teacher: Mrs. Coburn **Subject: Chemistry 1 (8)** **Grade: 10 - 12**

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Monday 12/04

Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** **A** **B** **C** **D**

Lesson Objective: **List** the six basic steps used in writing Lewis structures.

- **Explain** how to determine Lewis structures for molecules containing single bonds, multiple bonds, or both.

Details: -review WS 6.3
-Quiz 6.3 and diagram ionic compounds
-Begin notes on covalent bonding (6.2 pp. 178-183) and naming molecular compounds
-complete Covalent bonding WS in packet **Homework**: Adopt an Element W 12/06; Molecular Model Project research and brochure due date 12/20 W

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Tuesday 12/05

Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** **A** **B** **C** **D**

Lesson Objective:

- **Define** ionic bonding and ions molecule and molecular formula.
- **List** the six basic steps used in writing Lewis structures.
- **Explain** how to determine Lewis structures for molecules containing single bonds, multiple bonds, or both.

Details: -continue notes 6.2 (e- dot diagrams and Lewis structures p184-185)
-complete e- dot digrams WS & Lewis Structures WS in packet p. 32
Homework: Adopt an Element W 12/06; Molecular compounds ch 6-7 WS & More Fun with Lewis Structures; work Molecular Model Project research and brochure due date 12/20 W

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Wednesday 12/06

Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** **A** **B** **C** **D**

Lesson Objective:

- Describe the electron-sea model of metallic bonding, and explain why metals are good electrical conductors.
- Explain why metal surfaces are shiny.
- Explain why metals are malleable and ductile but ionic-crystalline compounds are not.
- Explain VSEPR theory.
- Predict the shapes of molecules or polyatomic ions using VSEPR theory.

Details: -turn in Adopt an Element Project
-check and review HW
-Quiz 6.4 & drawing Lewis structures (Ionic and covalent)
-Notes on VSEPR Shapes
-VSEPR Worksheet TRIO
Homework: VSEPR WS in packet; Molecular Model Project research and brochure due date 12/20 W

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Thursday 12/07

Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** **A** **B** **C** **D**

Lesson Objective: see Wednesday's objectives

- Describe dipole-dipole forces, hydrogen bonding, induced dipoles, and London dispersion forces and their effects on properties such as boiling and melting points.
- Explain what determines molecular polarity.

Details:-Check and review VSEPR
-Notes on 6.4 Intermolecular Forces
-<u>Lab</u>- VSEPR Shapes with marshmallows
Homework: complete practice IMF Ch 6 WS; Molecular Model Project research and brochure due date 12/20 W

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Friday 12/08

Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** **A** **B** **C** **D**

Lesson Objective: see Friday's objectives

- Describe dipole-dipole forces, hydrogen bonding, induced dipoles, and London dispersion forces and their effects on properties such as boiling and melting points.

- Explain what determines molecular polarity.

Details:-Check and review IMF WS
-Notes on 6.4 Intermolecular Forces
-Lab- IMFs **Homework**: Chapter 6 Test Review; Molecular Model Project research and brochure due date 12/20 W

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Monday **12/11**
Academic Standard # 3.4.10.A & B 3.4.12.A & B **Rigor/Relevance Quadrant** (A) (B) C D
Lesson Objective: **Review** all chapter 6 objectives
Details: Turn in Labs
-review Chapter 6 Test Review
Homework: Study for Chapter 6 TEST; Molecular Model Project research and brochure due date 12/20 W

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