

pHooey! HW

Read and outline (Cornel Style) **Section 19.3** in your chemistry textbook. ****Your section outlines must be at least ½ page with 3-4 topic questions**** Then answer the following assessment questions:

1. What is the relationship between the pH of a solution and the concentration of hydrogen ions in the solution?
2. If you know the pOH of a solution, how can you determine its pH?
3. Use your calculator to figure out the pH of each of the following solutions:
 - a. $[\text{H}^+] = 0.0014\text{M}$
 - b. $[\text{H}^+] = 6.0 \times 10^{-8}\text{M}$
 - c. $[\text{H}^+] = 4.2 \times 10^{-11}\text{M}$
 - d. $[\text{H}^+] = 1.5 \times 10^{-1}\text{M}$
 - e. $[\text{H}^+] = 2.4 \times 10^{-13}\text{M}$
 - f. $[\text{OH}^-] = 5.0 \times 10^{-4}\text{M}$
4. Which of the following solutions in question 3 are acids?