

Unit 4 The Periodic Law

Standards

SC4. Students will use the organization of the Periodic Table to predict properties of elements.

- a. Use the Periodic Table to predict periodic trends including atomic radii, ionic radii, ionization energy, and electronegativity of various elements.
- b. Compare and contrast trends in the chemical and physical properties of elements and their placement on the Periodic Table.

Essential Questions

1. What roles did Mosely and Mendeleev play in the development of the periodic table?
2. How is the organization of the periodic table related to electron configurations of the elements?
3. How is the periodic law used to predict physical and chemical properties of elements.

Calendar

September/October 2012				
Monday	Tuesday	Wednesday	Thursday	Friday
				21 Getting to know the periodic table activity
24 Development & arrangement of the periodic table notes	25 Periodic table notes TU Science Café (6:30 PM)	26 Lab—Density as a Periodic Trend	27 Graphing periodic trends	28 REPORTS DUE* Periodic Trends
October 1 Review Unit 4	2 Unit 4 Test	3 Review for benchmark	4 Benchmark	5 Discuss Mole Day projects

*Reports may be turned in by 3:05 PM, Monday, October 1 with no penalty. Reports turned in after that will receive a 10 point penalty. Reports will not be accepted after 3:05 PM, October 5.

Notebook Assignments

1	Syllabus		7		
2	Lab Safety		8		
3	PowerPoint Note Sheet		9		
4			10		
5			11		
6			12		