

Name: _____

Date: _____

Period: _____

Percentage Composition Worksheet

Give the % composition of all elements in these compounds. Show all work!

Remember: % composition of element = $\frac{\text{mass of element}}{\text{mass of compound}} \times 100$

When calculating mass of element don't forget to account for the subscripts.

1. ammonium sulfite $(\text{NH}_4)_2\text{SO}_3$ molar mass of compound:

% N _____

% H _____

% S _____

% O _____

2. aluminum acetate $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$ molar mass of compound:

% Al _____

% C _____

% H _____

% O _____

3. sodium bromide NaBr molar mass of compound:

% Na _____

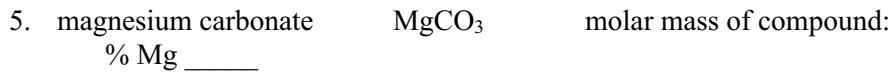
% Br _____

4. copper (II) hydroxide $\text{Cu}(\text{OH})_2$ molar mass of compound:

% Cu _____

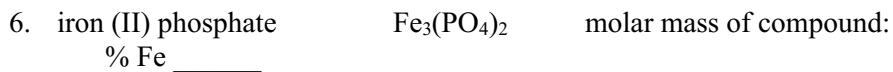
% O _____

% H _____



% C _____

% O _____



% P _____

% O _____

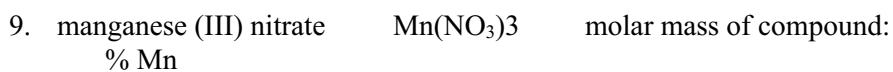


% N _____



% C _____

% N _____



% N _____

% O _____



% P _____

Percentage Composition Worksheet – Answer Key

Give the % composition of all elements in these compounds. **Show all work!**

1) ammonium sulfite % N **24.1**
 % H **6.9**
 % S **27.6**
 % O **41.3**

2) aluminum acetate % Al **13.2**
 % C **35.3**
 % H **4.4**
 % O **47.1**

3) sodium bromide % Na **22.4**
 % Br **77.6**

4) copper (II) hydroxide % Cu **65.1**
 % O **32.8**
 % H **2.1**

5) magnesium carbonate % Mg **28.8**
 % C **14.2**
 % O **56.9**

6) iron (II) phosphate % Fe **46.8**
 % P **17.3**
 % O **35.8**

7) beryllium nitride % Be **49.1**
 % N **50.9**

8) potassium cyanide % K **60.1**
 % C **18.4**
 % N **21.5**

9) manganese (III) nitrate % Mn **22.8**
 % N **17.4**
 % O **59.8**

10) lithium phosphide % Li **40.0**
 % P **60.0**

11) nickel (III) sulfate % Ni **28.9**
 % S **23.7**
 % O **47.3**