Name:	Date:	Period:
1.	<b>Using Percent Composition</b> You have a 236.10 g sample of sodium nitride what is the mass of sodium in the sample? (Hint find % composition of sodium first.)	
2.	There is a 769.28 g sample of barium phosphate what is the mass of HINT: Write chemical formula for Barium Phosphate, and then find % com a. Barium in the sample	position of element
	b. Phosphorous in the sample	
3.	Empirical Formula What's the empirical formula of a molecule containing 65.5% carbon, 5.5 oxygen?	% hydrogen, and 29.0%
4.	The compound benzamide has the following percent composition. What is $C = 69.40 \ \% \ H = 5.825 \ \% \ O = 13.21 \ \% \ N = 11.57 \ \%$	s the empirical formula?
5.	What is the empirical serine if it's percent composition is $34.95 \% C$ , $6.8 \cdot 13.59 \% N$ ?	44 % H, 46.56 % O and
6.	A 50.51 g sample of a compound made from phosphorus and chlorine is of products showed that 11.39 g of phosphorus atoms and 39.12 g of chlorin What is the empirical formula of the compound?	

## Molecular Formula and Empirical formula empirical formula of C<sub>2</sub>OH<sub>4</sub> and a molar mass of 88

1.	A compound with an empirical formula of C <sub>2</sub> OH <sub>4</sub> and a molar mass of 88 grams per mole.
2.	A compound with an empirical formula of C <sub>4</sub> H <sub>4</sub> O and a molar mass of 136 grams per mole.
3.	A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole.
4.	A compound whose percent composition by mass is 85.7% carbon and 14.3% hydrogen a. What is the empirical formula of the compound?
	b. If the compound has a molar mass of 56 g/mol, what is the molecular formula?
5.	A compound contains 64.3% carbon, 7.14% hydrogen and 28.6% oxygen. The molecular formula has a molecular mass of 448 g/mol.  a. What is the empirical formula for this compound?
	b. What is the molecular formula for this compound?