



## AP CHEMISTRY SUMMER HOMEWORK

The summer assignments for AP Chemistry will be a review of material you covered in Honors Chemistry last year. Since we are not able to pick up textbooks I have found an AP approved text on OPENSTAX. The link is shared below. All notes from Honors Chemistry last year should also help. The goal of this homework is to prepare you for AP Chemistry. You will not be asked to cover new material on your own. If you have any questions about homework assignments feel free to email me at the following:

[dgolshan@pittsburg.k12.ca.us](mailto:dgolshan@pittsburg.k12.ca.us) or

[golshanpitt@gmail.com](mailto:golshanpitt@gmail.com) or send me a message on

REMIND- AP Chemistry 20-21@b89d296

Make sure you register online to use the free copy of the openstax Chemistry2e textbook: It is a great reference and will help you with any questions you have with last year's material.

<https://openstax.org/details/books/chemistry-2e>

**Complete the following assignments by 8/12/20. The OpenStax reading will help with the problems assigned from the PDF document.**

Chapter reading from OpenStax Chemistry 2e	Exercises from PDF
1	<b>Assignment #1</b> p. 33 # 1, 3, 4, 5, 7, 8, 14, 16, 20, 30, 37, 41
2	<b>Assignment #2</b> p. 73 # 3, 4, 6, 17, 21, 24, 25, 30, 35, 37, 42, 49, 51, 56, 62, 70, 74, 80.
3	<b>Assignment #3</b> p. 112 # 5, 6, 12, 14, 20, 22, 26, 36, 38, 40, 48, 51, 54, 60, 62, 64, 78, 83

All exercises will be **checked and graded**.

- **Please make a SEPARATE packet or google doc for each problem set. They will be collected and graded individually.**

Each chapter will be **briefly** reviewed the first week of class.

**Other things to memorize:**

**SNAP solubility rules**

**The following ions are soluble in water**

<b>Nitrate</b>	<b><math>\text{NO}_3^{-1}</math></b>
<b>Sodium</b>	<b><math>\text{Na}^{+1}</math></b>
<b>Ammonium</b>	<b><math>\text{NH}_4^{+1}</math></b>
<b>Potassium</b>	<b><math>\text{K}^{+1}</math></b>

**Ion Charges from the periodic table (you should know how to find charges using only a periodic table)**

**Polyatomic Ions listed below:**

<b>Hydroxide</b>	<b><math>\text{OH}^{-1}</math></b>
<b>Nitrate</b>	<b><math>\text{NO}_3^{-1}</math></b>
<b>Acetate</b>	<b><math>\text{C}_2\text{H}_3\text{O}_2^{-1}</math></b>
<b>Cyanide</b>	<b><math>\text{CN}^{-1}</math></b>
<b>Permanganate</b>	<b><math>\text{KMnO}_4^{-1}</math></b>
<b>Carbonate</b>	<b><math>\text{CO}_3^{-2}</math></b>
<b>Sulfate</b>	<b><math>\text{SO}_4^{-2}</math></b>
<b>Dichromate</b>	<b><math>\text{Cr}_2\text{O}_7^{-2}</math></b>
<b>Phosphate</b>	<b><math>\text{PO}_4^{-3}</math></b>
<b>Ammonium</b>	<b><math>\text{NH}_4^{+1}</math></b>