

Name: _____ Date: _____ Period: _____

Using Percent Composition

1. You have a 236.10 g sample of sodium nitride what is the mass of sodium in the sample? (Hint find % composition of sodium first.)
2. There is a 769.28 g sample of barium phosphate what is the mass of
HINT: Write chemical formula for Barium Phosphate, and then find % composition of element
 - a. Barium in the sample
 - b. Phosphorous in the sample

Empirical Formula

3. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen?
4. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 %
5. What is the empirical serine if it's percent composition is 34.95 % C, 6.844 % H, 46.56 % O and 13.59 % N?
6. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms and 39.12 g of chlorine atoms were produced. What is the empirical formula of the compound?

