

# Percent Composition Problems

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1) A sample of a compound analyzed consists of 5.34 g carbon, 0.42 g hydrogen, and 47.08 g chlorine. What is the % Cl in this sample? (Round answer to the nearest tenth of a percent) \*

8 points

Your answer

2) Find the % tin in a compound containing tin and chlorine if 18.35 g of the compound contains 5.74 g tin. (Round answer to the nearest tenth of a percent) \*

8 points

Your answer



3) If 3.907 g of carbon combines completely with 0.874 g of hydrogen to form a compound, what is the percent carbon in this compound? (Round answer to the nearest tenth of a %) \*

8 points

Your answer

4) From the formula for calcium acetate  $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ , calculate the mass of carbon that can be obtained from 65.3 g of the compound. (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

5) How many grams of aluminum are in 25.0 g of  $\text{Al}_2\text{O}_3$ ? (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

6) How many grams of iron are in 21.6 g  $\text{Fe}_2\text{O}_3$ ? (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

7) Find the % C in  $\text{CBr}_4$ . (Round answer to the nearest tenth of a percent) \*

8 points

Your answer

8) Find the %N in  $\text{Ca}(\text{NO}_3)_2$ . (Round answer to the nearest tenth of a percent) \*

8 points

Your answer

9) How many grams of Cl can be extracted from 86.5 g  $\text{NH}_4\text{Cl}$ ? (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

10) How many grams of  $\text{CO}_2$  can be released from 48.5 g  $\text{K}_2\text{CO}_3$  when it reacts with an acid? (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

11) If 135 g of the hydrate  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  are heated, how many grams of water can be removed? (Round answer to the nearest tenth of a gram) \*

8 points

Your answer

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