

Chemistry SLO / Exam Review (Answers)

Chapter 2	Chapter 3	Chapter 4	Chapter 5	Chapter 6	Chapter 7	Chapter 8	Chapter 9	Chapter 10	Chapter 11	Chapter 12	Chapter 13
1 I	1 E	1 J	1 H	1 C	1 J	1 I	1 I	1 F	1 G	1 B	1 J
2 D	2 I	2 G	2 C	2 I	2 I	2 E	2 B	2 H	2 H	2 E	2 N
3 H	3 G	3 A	3 F	3 B	3 F	3 J	3 C	3 B	3 I	3 D	3 H
4 A	4 F	4 C	4 E	4 G	4 G	4 F	4 H	4 E	4 F	4 A	4 E
5 J	5 C	5 F	5 D	5 H	5 H	5 D	5 G	5 G	5 C	5 F	5 D
6 B	6 B	6 E	6 G	6 A	6 L	6 A	6 A	6 C	6 A	6 C	6 K
7 F	7 J	7 D	7 B	7 J	7 A	7 B	7 F	7 A	7 D	7 D	7 B
8 E	8 A	8 B	8 J	8 D	8 C	8 C	8 E	8 D	8 J	8 D	8 G
9 C	9 H	9 H	9 A	9 F	9 E	9 G	9 D	9 D	9 B	9 B	9 F
10 G	10 D	10 I	10 I	10 E	10 D	10 H	10 C	10 B	10 E	10 D	10 L
11 A	11 C	11 C	11 C	11 B	11 B		11 A	11 A	11 D	11 D	11 A
12 C	12 D	12 D	12 C	12 D	12 K		12 A	12 C	12 B	12 B	12 M
13 D	13 A	13 C	13 C	13 D	13 M		13 C	13 B	13 C	13 C	13 C
14 C	14 C	14 C	14 B	14 A			14 D	14 B	14 B	14 B	14 I
15 D	15 A	15 D	15 B	15 D			15 D	15 A	15 A	15 A	15 D
16 C	16 D	16 A	16 A	16 C			16 C	16 D	16 B	16 C	16 B
17 D	17 C	17 B	17 C	17 D			17 C	17 A	17 A	17 C	17 D
18 D	18 D	18 A	18 B	18 C			18 B	18 C	18 D	18 D	18 D
19 D	19 D	19 B	19 D	19 D			19 A	19 C	19 B	19 A	19 C
20 A	20 C	20 B	20 B	20 A			20 D	20 B	20 D	20 D	20 B
21 B	21 D	21 C	21 C	21 D			21 C	21 D	21 D	21 D	21 A
22 B	22 B	22 D	22 D	22 D			gain	22 A	22 A	22 A	22 A
23 A	23 A	23 C	23 C	23 B			lose	23 C	23 C	23 C	23 A
24 AT	24 C	24 B	24 B	24 A			cation	24 B	24 B	24 B	24 A
25 NT	25 A			25 C			definite proportions	25 D			25 A
				26 C			2 -	26 C			26 A
				27 A			NH4+	27			27
				28 B			magnesium nitrate	28			28
				29 B			"-ide"	29			29
				30 C			Ca3(PO4)2	30			30
							more	31			31

Name _____

Date _____

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ATOMIC STRUCTURE

25. Complete this table.

Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
9	19	9	10	9
7	14	7	7	7
20	41	20	21	20
13	27	13	14	13
26	58	26	30	26

26. List the number of protons, neutrons, and electrons in each of the following atoms.

	Protons	Neutrons	Electrons
$^{12}_6\text{C}$	6	7	6
$^{10}_5\text{B}$	4	6	4
$^{20}_{10}\text{Ne}$	10	10	10
$^{11}_5\text{B}$	5	6	5
$^{32}_{16}\text{S}$	16	17	16

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CHEMICAL NAMES AND FORMULAS

Chapter Test B

D. Questions

Answer the following questions in the space provided.

