

6 C Carbon 12.0107	2 He Helium 4.002602	25 Mn Manganese 54.938045	53 I Iodine 126.90447	16 S Sulfur 32.065	69 Tm Thulium 168.93421	86 Rn Radon [222]	39 Y Yttrium 88.90585
	91 Pa Protactinium 231.03588	6 C Carbon 12.0107	19 K Potassium 39.0983	63 Eu Europium 151.964	69 Tm Thulium 168.93421		

Name: _____ Period: _____

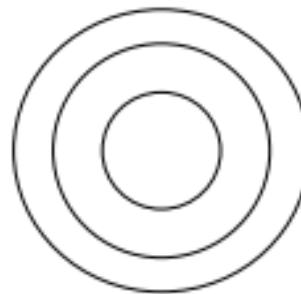
Targets

Students will evaluate and identify an unknown substance in terms of its physical and chemical properties

Students will develop a model that illustrates the arrangement of electrons using the Bohr Model and the Lewis Dot Diagram

Bohr Model Notes

- Used to show arrangement of electrons
- Electrons are placed on the _____ shell first
- Once full, extra electrons are placed in the next shells
- Maximum number of electrons on shells
 - 1st = _____ electrons
 - 2nd – _____ electrons
 - 3rd – _____ electrons
 - After element 18, shells fill differently so you might hear 2, 8, 18 in higher level chemistry classes
 - After element 18 the shells fill differently
- Periodic Table Tips
 - The periodic table can help you quickly complete the Bohr model
 - The number of periods shows you the number of _____
 - The number of groups shows you the number of _____ in the outer shell
- Bohr Model Practice
 - Fill in the blanks below.
 - Phosphorus
 - Symbol: _____
 - Atomic Number: _____
 - Atomic Mass: _____
 - Protons: _____
 - Neutrons: _____
 - Electrons: _____



Bohr Model Worksheet

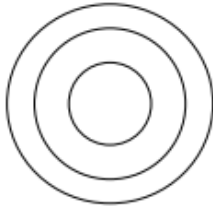
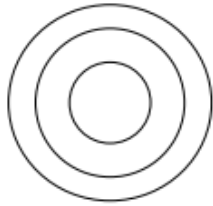
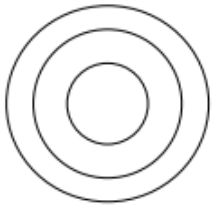
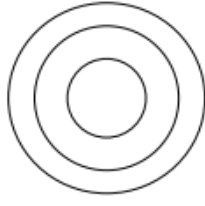
Use the description sheet and the periodic table to help you complete the following Bohr models.

1. How many electrons can each shell hold?

a. 1st = _____

b. 2nd = _____

c. 3rd = _____

Element	Atomic #	Atomic Mass	Protons	Neutrons	Electrons	Bohr Model
Carbon	6	12	6	6	6	
Hydrogen	1	1				
Lithium	3		3		3	
Magnesium	12	24				
Boron	5	11				