

Name: \_\_\_\_\_

Period: \_\_\_\_\_

## Ion Formation Practice

<b>Name</b>	<b>Symbol</b>	<b>Group Number</b>	<b>Valence Electrons</b>	<b>Ion Formed</b>
Oxygen				
	Mg			
Sodium				
	F			
Argon				
	Al			
Nitrogen				
	Br			
	K			
Neon				
Calcium				
	B			
	Sr			
Sulfur				

Directions: Complete the table below by using the formula of each compound to identify the elements that each compound contains and the number of atoms of each of these elements in a unit of the compound.

Formula	Element 1	Element 2	Element 3
CO <sub>2</sub>			-----
HClO <sub>4</sub>			
Pb(NO <sub>3</sub> ) <sub>2</sub>			
H <sub>2</sub> (SO <sub>4</sub> ) <sub>2</sub>			
C <sub>6</sub> H <sub>6</sub>			-----
C <sub>6</sub> H <sub>8</sub> O <sub>12</sub>			

Directions: Draw Electron Dot Diagrams for the elements listed.

a. Carbon

e. Argon

b. Sulfur

f. Phosphorus

c. Chlorine

g. Magnesium

d. Helium

h. Aluminum