82. For each pair of ions listed, write the correct formula and then name the compound formed by combining the two ions.

	Formula	Name
a. Ca^{2+} and NO_3^-	$\text{Ca}(\text{NO}_3)_2$	
b. Na^+ and SO_4^{2-}	Na_2SO_4	
c. Fe^{3+} and O^{2-}	Fe_2O_3	
d. Al^{3+} and CO_3^{2-}	$\text{Al}_2(\text{CO}_3)_3$	

83. Write the formulas for the following compounds.

a. silicon dioxide	SiO_2	e. nitric acid	HNO_3
b. carbon tetrafluoride	CF_4	f. silver nitrate	AgNO_3
c. zinc hydrosulfide	$\text{Zn}(\text{OH})_2$	g. ferric sulfide	
d. phosphorus tribromide	PBr_3	h. mercury(II) chloride	HgCl_2

84. Name the following compounds:

a. CS_2	Carbon disulfide	e. $\text{Sn}(\text{OH})_4$	Tin (IV) hydroxide
b. $(\text{NH}_4)_2\text{CO}_3$	Ammonium carbonate	f. H_2SO_4	Sulfuric acid
c. As_2O_5	Diphosphorus pentoxide	g. P_4	Phosphorus pentoxide
d. CO	Carbon monoxide	h. KMnO_4	Potassium permanganate

F F-34V

Naming Compounds Quiz

Name _____ Date _____ Period _____

Write the correct formula for the following compounds.

- 1) Calcium phosphate
- 2) Strontium perchlorate
- 3) Titanium(II) chloride
- 4) Iron(II) iodide
- 5) Aluminum sulfide
- 6) Tin(IV) permanganate
- 7) Sodium phosphate
- 8) Zinc cyanide
- 9) Barium chromate
- 10) Lithium chlorate
- 11) Potassium dichromate
- 12) Cobalt(II) acetate
- 13) Beryllium bromide
- 14) Cesium oxide
- 15) Chromium(VI) hydroxide
- 16) Krypton tetraiodide
- 17) Disulfur pentoxide
- 18) Hexachlorine difluoride
- 19) Carbon tetrachloride
- 20) Diphosphorous trioxide

$\text{Ca}_3(\text{PO}_4)_2$
 $\text{Sr}(\text{ClO}_4)_2$
 TiCl_2
 FeI_2
 Al_2S_3
 $\text{Sn}(\text{MnO}_4)_4$
 Na_3PO_4
 $\text{Zn}(\text{CN})_2$
 BaCrO_4
 LiClO_3
 $\text{K}_2\text{Cr}_2\text{O}_7$
 $\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2$
 BeBr_2
 Cs_2O
 $\text{Cr}(\text{OH})_6$
 KrI_4
 S_2O_5
 Cl_6F_2
 CCl_4
 P_2O_3

Name the following compounds.

- 21) CaBr_2
- * 22) $\text{Cr}_2(\text{SO}_4)_3$
- 23) $\text{Mg}(\text{OH})_2$
- * 24) TiF_4
- 25) $\text{Ba}(\text{NO}_3)_2$
- * 26) SnO_2
- 27) BiPO_4
- * 28) HgCl_2
- 29) SnS
- * 30) TeO
- 31) $(\text{NH}_4)_2\text{CO}_3$
- 32) KMnO_4
- 33) $\text{Al}(\text{CN})_3$
- 34) Li_2SO_4
- 35) AgCl
- 36) NF_3
- 37) CS_2
- 38) P_4O_{10}
- 39) SCl_6
- 40) XeBr_8

Calcium bromide
Chromium(III) sulfate
Magnesium hydroxide
Titanium(IV) fluoride
Barium nitrate
Tin(IV) oxide
Bismuth phosphate
Mercury(I) hypochlorite
Strontium sulfide
Iron(II) oxide
Ammonium carbonate
Potassium permanganate
Aluminum cyanide
Lithium sulfate
Silver chloride
Nitrogen trifluoride
Carbon disulfide
Tetraphosphorus decoxide
Sulfur hexachloride
Xenon octabromide