

Thermal Energy

A. Temperature & Heat

1. Temperature is related to the average kinetic energy of the particles in a substance.

The atoms in an object are in constant motion.



A When the horseshoe is hot, the particles in it move very quickly.

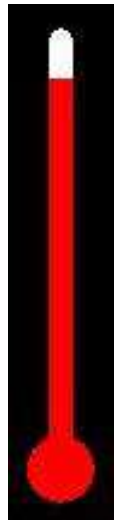


B When the horseshoe has cooled, its particles are moving more slowly.

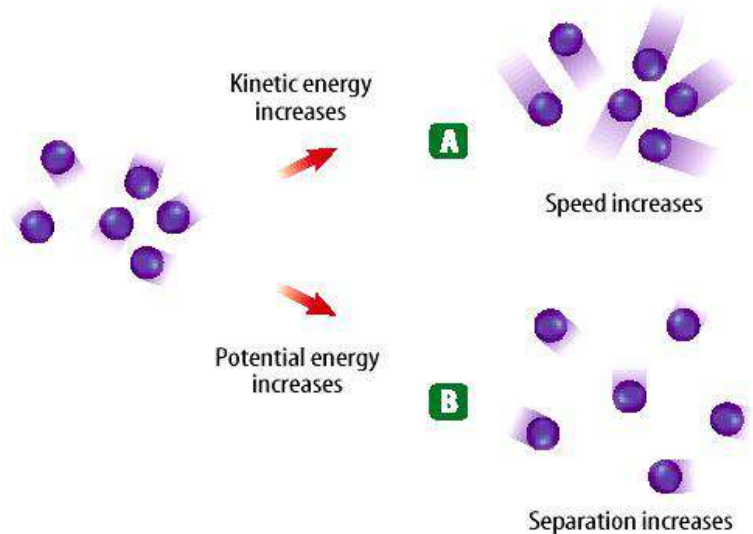
2. SI unit for temp. is the **Kelvin**

a. $K = C + 273$ ($10^{\circ}\text{C} = 283\text{K}$)

b. $C = K - 273$ ($10\text{K} = -263^{\circ}\text{C}$)



3. Thermal Energy – the total of **all** the kinetic and potential energy of all the particles in a substance.



4. Thermal energy relationships

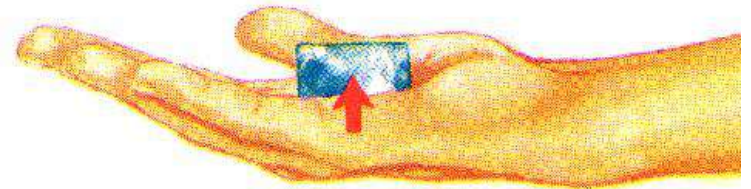
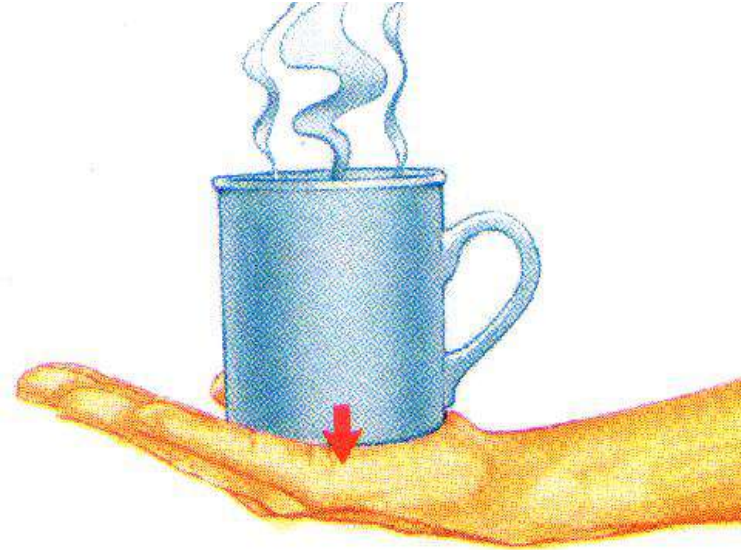
- a. As temperature increases, so does thermal energy (because the kinetic energy of the particles increased).
- b. Even if the temperature doesn't change, the thermal energy in a more massive substance is higher (because it is a total measure of energy).

5. Heat

a. The *flow* of thermal energy from one object to another.

b. Heat *always* flows from **warmer** to **cooler** objects.

Cup gets cooler while hand gets warmer



Ice gets warmer while hand gets cooler

6. Specific Heat

a. Some things **heat up** or **cool down** faster than others.



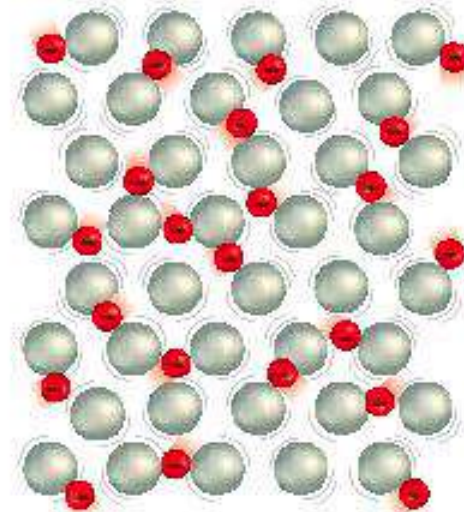
b. Specific heat is the amount of heat required to raise the temperature of 1 kg of a material by one degree (C or K).

$$1) \ C_{\text{water}} = 4184 \text{ J / kg C}$$

$$2) \ C_{\text{sand}} = 664 \text{ J / kg C}$$

This is why land heats up quickly during the day and cools quickly at night and why water takes longer.

Why does water have such a high specific heat?



metal

Water molecules form strong bonds with each other; therefore it takes more heat energy to break them. Metals have weak bonds and do not need as much energy to break them.

How to calculate changes in thermal energy

$$Q = m \times \Delta T \times C_p$$

Q = change in thermal energy

m = mass of substance

ΔT = change in temperature ($T_f - T_i$)

C_p = specific heat of substance

c. A calorimeter is used to help measure the specific heat of a substance.

Knowing its Q value,
its mass, and its
 ΔT , its C_p can be
calculated

