

Report for Experiment 15

Forming and Naming Ionic Compounds

Prelaboratory Questions

- For the following pairs of ions, write the formula of the compound that you would expect them to form:
 - barium and hydroxide
 - cobalt (III) and phosphate
 - iron (II) and sulfate
 - silver and hydrogen carbonate
- Platinum is a transition metal and forms Pt^{2+} and Pt^{4+} ions. Write the formulas for the compounds for each of these ions with
 - bromide ions
 - carbonate ions

Data/Observations

	Na_3PO_4 PO_4^{3-}	KOH OH^-	$\text{K}_4\text{Fe}(\text{CN})_6$ $\text{Fe}(\text{CN})_6^{4-}$	Na_2CO_3 CO_3^{2-}	KI I^-	$\text{Na}_2\text{C}_2\text{O}_4$ $\text{C}_2\text{O}_4^{2-}$
CoCl_2 Co^{2+}						
$\text{Pb}(\text{NO}_3)_2$ Pb^{2+}						
CuSO_4 Cu^{2+}						
$\text{Fe}(\text{NO}_3)_3$ Fe^{3+}						
NiCl_2 Ni^{2+}						
SrCl_2 Sr^{2+}						
AgNO_3 Ag^+						

Analysis and Conclusions

1. For each case in which you found a reaction occurred, write the correct formula for the substance formed.

	Na_3PO_4 PO_4^{3-}	KOH OH^-	$\text{K}_4\text{Fe}(\text{CN})_6$ $\text{Fe}(\text{CN})_6^{4-}$	Na_2CO_3 CO_3^{2-}	KI I^-	$\text{Na}_2\text{C}_2\text{O}_4$ $\text{C}_2\text{O}_4^{2-}$
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NiCl_2 Ni^{2+}						
SrCl_2 Sr^{2+}						
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