

## Balancing Equations Test

### Terms

1. when one element replaces another element in a compounds, the reaction is \_\_\_\_\_.
2. a reaction in which thermal energy is released \_\_\_\_\_.
3. Each substance to the right of the arrow in the chemical reaction is a \_\_\_\_\_.
4. numbers that precede (come before) symbols and formulas in a chemical equation are called \_\_\_\_\_.
5. A chemical reaction in which tow or more substances combine to form another substance is called a \_\_\_\_\_ reaction,
6. Each substance to the left of the arrow is called \_\_\_\_\_.
7. The arrow in a chemical equation means \_\_\_\_\_.
8. The symbol (aq) indicates the element or compound is a \_\_\_\_\_
9. When the positive part of one compound unites with the negative part of another compound, the reaction is a \_\_\_\_\_.
10. When one element displaces another element in a compound, the reaction is called a \_\_\_\_\_.
11. Changes that take place during a chemical reaction are shown in a \_\_\_\_\_.
12. The term that means “putting together” is \_\_\_\_\_?

### Balanced or not

13.  $\text{Mg}(\text{NO}_3)_2 + \text{K}_3\text{PO}_4 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{KNO}_3$
14.  $2\text{S} + 3\text{O}_2 \rightarrow 2\text{SO}_3$
15.  $3\text{KI} + \text{Pb}(\text{NO}_3)_2 \rightarrow 3\text{KNO}_3 + \text{PbI}_2$
16.  $2\text{Na} + \text{I}_2 \rightarrow 2\text{NaI}$
17.  $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
18.  $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow 2\text{Fe}_3\text{O}_2 + \text{H}_2$

### Identify the type

19.  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$
20.  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
21.  $2\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$
22.  $\text{Na}_2\text{S} + 2\text{Ag}_2\text{NO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$

23. **Balance the following equation the type of reaction.**

