

Pure Substances, Mixtures

- Pure substance: matter that has a fixed (constant) composition and unique properties. Contains only 1 type element or compound; homogeneous
- Mixture: Contains at least 2 PHYSICALLY combined compounds; can be homogeneous or heterogeneous

### Homogeneous Substances

- Means same throughout
- 1) element: only 1 type of atom
- 2) compound: 2 or more CHEMICALLY combined elements (not easily separated from each other)
   ex: water, CO<sub>2</sub>
- Solution: a special kind of mixture 2 phases/parts (SOLUTE dissolves & SOLVENT does the dissolving)

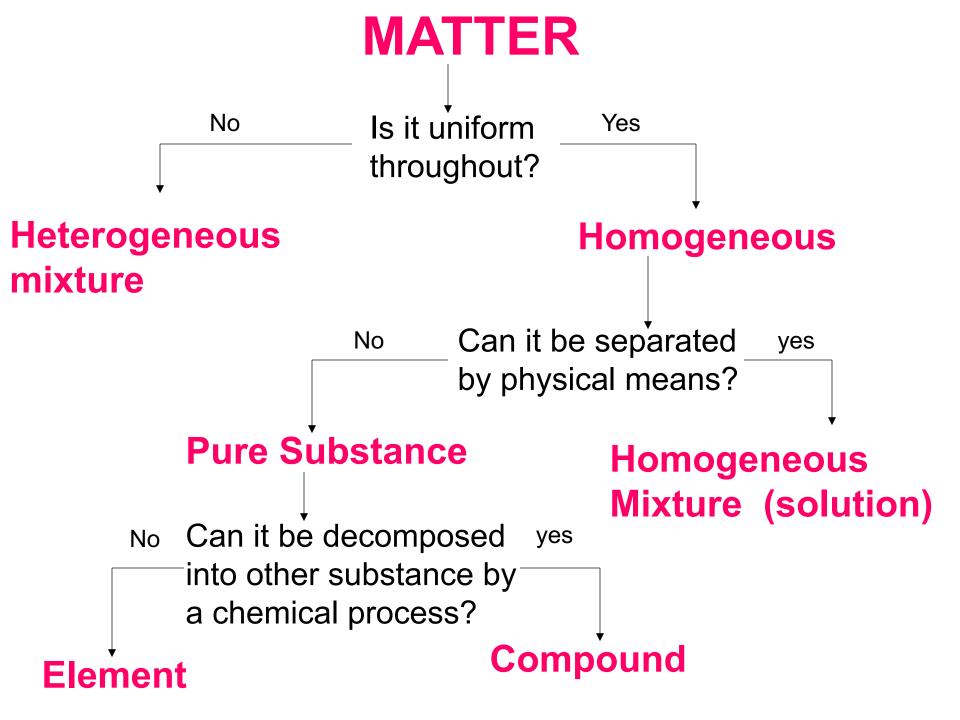
ex: moist air (H<sub>2</sub>O in Air); sterling silver (Cu in Ag...called an alloy)

# Heterogeneous matter

- Means different throughout
- Always a MIXTURE (solutions are mixtures that are NOT heterogeneous)
  - 2 or more PHYSICALLY combined substances (elements/compounds)
  - ex: air, muddy water

#### Classification of Matter

- Matter can also be classified according to its composition. Mixtures can be homogeneous or heterogeneous.
- Mixtures can be separated into <u>pure</u>
  <u>substances</u>, and pure substances can be
  either compounds or elements.



## Homogeneous mixtures

Homogeneous mixtures: is a mixture in which the components are evenly distributed among each other. You can't see the component parts. It has a constant composition throughout. Homogeneous mixtures are also

**Examples:** Salt dissolved in water, sugar dissolved in water, apple juice, tea, copper (II) sulfate solution in water, alloys....

called **SOLUTIONS** 

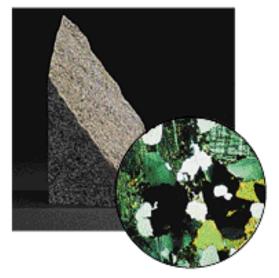


C Copper(II) sulfate (CuSO<sub>4</sub>) in water, a homogeneous mixture (solution)

# heterogeneous mixtures

Heterogeneous mixture: the components are not evenly distributed among each other. An heterogeneous mixture has two or more distinct phases that are usually detectable. This type of mixture does NOT have uniform properties.

**Examples**: Sand water, oil and water, sulfur and iron, granite,...



A Granite, a heterogeneous mixture