Date: _____ Period: _____

Name: _____

Molecular Geometry Lab

Complete the following data table

- 1. Draw the Lewis line structure for the covalent molecule
- 2. Locate and identify the central atom (highlight, make with different color). IF there are only two atoms total select one to be your central atom.
- 3. State the number of bonds the central atom is forming
- 4. State the number of lone pairs on the central atom
- 5. Using the modeling kit build the molecule
- 6. State the molecular geometry around the central atom

Covalent Molecule	Lewis Line Structure	# of bonds on central atom	# of lone pairs of electrons on central	Geometry
			atom	
1.Cl ₂				
2.02				
3.P ₂				
4.BH ₃				
5 C.H.				
J.C2116				
0.312				
7.CF ₄				

8.SBr ₂		
9.NCl₃		
10. H ₂ S		
11. PH ₃		
12. BF ₃		
13. NBr ₃		
14. PBr ₂ I		
15. SiBr ₂ H ₂		

Post Lab Questions:

- 1. What is the molecular geometry of all diatomic compounds?
- 2. What is one way to recognize that a compound has bent geometry?
- 3. What is one way to recognize that a compound has trigonal planar geometry?
- 4. What is one way to recognize that a compound has tetrahedral geometry?
- 5. What is one way to recognize that a compound has trigonal pyramid geometry?

